



Metal-Free Nitrogen-Doped Carbon Foam Electrocatalysts for the Oxygen Reduction Reaction in Acid Solution

J. Liu,^{a,b} S. Yu,^c T. Daio,^{a,d} M. S. Ismail,^e K. Sasaki,^{a,c,d,f} and S. M. Lyth^{e,f,*}

^aFaculty of Engineering, Kyushu University, Fukuoka 819-0395, Japan

^bCollege of Energy and Mechanical Engineering, Shanghai University of Electric Power, Shanghai 200090, People's Republic of China

^cInternational Research Center for Hydrogen Energy, Japan

^dNext-Generation Fuel Cell Research Center (NEXT-FC)

^eDepartment of Mechanical Engineering, University of Sheffield, South Yorkshire S3 7RD, United Kingdom

^fInternational Institute for Carbon-Neutral Energy Research (WPI-I2CNER), Kyushu University, Nishi-ku, Fukuoka, Japan

Metal-free, nitrogen-doped carbon foam is utilized as a model non-precious electrocatalyst system to investigate the role of nitrogen in the oxygen reduction reaction (ORR) in the absence of iron contamination. This graphene-like foam displays relatively high activity for the ORR in acid, despite being proven free from transition-metal impurities. The onset potential is 0.85 V_{RHE}, the mass activity is 2.8 A/g at 0.6 V_{RHE}, and the current density is -4.0 mA/cm². The maximum electron transfer number is calculated to be 3.6, revealing that a 4-electron pathway is possible in nitrogen-doped carbon, even in the absence of transition-metal coordination sites. The excellent electrochemical activity is attributed to the large surface area (700 m²/g), improved conductivity after graphitization, and the relatively high proportion of tertiary (graphite-like) nitrogen.

© The Author(s) 2016. Published by ECS. This is an open access article distributed under the terms of the Creative Commons Attribution 4.0 License (CC BY, <http://creativecommons.org/licenses/by/4.0/>), which permits unrestricted reuse of the work in any medium, provided the original work is properly cited. [DOI: 10.1149/2.0631609jes] All rights reserved.

Manuscript submitted May 20, 2016; revised manuscript received June 29, 2016. Published July 12, 2016.

Non-platinum group metal (non-PGM) electrocatalysts for the oxygen reduction reaction (ORR) are desired in order to reduce the cost and improve the durability of proton exchange membrane fuel cells (PEMFCs). Transition metal-based electrocatalysts have long been suggested as alternative electrocatalysts. Starting in the 1970's, early attempts were made to utilize transition metal-containing macrocyclic complexes as biomimetic ORR electrocatalysts, taking inspiration from their usefulness in nature.¹ However, soon it was discovered that heat-treatment of these macrocycles resulted in better activity and durability,² leading to debate over the precise chemical composition of the active site. Later still, it was realized that the macrocyclic complex was not necessary at all, but that seemingly arbitrary mixtures of Fe (or Co), nitrogen and carbon could lead to very active electrocatalysts.³ Since then, a vast amount of research output has been produced from a number of groups across the world.⁴⁻⁸ In recent years, such electrocatalysts have become a mainstream research topic, with the DOE setting specific targets for a non-PGM cathode in a H₂/air PEMFC; namely a volumetric activity of 300 A/cm³ measured at 800 mV_{iR-free} by 2017.⁹ These targets have been all but met,¹⁰ and non-PGM electrocatalysts are now commercially available.

The issue of the nature of the active site has never gone away, and is still frequently debated.¹¹ Complicating the issue is the fact that the ORR can proceed via either a 4- or 2-electron pathway to produce water or hydrogen peroxide (which can then further decompose to water), respectively. The 4-electron pathway is most desired in fuel cells, but the 2-electron pathway is significantly present in most non-precious electrocatalysts. The consensus in recent years has drifted toward the active site being comprised of pyridinic nitrogen atoms near a defect or edge, coordinated with a transition metal atom, with strong evidence of this provided by Mossbauer spectroscopy.¹² The clustering of nitrogen atoms in this manner has been observed in graphene in high resolution scanning tunneling microscopy (STM) measurements, and is possible because this is a low energy atomic configuration.¹³ However, other models have been proposed describing a metal-free active site,¹⁴⁻¹⁷ in which, for example, tertiary bonded electronegative nitrogen atoms (near edges or defects) withdraw electrons from neighboring edge carbon atoms, increasing their positive charge.¹⁸⁻²¹ This increases the binding energy with incoming oxygen molecules.

A third situation is also likely, in which a plurality of different configurations of ORR active sites for the ORR exist, and that both of the above proposed models are active.

Here, we focus on the second mechanism, which has received much less experimental attention in the literature. It must here be noted that metal-free electrocatalysts have been widely studied as ORR electrocatalysts in *alkaline* media where the cathode kinetics are much faster and non-PGM electrocatalysts can compete easily with Pt.^{15,22-25} However, this research is only relevant to alkaline fuel cells (AFCs), which are promising, but much further behind in commercial development than PEMFCs. The study of metal-free nitrogen-doped carbon electrocatalysts in acid media may have been neglected due to the much lower and therefore less "exciting" electrochemical activities that can be expected compared to their transition-metal-containing counterparts, or by measuring in alkaline medium. An additional issue is the difficulty in ensuring that samples are truly metal-free. Common carbon-based materials used in such research are carbon nanotubes or carbon black, both of which contain significant amounts of transition metal contamination.^{26,27} This fact is often not sufficiently discussed in such "metal-free" studies.

In 2009, our group synthesized carbon nitride (g-C₃N₄) as the ultimate metal-free nitrogen-doped carbon, and measured the ORR activity in acid medium.²⁸ Although the activity was greater than that of carbon black, the current density was relatively bad due to the low surface area (<10 m²/g), and poor electronic conductivity. We successfully increased the current density by supporting carbon nitride on carbon black,²⁹ or carbon nanotubes,²⁸ and by heat-treatment. However, metal contamination was a constant issue and could not be discounted. To avoid the presence of transition metals, our group began synthesis of nitrogen-doped graphene foam via thermal decomposition of sodium alkoxides and subsequent heat-treatment.^{30,31} Due to the high surface area and unique structure of these electrocatalysts, the ORR currents achieved were much higher than those previously obtained. Due to the purity of the precursors, and the synthesis in Teflon vessels, such samples are proven to be free of metal contamination.

Despite the difficulties involved, several other groups have made attempts to study metal-free nitrogen-doped carbon materials for the ORR in acid. For example, Sugimoto et al. studied pyrolyzed silk, and reported surface areas of 1018 m²/g, onset potentials of 0.83 V_{RHE}, and current densities of -5 mA/cm².^{32,33} They concluded that the large mesopore volume and surface area contributed significantly to the high

*Electrochemical Society Member.

^zE-mail: lyth@i2cner.kyushu-u.ac.jp

activity, whilst tertiary (graphite-like) nitrogen plays an important role in 4-electron ORR. Chokai et al. prepared metal-free electrocatalysts with large surface area (258 m²/g) by pyrolysis of polyimide, with onset potentials of 0.8 V_{RHE}.³⁴ Rao et al. synthesized nitrogen-doped carbons with unspecified surface area utilizing sacrificial alumina templates. The best sample had 8.4 at% nitrogen, displayed an onset potential of 0.455 V_{Ag/AgCl} (~0.67 V_{RHE}), and a current density of -0.53 mA/cm². The catalytic activity was attributed to pyridinic nitrogen.³⁵ Yu et al. synthesized metal-free nitrogen-doped single wall carbon nanotubes (with 3.6 at% nitrogen) via a plasma etching process and achieved an onset potential of 0.5 V_{Ag/AgCl} (~0.71 V_{RHE}) and a current density of -2 mA/cm², also attributed to pyridinic-type nitrogen.²⁴ Finally, Wang et al. synthesized ammonia-treated mesoporous carbons achieving a surface area of 1923 m²/g, a nitrogen content of 4.6 at%, an onset potential of 0.72 V_{NHE} (i.e. ~0.72 V_{RHE}), and a current density of -4.2 mA/cm².³⁶ These are probably the best values published for a metal-free electrocatalyst in acid media to date. Some other groups have reported “metal-free” ORR electrocatalysts, but admit to up to 0.5 wt% iron contamination.^{22,37,38} The history and recent advances in Fe/N/C and their metal-free counterpart electrocatalysts have been presented in great detail in several recent reviews.^{11,39,40}

Here, we expand on our previous research by performing high temperature graphitization of optimized nitrogen-doped carbon foam samples, and exploring how this affects the electrochemical activity.

Experimental

Synthesis of nitrogen-doped carbon foam.—Chemicals were used as received without further purification. Nitrogen-doped carbon foam was synthesized via combustion of nitrogen-containing sodium alkoxide. Anhydrous ethanol (Japanese pharmacopeia grade, Wako, >99.5%) and diethanolamine (Japanese Industrial Standards (JIS) special grade, Wako, >99%) were both added to a Teflon reactor, followed by an equimolar amount of sodium (Sigma-Aldrich, >99%). The reactor was sealed, heated to 150°C, and stirred magnetically overnight, before being cooled to room temperature. The resulting alkoxide powder was ignited, burning with a steady flame to form a black powder. This was crushed, sonicated, and washed in deionized water to remove residual sodium oxides, and then dried under vacuum.

The first sample was heat treated at 1000°C for 2 hours under nitrogen flow, and the again washed in deionized water. A second calcination step was performed at 800°C for 2 hours under 5 sccm H₂ and 95 sccm N₂ gas flow. This sample is herein referred to as CN_x-1000.³¹ The second sample was subjected to the same process, and then graphitized at 1400°C under N₂ flow for 15 hours. This sample is herein referred to as CN_x-1400. Great care was taken throughout to avoid contact with any transition metal species through either cross contamination, or even by contact with stainless steel instruments. Despite this, some trace impurities may present in the reactants, especially in the sodium metal.

Materials characterization.—Samples were characterized by scanning electron microscopy (SEM, S-5200, Hitachi, Japan); transmission electron microscopy (TEM, HD-2300, Hitachi, Japan); BET surface area analysis (Belsorp mini II-VS, Bel Japan, Inc.); X-ray photoelectron spectroscopy (XPS, ESCA-3400, Kratos Analytical Ltd., UK); and inductively coupled plasma - atomic emission spectrometry (ICP-AES, Shimadzu Corp., ICPE-9000).

Electrode preparation.—Working electrodes were prepared according to the methods recommended by the Fuel Cell Commercialization Conference of Japan (FCCJ). Electrocatalysts were dispersed in a mixture of deionized water, ethanol (Wako) and 5 wt% Nafion solution (Wako) in a volume ratio of 3:3:1, then sonicated for 30 min. A glassy carbon-disk platinum-ring electrode (0.196 cm², Hokuto Denko Corp.) was coated with 8 μL of catalyst ink, then dried at 60°C for 15 min. The resulting density of the electrocatalyst material on the glassy carbon electrode was 0.58 mgcm⁻².

Cyclic voltammetry (CV) and linear sweep voltammetry (LSV).—Electrochemical measurements were performed using a typical rotating ring-disk electrode (RRDE) system (HX-107, Hokuto Denko Corp., Japan), in N₂-saturated 0.1 M perchloric acid (Aldrich). All potentials were calculated relative to the reversible hydrogen electrode (RHE). The working electrode was a glassy-carbon-disk platinum-ring electrode, the counter electrode was a platinum wire, and Ag/AgCl was used as the reference electrode. Cyclic voltammograms (CV) were obtained in the range 0.05 to 1.2 V_{RHE}, with a scan rate of 50 mVs⁻¹ in N₂-saturated solution. Linear sweep voltammograms (LSV) were performed in the range 0.2 to 1.2 V_{RHE}, with a scan rate of 10 mVs⁻¹ in O₂ saturated electrolyte at 400, 900, 1600, and 2500 rpm. Onset potentials were determined using 1st derivative plots. The kinetically controlled current, *i_k*, is determined from the intercept of Koutecky-Levich plots. The mass activity at 0.60 V_{RHE} is based on ORR measurements at different rotating speeds. In general the mass activity is measured at 0.9 V_{RHE}, but since this would be zero in the case of most metal-free catalysts, 0.60 V_{RHE} is used for comparison of these samples. The electron transfer number (*N*) was obtained from the ring current (*I_R*), the disk current (*I_D*), and the H₂O₂ collection efficiency (*n* = 0.4735), using Equation 1;

$$N = 4I_D / [(I_R/n) + I_D] \quad [1]$$

Results and Discussion

Scanning electron microscopy (SEM) and transmission electron microscopy (TEM).—The microstructure of the samples was characterized by SEM (Figure 1). Figure 1a highlights the three-dimensional structure of a fragment of CN_x-1000. A spheroidal micron-scale pore / void is observed at the bottom of the fragment. Large, flat, graphene-like areas of carbon are visible, some of which protrude out-of-plane. Holes in the sheets are also present, as are many edges. CN_x-1000 and CN_x-1400 both have almost identical structure (Figures 1b, 1c); an interconnected, highly porous, three dimensional carbon framework with micron-scale macropores separated by thin carbon walls. The unique structure prevents restacking of the graphene-like sheets and subsequent loss of surface area. Restacking generally restricts the surface area of commercially available graphene powders to around 600 m²/g. The large pores in this carbon foam are expected to aid mass transfer of reactants and products through the electrocatalyst layer. TEM images (Figures 1d, 1e) show higher magnification images at the edge of the graphene-like sheet, revealing a multilayered structure, and suggesting that the large sheets are essentially few-layer nitrogen-doped graphene, albeit without the highly crystalline chemical structure.

BET surface area and pore size distribution.—The specific surface area and average pore size were calculated using the BET method, based on nitrogen adsorption/desorption isotherms. CN_x-1000 has a surface area of 705 m²/g, and CN_x-1400 has a very similar value of 700 m²/g. The average pore size of CN_x-1000 is 2.0 nm, whilst CN_x-1400 has a slightly higher value of 2.5 nm. This suggests that graphitization of this material slightly increases pore size, but does not have a large impact on the overall microstructure. This is important in the interpretation of the electrochemical results, since changes in microstructure due to heat-treatment can be discounted in this case. The theoretical maximum surface area of single-layer graphene is 2630 m²/g. Dividing that value by the surface area of our samples gives a rough estimate of the number of layers, namely 3.7 in this case. This is consistent with AFM results previously reported by our group on similar samples.³¹ Figure 2 shows the pore size distribution; CN_x-1000 has a majority of pore sizes smaller than 5 nm in radius, whilst they are slightly bigger in the graphitized CN_x-1400 sample.

Inductively charged plasma (ICP).—The aim of this work is to investigate the ORR in metal-free nitrogen-doped carbon. Therefore, it is imperative that the presence of iron in the samples be ruled out. The sensitivity of e.g. XPS is insufficient to detect low concentrations

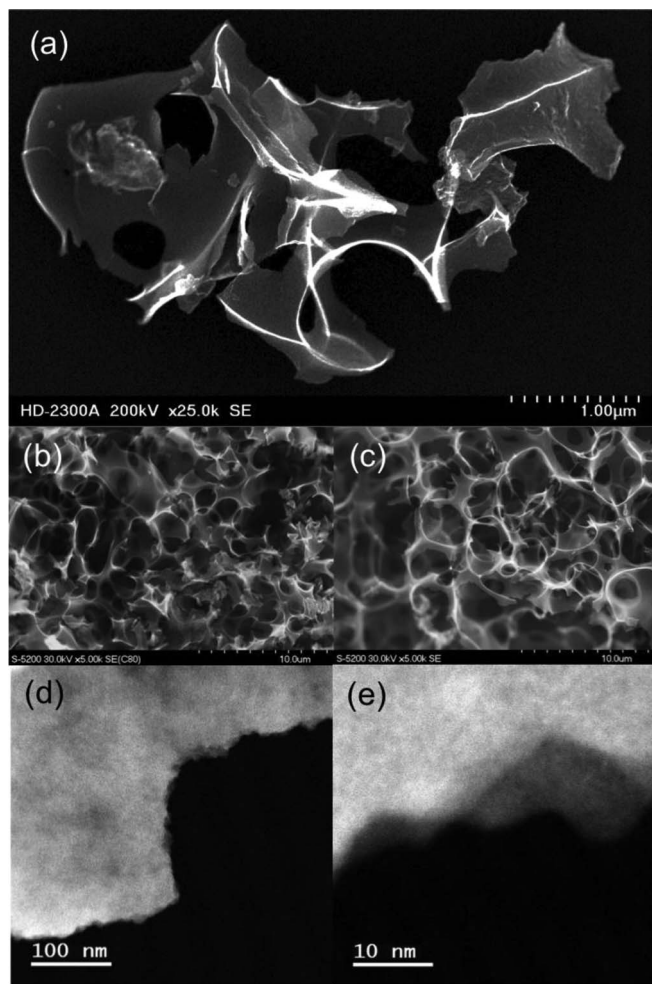


Figure 1. SEM images of (a,b) CN_x-1000, and (c) CN_x-1400. (d,e) TEM images of CN_x-1400.

of metallic elements (<1 at%). However low concentration metal inclusions still have a large impact on catalytic activity. ICP is very sensitive to metal cations, and therefore this technique was used. 4.2 mg GF_N was mixed with 20 ml of concentrated HCl overnight. 3 ml of the resulting solution was made up to 30 ml with deionized water, and this was analyzed. The resulting concentration of Fe⁺ ions was below the detection limit of the machine (1 ppb), confirming that the iron content in GF_N is negligible. Other electrochemically active trace impurities such as Ni or Co may be contained e.g. in the sodium metal, however these are generally accepted to be less active than iron, and less abundant.

X-ray photoelectron spectroscopy (XPS).—X-ray photoelectron spectroscopy (XPS) was performed for elemental analysis (Figure 3). The wide-scan spectrum (Figure 3a) shows a strong C 1s signal at 284.5 eV; a weak N 1s signal at ~400 eV; and a small O 1s signal at ~532 eV for both samples. The oxygen content is slightly reduced from 9.1 at% (CN_x-1000) to 8.5 at% after graphitization (CN_x-1400). The nitrogen content is decreased from 1.3 at% for CN_x-1000, to 0.65 at% for CN_x-1400. High temperature heat-treatment is therefore shown to decrease the amount of oxygen and nitrogen in the samples, as expected. Carbonyl bonds (C=O, 532.6 eV) are the major component in the O 1s signal for both samples.

The C 1s spectra are shown in Figures 3b and 3d, centered at 284.5 eV. The N 1s spectra (Figures 3d, 3e) are deconvoluted into five peaks; (1) amine functional groups at ~401.8 eV; (2) graphitic / tertiary nitrogen at ~400.8 eV; (3) pyrrolic nitrogen at ~399.8 eV; (4)

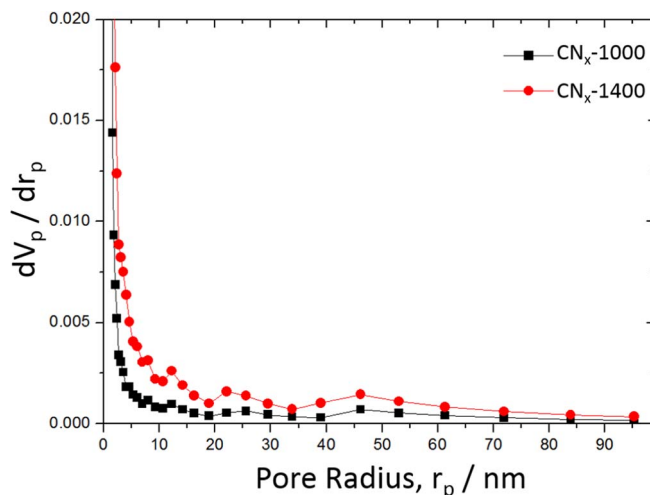


Figure 2. Pore size distributions for CN_x-1000 and CN_x-1400.

pyridinic nitrogen at ~398.9 eV; and a small shoulder (5) at ~397.8 eV possibly corresponding to imine groups.⁴¹ The main difference between the two samples is that the proportion of tertiary (graphitic) nitrogen relative to pyridinic nitrogen significantly increases after graphitization. Edge-based pyridinic nitrogen has only two bonds and is removed more easily by heat-treatment than graphitic nitrogen with three bonds in the basal plane. Thus, tertiary nitrogen is the main chemical state in CN_x-1400, whilst an approximately equal mix of tertiary and pyridinic nitrogen is present in CN_x-1000.

Cyclic Voltammetry (CV).—CV measurements were performed in N₂- and O₂-saturated 0.1 M HClO₄ solution (Figure 4). The double layer capacitance of CN_x-1000 is higher than CN_x-1400, despite the similar surface area. Therefore, the decrease in double layer capacitance may be attributed to the larger pores in CN_x-1400.⁴² Under N₂ flow, the CV curves are largely featureless, which is typical of carbon-based materials. Under O₂ flow, the ORR activity of CN_x-1400 is slightly improved compared with CN_x-1000.

Rotating ring-disk electrode voltammetry (RRDE).—RRDE measurements were performed in O₂-saturated 0.1 M HClO₄. The mass activity of CN_x-1000 (Fig. 5a) is 0.46 A/g at 0.6 V_{RHE}. CN_x-1000 does not show any mass diffusion limited current, however the maximum current density is -3.4 mA/cm² at 2500 rpm. The onset potential is around 0.8 V_{RHE} for CN_x-1000, calculated from the 1st derivative of the current density (*inset*). This is an unusual method for the determination of onset potential, but provides an accurate and consistent value compared with simply estimating the point where an arbitrary current density is reached. Whilst this is low compared with Pt (e.g. 1.08 V_{RHE}), it must be noted that this is a completely metal-free electrocatalyst. The electron transfer number ranges from around 2.5 at low current densities, to 3.3 at higher current density.

After graphitization, CN_x-1400 (Fig. 5b) has a mass activity of 2.79 A/g at 0.6 V_{RHE}, which is six times higher than that of CN_x-1000. A mass diffusion limited current is not quite reached, but saturation does start to occur (-4.0 mA/cm² at 2500 rpm). The gradients of the slopes are greater than for CN_x-1000, indicating lower resistance due to the graphitization step. The onset potential is ~0.85 V_{RHE} (*inset*). This is higher than CN_x-1000, and compares favorably even with many Fe/N/C-based electrocatalysts, despite the metal-free nature of this sample. It also indicates a significant proportion of oxygen reduction via the 4-electron ORR pathway, since the theoretical overpotential for 2-electron ORR is much lower, at 0.695 V_{RHE}. The electron transfer number varies from ~3.0 to 3.6, also indicating that the graphitization step results in an increase in the proportion of 4-electron transfer.

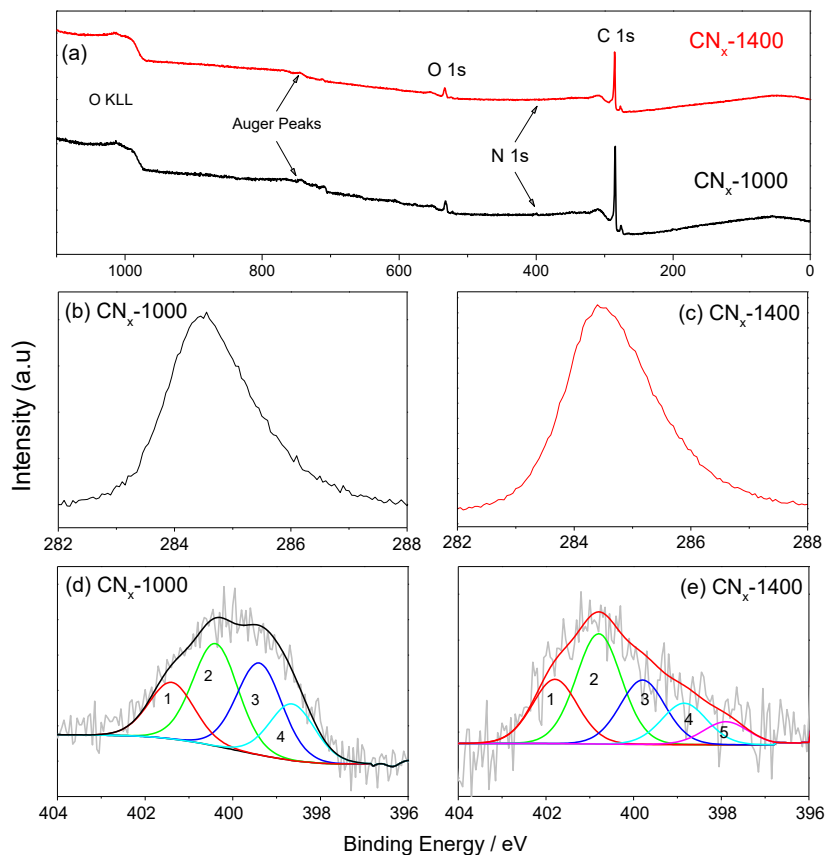


Figure 3. (a) XPS wide scan spectra. C 1s spectra for (b) CN_x -1000 and (c) CN_x -1400. N 1s spectra for (d) CN_x -1000 and (e) CN_x -1400.

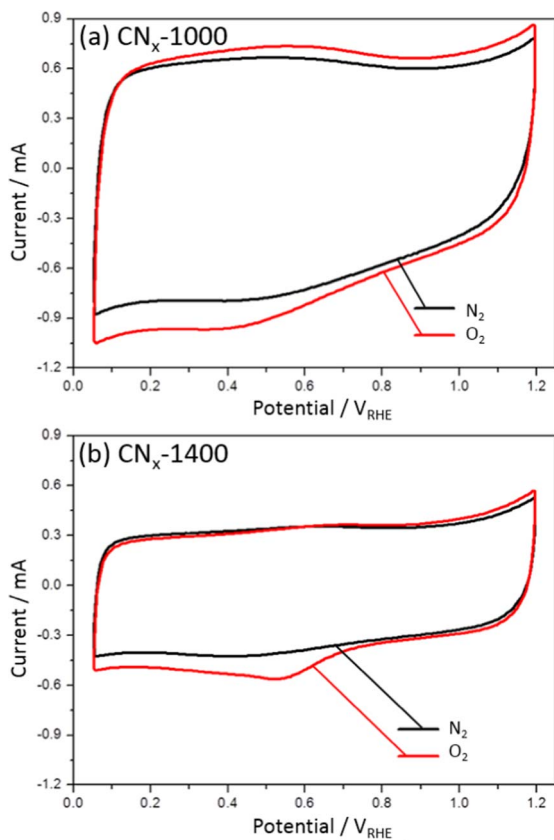


Figure 4. Cyclic voltammograms of (a) CN_x -1000 and (b) CN_x -1400, on a glassy carbon-disk / Pt-ring electrode in N_2 - and O_2 -saturated 0.1 M HClO_4 at 25°C, with a scan rate of 50 mV/s.

Discussion

Whilst it is widely accepted that iron-coordinated nitrogen active sites are active for the ORR with a 4-electron pathway, there is still debate about whether metal-free nitrogen-doped carbon can catalyze a 4-electron ORR pathway in acid medium. It is often assumed that such materials will mainly catalyze oxygen reduction in a 2-electron pathway to form hydrogen peroxide. Here however, we observe a majority 4-electron pathway, despite the proven metal-free nature of the samples. In addition, we observe high onset potential and current density compared with other “metal-free” electrocatalyst systems reported in the literature.^{32–34} The mass activities reported here even compare well with many reported Fe/N/C-based electrocatalysts,¹¹ although are clearly some way off competing with the best reported values.^{6,7} It is also interesting to note that for these Fe-free carbon foam electrocatalysts, the activity is much improved by graphitization at high temperature (in this case, 1400°C). This is in contrast to the case of the transition-metal containing electrocatalysts, in which the ORR activity peaks at e.g. 600 to 800°C and then is drastically reduced at higher temperature. This is further evidence of the fundamentally different nature of metal-free versus metal-containing catalysts for this reaction.

The reasons for the relatively high electrochemical activity achieved in our work are probably threefold. Firstly, the **surface area** is large, at 700 m²/g. This increases the triple phase boundary, thereby increasing the current density. Generally, much lower surface areas are reported for metal-free studies in the literature.³⁴ Additionally, the pore-size is very large compared with e.g. mesoporous carbons, resulting in increased mass transport, also a major factor in catalyst performance.⁴³ However, the higher activity of CN_x -1400 compared with CN_x -1000 cannot be explained by surface area, as they are both almost identical. The second **possible factor** leading to high electrochemical activity is therefore **conductivity**. Graphitization increases conductivity in carbon,^{44–46} in this case resulting in improved electron transport through the electrocatalyst layer. Evidence for this is in

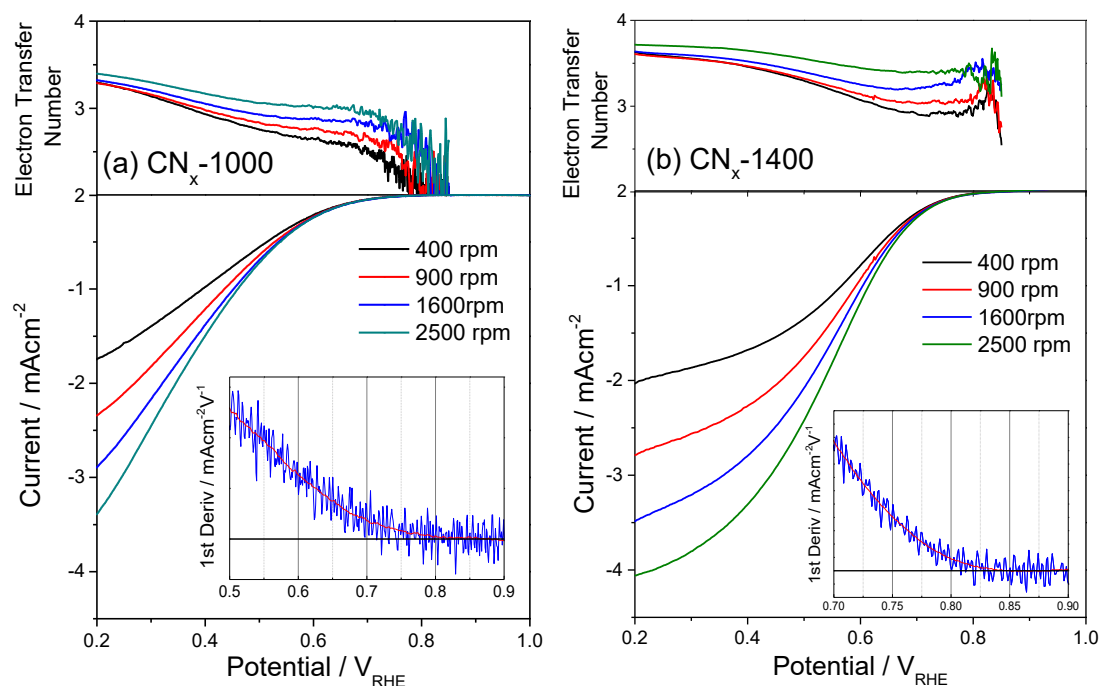


Figure 5. Rotating disk current density; electron transfer number; and 1st derivative of the onset region for: (a) CN_x-1000 and (b) CN_x-1400, in O₂-saturated 0.1 HClO₄ on a glassy carbon-disk Pt-ring electrode at 25°C, with a scan rate of 10 mV/s.

the steeper gradient of the LSV curves after graphitization, indicating lower resistance and reduced iR losses. The third possibility is that CN_x-1400 has a greater proportion of 4-electron pathway active sites, namely *tertiary nitrogen* sites. There is a significantly higher proportion of tertiary to pyridinic nitrogen in CN_x-1400, as evidenced by XPS measurements. This nitrogen configuration has been suggested to be active for the ORR in several theoretical models.^{18–20,47,48} *Our results provide some experimental verification of these models.* It should be noted that evidence of a 4-electron pathway in metal-free systems does not necessarily mean that such metal-free active sites are responsible for the ORR activity of Fe-containing electrocatalyst systems. The Fe-coordinated pyridinic nitrogen active site is still likely much more active than the Fe-free system. However, such metal-free tertiary nitrogen active sites may contribute significantly to the overall ORR current density in Fe/N/C-based non-precious electrocatalyst systems, especially at high current density (<0.6 V_{RHE}).

Conclusions

Nitrogen-doped carbon foams with large surface area (700 m²/g) were synthesized and pyrolyzed at different temperatures. These were applied as metal-free electrocatalysts for the ORR in acid, in order to shed light on the active site in non-PGM electrocatalysts. SEM images revealed the highly porous nature of the material. The proportion of tertiary bonded nitrogen increased significantly after graphitization at 1400°C. Heat treatment also significantly improved the electrochemical activity. An onset potential of 0.85 V_{RHE}, a mass activity of 2.79 A/g at 0.6 V_{RHE}, and a current density of -4 mA/cm² were recorded, and the maximum electron transfer number was 3.6. These results confirm that a 4-electron ORR pathway is indeed possible in nitrogen-doped carbon in the absence of Fe atoms. The increase in electron transfer number with the proportion of tertiary nitrogen suggests that this arrangement of nitrogen and carbon is responsible for the 4-electron oxygen reduction reaction in acid. Given the above findings, we are currently focusing on increasing the surface area, conductivity, and nitrogen content of our nitrogen-doped carbon foams in order to further improve the electrochemical activity of this unique material.

Acknowledgments

The authors gratefully acknowledge the support of the International Institute for Carbon Neutral Energy Research (WPI-I2CNER), sponsored by the World Premier International Research Center Initiative (WPI), MEXT, Japan; and the International Research Center for Hydrogen Energy, Kyushu University. This work was supported by JSPS KAKENHI Grant-in-Aid for Encouragement of Young Scientists (B), grant Number 15K17898; and a Royal Society of Chemistry Mobility Fellowship.

References

- H. Jahnke, M. Schönborn, and G. Zimmermann, *Top. Curr. Chem.*, **61**, 133 (1976).
- V. S. Bagotzky, M. R. Tarasevich, K. A. Radyushkina, O. A. Levina, and S. I. Andrusyova, *J. Catal.*, **2**, 8 (1973).
- S. Gupta, D. Tryk, I. Bae, W. Aldred, and E. Yeager, *J. Appl. Electrochem.*, **19**, 19 (1989).
- H. Wang, R. Côté, G. Faubert, D. Guay, and J. P. Dodelet, *J. Phys. Chem. B*, **103**, 2042 (1999).
- F. Jaouen, S. Marcotte, J.-P. Dodelet, and G. Lindbergh, *J. Phys. Chem. B*, **107**, 1376 (2003).
- F. Jaouen et al., *Energy Environ. Sci.*, **4**, 114 (2011).
- M. H. Robson, K. Artyushkova, W. Patterson, P. Atanassov, and M. R. Hibbs, *Electrocatalysis* (2013).
- Y. Nabaie et al., *ECS Trans.*, **25**, 463 (2009).
- H. A. Gasteiger, S. S. Kocha, B. Sompalli, and F. T. Wagner, *Appl. Catal. B Environ.*, **56**, 9 (2005).
- S. Stariha, A. Scrov, B. Halevi, and P. Atanassov, in *2014 ECS and SMEQ Joint International Meeting (October 5-9, 2014)*, ECS (2014).
- J.-P. Dodelet, in *Electrocatalysis in Fuel Cells*, Lecture Notes in Energy. D. Shao, D. MinHiggins, and Z. Chen, Editors, vol. 9, p. 271, Springer London, London (2013).
- A. Zitolo et al., *Nat. Mater.*, **14**, 937 (2015).
- Y. Tison et al., *ACS Nano* (2015).
- K. Wiesener, *Electrochim. Acta*, **31**, 1073 (1986).
- S. Maldonado and K. J. Stevenson, *J. Phys. Chem. B*, **108**, 11375 (2004).
- P. Matter, L. Zhang, and U. Ozkan, *J. Catal.*, **239**, 83 (2006).
- N. P. Subramanian et al., *J. Power Sources*, **188**, 38 (2009).
- T. Ikeda et al., *J. Phys. Chem. C*, **112**, 14706 (2008).
- H. Niwa et al., *J. Power Sources*, **187**, 93 (2009).
- F. Vallejos-Burgos et al., *Fuel*, **99**, 106 (2012).
- J. A. Menéndez, J. Phillips, B. Xia, and L. R. Radovic, *Langmuir*, **12**, 4404 (1996).
- L. Qu, Y. Liu, J.-B. Baek, and L. Dai, *ACS Nano*, **4**, 1321 (2010).
- K. Gong, F. Du, Z. Xia, M. Durstock, and L. Dai, *Science*, **323**, 760 (2009).

24. D. Yu, Q. Zhang, and L. Dai, *J. Am. Chem. Soc.*, **132**, 15127 (2010).
25. D. Geng et al., *Energy Environ. Sci.*, **4**, 760 (2011).
26. F. G. Lepri et al., *Talanta*, **81**, 980 (2010).
27. S. M. Lyth and S. R. Silva, *Appl. Phys. Lett.*, **90**, 173124 (2007).
28. S. M. Lyth et al., *J. Nanosci. Nanotechnol.*, **12**, 4887 (2012).
29. S. M. Lyth et al., *J. Electrochem. Soc.*, **158**, B194 (2011).
30. S. M. Lyth et al., *e-Journal Surf. Sci. Nanotechnol.*, **10**, 29 (2012).
31. J. Liu, D. Takeshi, D. Orejon, K. Sasaki, and S. M. Lyth, *J. Electrochem. Soc.*, **161**, F544 (2014).
32. T. Iwazaki, R. Obinata, W. Sugimoto, and Y. Takasu, *Electrochem. Commun.*, **11**, 376 (2009).
33. T. Iwazaki, H. Yang, R. Obinata, W. Sugimoto, and Y. Takasu, *J. Power Sources*, **195**, 5840 (2010).
34. M. Chokai et al., *J. Power Sources*, **195**, 5947 (2010).
35. C. V. Rao, C. R. Cabrera, and Y. Ishikawa, *J. Phys. Chem. Lett.*, **1**, 2622 (2010).
36. X. Wang et al., *Chem. Mater.*, **22**, 2178 (2010).
37. Z. Lin et al., *Phys. Chem. Chem. Phys.*, **14**, 3381 (2012).
38. O.-H. Kim et al., *Sci. Rep.*, **5**, 8376 (2015).
39. M. Shao, Q. Chang, J.-P. Dodelet, and R. Chenitz, *Chem. Rev.*, **116**, 3594 (2016).
40. N. Daems, X. Sheng, I. F. J. Vankelecom, and P. P. Pescarmona, *J. Mater. Chem. A*, **2**, 4085 (2014).
41. K. L. Tan, B. T. G. Tan, E. T. Kang, and K. G. Neoh, *J. Mater. Sci.*, **27**, 4056 (1992).
42. P. Simon and A. Burke, *Electrochem. Soc. Interface* (2008).
43. J. Park, Y. Nabaee, T. Hayakawa, and M. Kakimoto, (2014).
44. J. Liu, K. Sasaki, and S. M. Lyth, *ECS Trans.*, **64**, 1161 (2014).
45. M. Sevilla and A. B. Fuertes, *Carbon N. Y.*, **44**, 468 (2006).
46. X. Zhao et al., *Electrochim. Acta*, **97**, 33 (2013).
47. L. R. Radovic, *J. Am. Chem. Soc.*, **131**, 17166 (2009).
48. M. Kobayashi et al., *Electrochim. Acta*, **74**, 254 (2012).