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**Al-modified Pd@mSiO<sub>2</sub> core-shell catalysts for the selective hydrodeoxygenation of fatty acid esters: Influence of catalyst structure and Al atoms incorporation**

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**Abstract:** The development of robust catalysts for selective hydrodeoxygenation (HDO) of fatty acid esters is key for production of the diesel-range alkanes under mild conditions. Here, we report a stable and efficient HDO catalyst (Pd@Al<sub>x</sub>-mSiO<sub>2</sub>) that composed of inner uniformly dispersed metallic palladium core and the outer mesoporous silica shell doped with aluminum atoms. Using methyl palmitate as a model compound, Al-modified Pd@mSiO<sub>2</sub> catalyst (Pd@Al<sub>3</sub>-mSiO<sub>2</sub>) exhibited higher catalytic performance (98% conversion rate and 99% selectivity towards diesel-range alkanes) as compared with the Pd@mSiO<sub>2</sub> and conventional Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalysts (35% and 70% conversion rates, respectively) at 260 °C and 3.0 MPa H<sub>2</sub>. Further, when using vegetable oils such as soybean and palm oils as raw materials, high yields of diesel-range alkanes (> 80 wt%) can be obtained under the mild conditions. The HDO reaction pathway was more dominant than the decarbonylation pathway when the reaction was catalyzed by Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, thereby reducing the loss of carbon atoms. Detailed characterization (<sup>27</sup>Al NMR, NH<sub>3</sub>-TPD, and in situ Py-FTIR) suggests that the incorporation of aluminum atoms brings not only Lewis acid sites, but also Brønsted acid sites via the formation of

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Si-OH-Al bonds. The synergy between the metallic Pd, Lewis- and Brønsted-acid sites is responsible for its high HDO activity under the mild conditions. Additionally, the core-shell structure enables fatty aldehyde intermediate preferentially adsorbed on the aluminum atoms in the outer silica shell and avoids direct contacting with metallic Pd, which inhibits the cleavage of C-C bonds to some extent. On the other hand, due to the protective effect of outer silica shell that inhibits the leaching and agglomeration of metallic Pd, the synthesized Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst showed good stability with a slight loss (conversion decreased from 98% to 90%) over five cycles.

**Keywords:** Core-shell structure; Aluminum atoms incorporation; Selective hydrodeoxygenation; Diesel-range alkanes; Sintering resistant

## 1. Introduction

With the continuous depletion and consumption of fossil fuels, the development of renewable and clean energy production has become increasingly important. Biodiesel derived from biomass conversion has attracted great attention due to its reusability and low emission of pollutants [1, 2]. Currently, fatty acid methyl ester (FAME), as the first generation of biodiesel, has been produced with a large scale by transesterification and annual output exceeds 20 million tons [3]. However, high oxygen content and cold filter plugging point issues make it difficult to application as a high-grade transportation fuel [4]. The second-generation biodiesel derived from the direct conversion of fatty acid esters can solve the issues of the FAME because the as-prepared biodiesel is long-chain alkanes (C<sub>12</sub>-C<sub>22</sub>) and does not contain any oxygen element [5].

To date, two catalytic hydrodeoxygenation (HDO) catalysts have been extensively reported for production of diesel-range alkanes from fatty acid esters. The first employs sulfide transition metal catalysts, such as NiMoS<sub>4</sub> and CoMoS<sub>4</sub>, but strict environmental regulations on the sulfur content of

biofuels hinder its application [6-8]. The second relies on supported metal catalysts in a reduced state, such as supported Ni [9-11], Co [12, 13], Cu [14], Pd [15, 16], Pt [17] and Ir [18] catalysts. Compared with the non-precious catalysts with high metal loadings and harsh reaction conditions, noble catalysts seem to be attractive due to its high catalytic activity and mild reaction conditions. However, the problems of instability and deactivation due to the leaching and agglomeration of metal nanoparticles during the reactions greatly limit its development. Besides, these catalysts often favor the decarbonylation and/or decarboxylation, which leads to some carbon atoms loss [16]. Therefore, in the process of fatty acid esters conversion, developing a noble metal catalyst with good reusability and high HDO selectivity towards target products has turned into a great challenge.

To inhibit the leaching and agglomeration of noble metal catalysts, great efforts have been devoted to improve the catalyst stability and reusability, which includes developing catalysts with stronger electron interaction between nanoparticles and support [19, 20], embedding active sites into controlled cavities or pore channel of pre-synthesized molecular sieves [21, 22], and constructing core/yolk shell nanostructures by coating a layer of shell on the pre-synthesized nanoparticles [23-29]. Among these methods, designing a nanocatalyst into core/yolk shell structure is a promising method to improve catalyst stability, since it can prevent the migration and sintering of nanoparticles. On the other hand, it is known that catalyst acidity plays an important role in promoting the conversion of esters and improving the HDO selectivity of products [11, 30]. For example, the synergy between the hydrogenation metal and Lewis acid sites has been demonstrated for promoting the conversion of fatty acids/esters [18, 31], while Brønsted acid sites are beneficial to enhance the HDO selectivity of products [32]. However, traditional core-shell structured catalysts have weak acidity due to its neutral silicon framework, which greatly limits their application in the conversion of fatty esters.

Based on the above analysis and an inspiration of inserting aluminum atoms into the silica framework to increase catalyst acidity in the molecular sieve catalysts via formation of the Si-OH-Al bonds [33-35], a series of Pd@Al<sub>x</sub>-mSiO<sub>2</sub> core-shell catalysts with different framework aluminum content were synthesized and applied to convert fatty acid esters into diesel-range alkanes. To shed light on the structure-activity relationship, the catalyst structure was characterized by detailed characterization (XRD, BET, TEM, HRTEM and <sup>27</sup>AlNMR), and associated with their intrinsic catalytic activity. Additionally, several crucial parameters such as aluminum content, catalyst stability, reaction temperature and H<sub>2</sub> pressure were also investigated. To the best our knowledge, only few literatures were reported using core-shell-type catalysts for the conversion of fatty acid esters. This research provides a paradigm for the application of core-shell-type catalysts in the HDO of fatty acid esters.

## 2. Experimental section

### 2.1 Material

Cetyltrimethylammonium bromide (CTAB, 99%), tetraethyl orthosilicate (TEOS, 98%), ascorbic acid (>99%), sodium aluminate (NaAlO<sub>2</sub>, 98%), potassium tetrachloropalladate (K<sub>2</sub>PdCl<sub>4</sub>, 98%), methyl palmitate (95%), γ-Al<sub>2</sub>O<sub>3</sub> were all purchased from Shanghai Macklin Company. Soybean and palm oils were purchased from commercial sources. All chemicals were used without any further treatment.

### 2.2 Catalyst preparation

The Pd@Al<sub>x</sub>-mSiO<sub>2</sub> core-shell catalysts were synthesized through a facile one-pot method. Typically, CTAB (0.1 g), as a template agent due to its appropriate carbon chain length, was added into 45 mL deionized water and stirred at 95 °C for 0.5 h. Then, an aqueous solution of K<sub>2</sub>PdCl<sub>4</sub> (0.02 M, 5

mL) was slowly added to the mixture, and stirred at 95 °C for 1.0 h. Afterwards, ascorbic acid (0.1 M, 1.7 mL) was added, and the mixture gradually changed into black. After stirring at 95 °C for 1.0 h, a NaOH solution was introduced into the above mixture to adjust the pH to around 10-11, and then 0.4 mL of TEOS was added as silica precursor. The mixture was stirred at 95 °C for 3.0 h, at which Pd core coated with silica shells was formed. Subsequently, a certain amount of NaAlO<sub>2</sub> solution (0.05 M) was introduced into the above mixture and was stirred for 1.0 h. The gray product was collected by centrifugation, repeatedly washed by using the distilled water and absolute ethanol until no chloride and bromine ions were detected by AgNO<sub>3</sub> aqueous solution (1.0 M), and then dried at 80 °C for 12.0 h. The mesoporous silica shell was achieved by removing the template agent of CTAB at 500 °C calcination for 4.0 h in air. In order to obtain Pd@Al<sub>x</sub>-mSiO<sub>2</sub> core-shell catalysts with hydrogenation activity, the calcined catalyst was reduced at 200 °C for 2.0 h in a H<sub>2</sub> atmosphere. The synthesized catalysts were denoted as the Pd@Al<sub>x</sub>-mSiO<sub>2</sub>, where x refers to the volume of NaAlO<sub>2</sub> solution. For comparison, the conventional Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst was prepared by impregnation method at the same Pd content as the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, as described previously [32].

### 2.3 Catalyst characterization

In situ X-ray diffraction (XRD) patterns of catalysts were performed on a Panalytical Empyrean X-ray diffractometer using Cu Kα ( $\lambda = 0.15418$  nm) radiation under N<sub>2</sub> flow to investigate the changes of metal nanoparticles and the thermal stability of the catalyst with increasing temperature. The XRD patterns were recorded in a  $2\theta$  angle ranging from 10° to 80° and the step size of 0.02°. The unknown diffraction peaks were identified by using the PDF-2 2018 database. Prior to test, the sample (50 mg) was pretreated at 100 °C for 30 min under N<sub>2</sub> flow. During in situ XRD experiment, the tested sample was heated stepwise (using a 5 °C/min heating rate) from room temperature to 600 °C. The XRD

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patterns were recorded at 25, 50, 100, 200, 300, 400, 500 and 600 °C. Specific surface areas were measured on a Micromeritics ASAP 2460 instrument. The transmission electron microscopy (TEM) images were obtained on JEOL JEM-2010 high-resolution electron microscope, and the accelerating voltage was at 200 kV. <sup>27</sup>Al and <sup>29</sup>Si MAS NMR were carried out on an Bruker AVANCE III 600M.

NH<sub>3</sub> temperature-programmed desorption (NH<sub>3</sub>-TPD) was performed in Auto Chem II 2920 equipment (Micromeritics). Typically, approximately 100 mg of catalyst was pre-reduced at 200 °C for 1.0 h under flowing H<sub>2</sub>, and then the catalyst was cooled to room temperature in a He flow. NH<sub>3</sub>-TPD was performed under 10% NH<sub>3</sub>/He flow for 0.5 h, and then the temperature was increased from 50 °C to 800 °C in a heating rate of 10 °C min<sup>-1</sup>. In situ FTIR spectra of pyridine, methyl octanoate and octanal were performed on a Tensor 27 instrument with modified transmission FTIR cell. Detailed procedures of the in situ FTIR were described in [Supporting Information](#).

#### 2.4 Reactivity measurements and product analysis

The hydrodeoxygenation experiments of reactants were performed in a 50 mL stainless steel autoclave. Typically, 0.1 g of reactant, 0.03 g of the activated catalyst, and 10 mL of cyclohexane solvent were added into the reactor. Prior to reaction, the reactor was purged with H<sub>2</sub> (2.0 MPa) for three times to remove the inside air, and then was pressurized into 3.0 MPa H<sub>2</sub> at room temperature. Subsequently, the mixture was heated to 260 °C with stirring speed of 1000 rpm. After reaction, the liquid products were collected and analyzed by gas chromatography (GC, Agilent 7890A/5975C). Quantitative analysis of liquid product was performed by using Eicosane as internal standard. The detailed analysis and measured procedures were described in our prior work [\[32\]](#).

To ensure reproducibility of experimental data, every experiment was performed at least two times. The conversion and product selectivity are calculated as follows (1-2):

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$$\text{Conversion} = \frac{\text{mol}_{\text{initial reactant}} - \text{mol}_{\text{unreacted reactant}}}{\text{mol}_{\text{initial reactant}}} \times 100\% \quad (1)$$

$$\text{Selec. (detec. products)} = \frac{\text{mol}_{\text{product}} \times n_{\text{C atoms in product}}}{\text{mol}_{\text{total C atoms of in detected liquid products}}} \times 100\% \quad (2)$$

Where n is the number of C atoms. Taking methyl palmitate model compound as an example, the ratio of HDO/DCO (hydrodeoxygenation/(decarboxylation + decarbonylation)) is calculated as follows

(3) [36]:

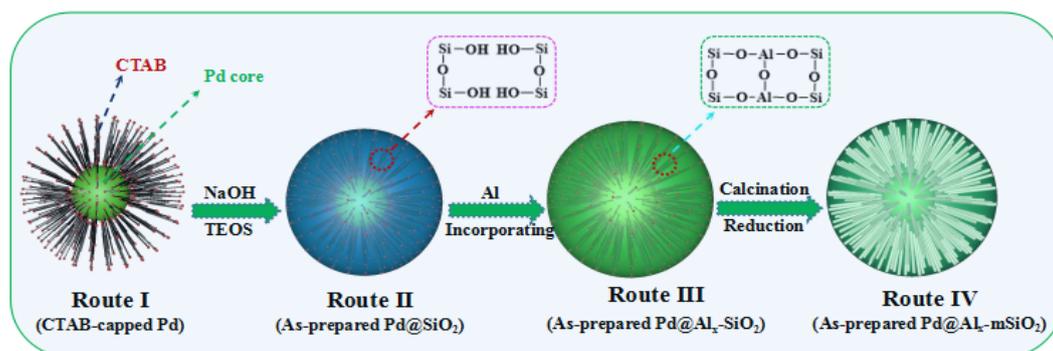
$$\text{HDO/DCO} = C_{16} / C_{15} \times 100\% \quad (3)$$

Where C<sub>15</sub> and C<sub>16</sub> are relative carbon moles of pentadecane and hexadecane, respectively.

### 3. Results and discussion

#### 3.1 Investigations of catalyst structure and physicochemical properties

In this work, the core-shell structured Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts were prepared via four steps (Scheme 1): (1) the synthesis of colloidal Pd nanoparticles using CTAB as the capping agent, (2) silica polymerization around the colloidal Pd cores under the suitable pH value, and the generation of as-synthesized Pd@SiO<sub>2</sub> catalyst, (3) incorporating Al atoms into the silica framework to produce Pd@Al<sub>x</sub>-SiO<sub>2</sub> catalysts, (4) the removal of CTAB template agent and the reduction of the as-synthesized catalyst to obtain Pd@Al<sub>x</sub>-mSiO<sub>2</sub> core-shell catalyst with mesopores and hydrogenation activity.



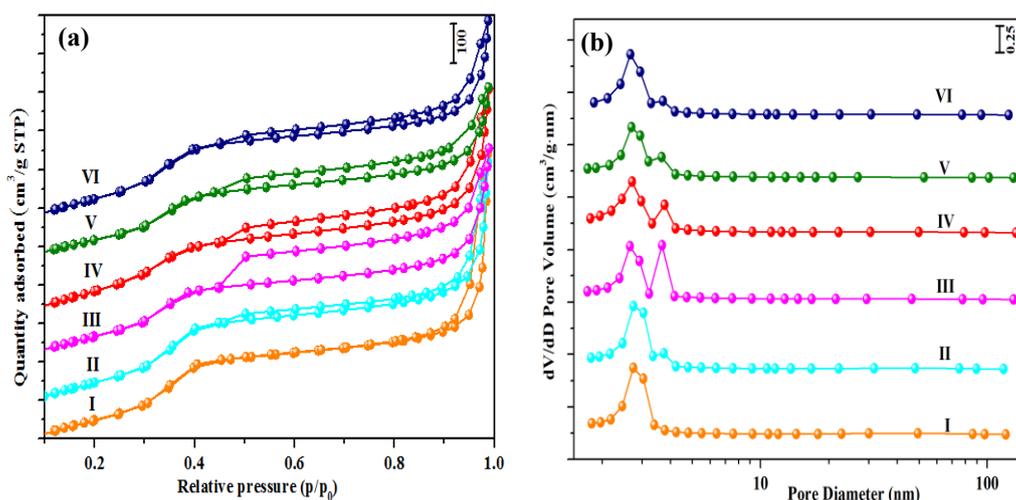
**Scheme 1.** Schematic diagram of the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts prepared by one-pot method.



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HRTEM images of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst; (e) and (f) STEM and corresponding STEM elemental mapping of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, respectively; (g) HAADF-STEM line scanning of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst.

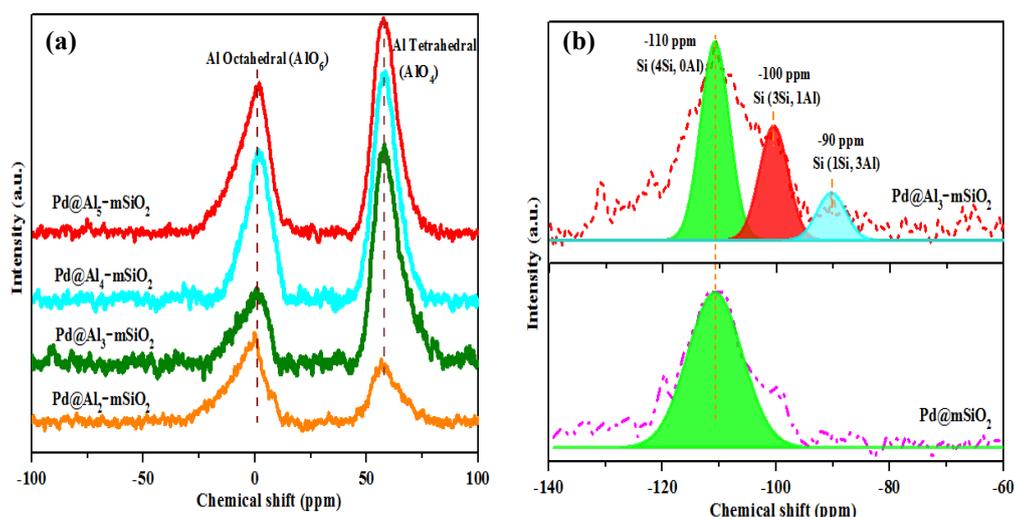
The different elemental distribution of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst was investigated by HAADF-STEM-EDS. STEM and corresponding elemental mapping (Figure 1(f)) showed that the isolated Pd metal core was closely surrounded by silica shell with aluminum doping. This phase-separated in the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst can also be indicated by EDS line scan (Figure 1(g)). Additionally, by studying the effect of aluminum content on the core-shell structure of the catalysts, it was found that, with an increase of aluminum content, the core-shell structure did not occur obvious changes as compared to Pd@mSiO<sub>2</sub> catalyst (See Figure S1).



**Figure 2.** (a) N<sub>2</sub> adsorption and desorption isotherm; (b) BJH pore size distribution curves: (I) Pd@mSiO<sub>2</sub>; (II) Pd@Al<sub>1</sub>-mSiO<sub>2</sub>; (III) Pd@Al<sub>2</sub>-mSiO<sub>2</sub>; (IV) Pd@Al<sub>3</sub>-mSiO<sub>2</sub>; (V) Pd@Al<sub>4</sub>-mSiO<sub>2</sub>; (VI) Pd@Al<sub>5</sub>-mSiO<sub>2</sub>.

Nitrogen adsorption/desorption was conducted to determine the surface areas and pore structures of the synthesized catalysts. As shown in Figure 2(a), the Pd@mSiO<sub>2</sub> and Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts all presented the typical type-IV isotherm with a characteristic type-H1 hysteresis loops at P/P<sub>0</sub> = 0.6-0.1 [38], suggesting that the outer silica shell was mesoporous structure, which is consistent with the observation in the HRTEM images. The pore sizes of all tested samples were mainly distributed in the range of 2-6 nm (Figure 2 (b)), which also confirms that the catalysts are mesoporous structure. The

parameters of the surface area, pore volume and pore diameters are listed in Table S1. As the aluminum species addition increased from 1.0 mL to 3.0 mL, the surface areas and pores of the catalysts did not occur obviously changes in comparison with the Pd@mSiO<sub>2</sub> catalyst. However, as further increased from 3.0 mL to 5.0 mL, the surface area and average pore size of the Pd@mAl<sub>x</sub>-SiO<sub>2</sub> catalysts (Pd@Al<sub>4</sub>-mSiO<sub>2</sub> and Pd@Al<sub>5</sub>-mSiO<sub>2</sub>) gradually decreased, indicating that the optimum addition amount of the aluminum species was 3.0 mL. Combined the <sup>27</sup>Al NMR spectroscopy results (Figure 3(a)), the decreased surface areas and average pore sizes in the Pd@Al<sub>4</sub>-mSiO<sub>2</sub> and Pd@Al<sub>5</sub>-mSiO<sub>2</sub> may be explained by the fact that in addition to a part of aluminum species incorporated into the silica framework, another part may be deposited on the surface and pores of the silica shell, and thus results in the decrease of surface areas and average pore sizes.

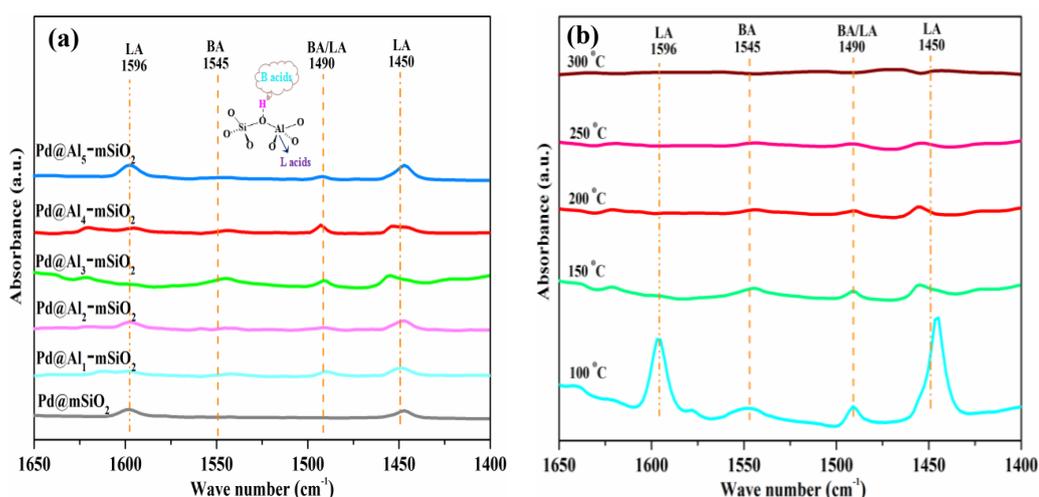


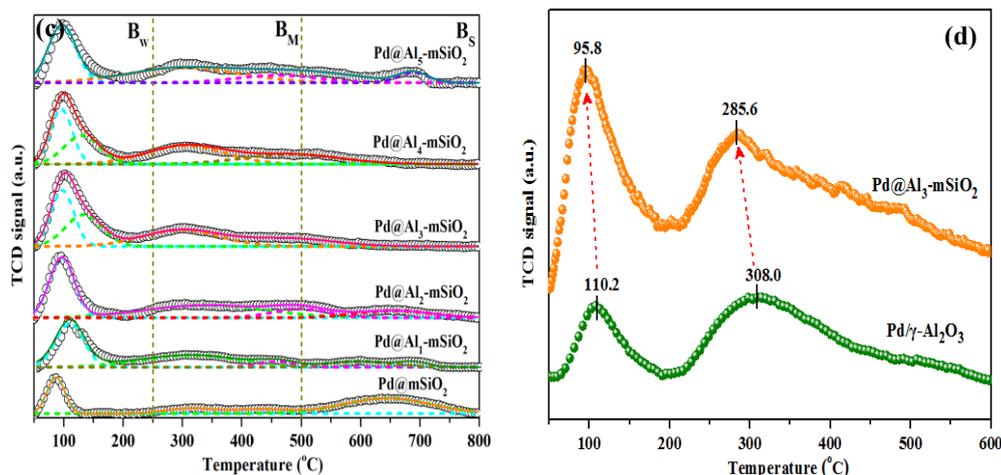
**Figure 3.** (a) <sup>27</sup>Al NMR (b) <sup>29</sup>Si NMR spectroscopy of the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts.

EDS mapping images (Figure 1(f)) only reflect the distribution of different elements on the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts. It can not illustrate the coordinated state of different species and whether the Al atoms were inserted into silica framework. In order to analyze the local environments of Al species, <sup>27</sup>Al NMR spectroscopy was performed and the results are shown in Figure 3(a). In the <sup>27</sup>Al NMR spectroscopy, the signals centered at 0 and 53 ppm were ascribed to octahedrally coordinated (AlO<sub>6</sub>

structure unit) and tetracoordinated aluminum (AlO<sub>4</sub> structure unit), respectively [39]. The 6-coordinated Al unit represented the extra-framework Al species, whereas the 4-coordinated Al unit represented the framework Al species. As shown in Figure 3(a), with aluminum species increased from 2.0 to 3.0 mL, the signal of AlO<sub>4</sub> structure unit gradually enhanced. As the aluminum content further increased from 3.0 to 5.0 mL, however, the signal of AlO<sub>6</sub> structure unit gradually enhanced, which indicates that a part of aluminum species was deposited on the extra surface of silica shell and formed AlO<sub>6</sub> structure unit, namely extra-framework aluminum.

To further confirm that aluminum atoms were inserted into the silica framework of the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts, <sup>29</sup>Si NMR characterization was performed (Figure 3(b)). Compared with the Pd@mSiO<sub>2</sub> catalyst with only one peak appeared at -110 ppm (corresponding to Si (4Si, 0Al)), two new peaks located at -100 and -90 ppm were observed in the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> spectra, which corresponded to Si (3Si, 1Al) and Si (1Si, 3Al), respectively [40]. The above <sup>27</sup>Al and <sup>29</sup>Si NMR results suggest that the aluminum atoms are successfully incorporated into the silica framework of the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts and form the Si-OH-Al bonds.





**Figure 4.** (a) FTIR spectra of pyridine on various catalysts at 150 °C; (b) in situ FTIR spectra of pyridine on the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst at different temperatures; (c) NH<sub>3</sub>-TPD profiles of different catalysts; (d) H<sub>2</sub>-TPD profiles of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> and Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalysts.

According to prior literature [24, 33], it is known that AlO<sub>4</sub> structure unit is helpful to increase the number of weak and medium acid sites, and thus influences the distribution of acid types because it forms Si-OH-Al bonds (See insert figure in Figure 4(a)). In order to analyze the surface acidity of the Pd@mSiO<sub>2</sub> and Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts, in-situ Py-FTIR and NH<sub>3</sub>-TPD measures were performed. In the Py-FTIR spectra (Figure 4(a)), the characteristic peaks appearing at approximately 1450 cm<sup>-1</sup> and 1540 cm<sup>-1</sup> were assigned to Lewis and Brønsted acid sites, respectively [41]. The adsorption peak appearing at approximately 1489 cm<sup>-1</sup> was ascribed to both Lewis and Brønsted acid sites [41]. As shown in Figure 4(a), the Pd@mSiO<sub>2</sub> catalyst only showed Lewis acid sites, while the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts displayed both Lewis and Brønsted acid sites. With aluminum content increased from 1.0 to 5.0 mL, the amounts of the total acid sites gradually increased (Table 1). These results indicate that introducing aluminum species into the Pd@mSiO<sub>2</sub> catalyst enhanced the catalyst acidity, and generated Brønsted acid sites.

To investigate the acid strength, in-situ Py-FTIR of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst was recorded. As shown in Figure 4(b), after evacuation at 100 °C, the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst showed strong peak intensity at 1450 cm<sup>-1</sup> (Lewis acid sites) and 1540 cm<sup>-1</sup> (Brønsted acid sites). With the evacuation

temperature increased, the intensity of these peaks gradually decreased. Notably, as the temperature increased to 300 °C, these peaks all disappeared, which indicates that the weak and medium acid sites were predominant in the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst. Additionally, NH<sub>3</sub>-TPD was conducted to further determine the acid strength of the Pd@mSiO<sub>2</sub> and Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts. According to the NH<sub>3</sub> desorption temperature, the acid sites were divided into weak (100-250 °C), medium (250-500 °C) and strong acid sites (>500 °C) [42]. As shown in Figure 4(c), the Pd@mSiO<sub>2</sub> catalyst only showed a weak peak centered at 100 °C, while the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> (x = 1-4) catalysts showed two NH<sub>3</sub> desorption temperature centered at 100 °C and 265 °C, ascribing to the weak and medium acid sites, respectively. As for the Pd@Al<sub>5</sub>-mSiO<sub>2</sub> catalyst, a novel NH<sub>3</sub> desorption peak was observed in the temperature range 600-700 °C, which was attributed to the strong acid sites. According to the results listed in Table 1, it is known that the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst had more weak and medium acid sites as compared with other Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts.

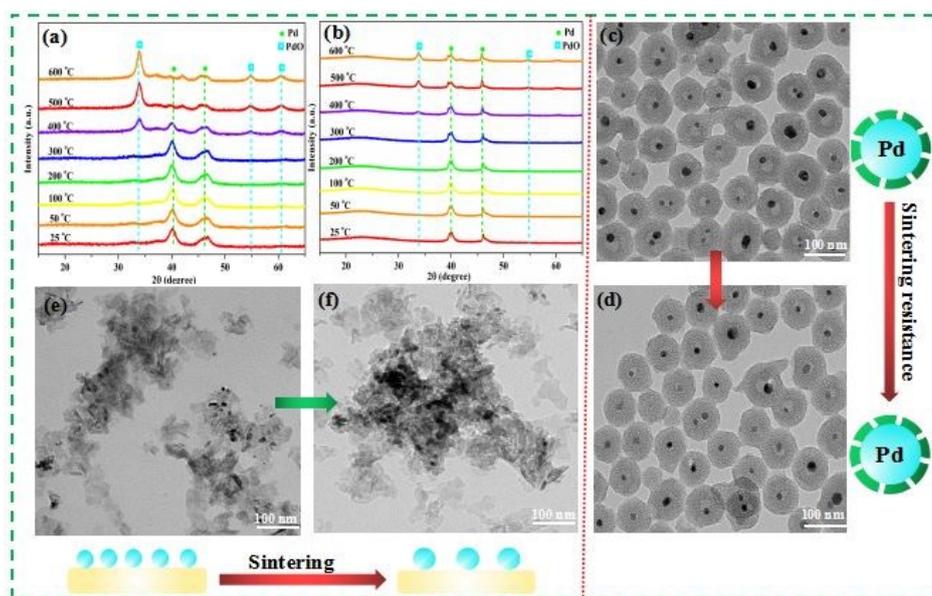
**Table 1.** Acid distributions of different catalysts

Entry	Catalyst	BA <sup>a</sup> (mmol·g <sup>-1</sup> )	LA <sup>a</sup> (mmol·g <sup>-1</sup> )	Acid Sites (mmol·g <sup>-1</sup> )		
				B <sub>W</sub> <sup>b</sup>	B <sub>M</sub> <sup>b</sup>	B <sub>T</sub> <sup>b</sup>
1	Pd@mSiO <sub>2</sub>	0.001	0.032	0.211	0.091	0.308
2	Pd@Al <sub>1</sub> -mSiO <sub>2</sub>	0.012	0.053	0.654	0.307	0.965
3	Pd@Al <sub>2</sub> -mSiO <sub>2</sub>	0.020	0.067	0.632	0.574	1.206
4	Pd@Al <sub>3</sub> -mSiO <sub>2</sub>	0.053	0.085	0.850	0.623	1.473
5	Pd@Al <sub>4</sub> -mSiO <sub>2</sub>	0.032	0.102	0.769	0.678	1.450
6	Pd@Al <sub>5</sub> -mSiO <sub>2</sub>	0.025	0.128	0.589	0.827	1.456

<sup>a</sup> BA and LA are the concentration of Brønsted and Lewis acid sites determined by Py-FTIR at 150 °C, respectively. <sup>b</sup> B<sub>W</sub>, B<sub>M</sub>, and B<sub>T</sub> are the concentration of weak, medium, and total acid sites determined by NH<sub>3</sub>-TPD, respectively.

For the hydrogenation reactions, the ability to activate and dissociate hydrogen is important for the hydrogenation catalysts. To study the ability of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, H<sub>2</sub>-TPD experiment was carried out. For comparison, the conventional Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst was also tested. As shown in Figure 4(d), compared with the Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst, the H<sub>2</sub> desorption peaks of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst all

shifted to lower temperatures and both peak area and peak intensity increased. This result indicates that Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst is easier to activate and dissociate H<sub>2</sub> than the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst, and therefore contribute to the improvement of catalytic activity.



**Figure 5.** In situ XRD patterns of Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> (a) and Pd@Al<sub>3</sub>-mSiO<sub>2</sub> (b) catalysts; TEM images of (c) Pd@Al<sub>3</sub>-mSiO<sub>2</sub> (25 °C); (d) Pd@Al<sub>3</sub>-mSiO<sub>2</sub> (600 °C); (e) Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> (25 °C); (f) Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> (600 °C).

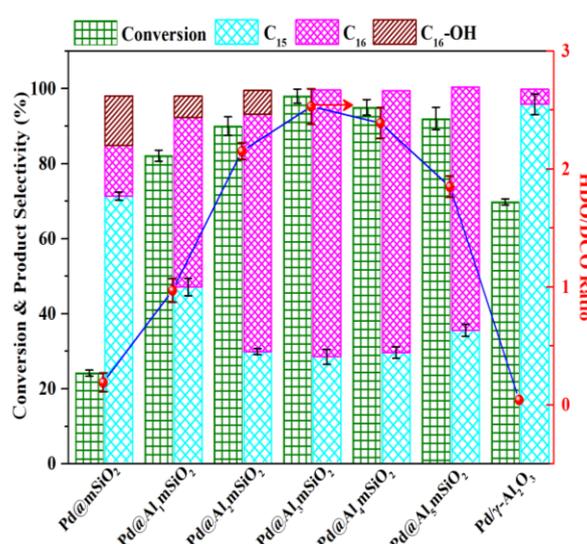
The thermal stability of the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> and Pd@Al<sub>3</sub>-mSiO<sub>2</sub> was analysed by using in-situ XRD in the temperature range 25-600 °C under N<sub>2</sub> atmosphere. Although these catalysts were tested under N<sub>2</sub> atmosphere, it is inevitable that the metal nanoparticles are oxidized during the heating process, especially in the high temperature environment. As shown in Figure 5(a), at room temperature (25 °C), the characteristic diffraction peaks of metallic Pd (JCPDS 46-1043) were observed in the XRD pattern of the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst. As test temperature increased from 25 to 300 °C, the peak intensity corresponding to Pd metal at  $2\theta = 50.2$  did not occur obviously changes. But, when the temperature increased from 300 to 400 °C, the diffraction peaks of the PdO (JCPDS 41-1107) were detected and the peak intensity gradually became strong as temperature increased. The Pd metal was completely oxidized into PdO at 600 °C. This result indicates that the dispersed Pd metal nanoparticles on the surface of Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst were easily oxidized and sintered at high temperatures, which affects

catalytic activity. In contrast, the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> core-shell catalyst preserved its XRD peak intensity, and only minor Pd metal was oxidized into PdO nanoparticles even temperature up to 600 °C (Figure 5(b)). This result indicates that the synthesized Pd@Al<sub>3</sub>-mSiO<sub>2</sub> core-shell catalyst not only has high thermal stability, but also has good oxidation resistance at high temperatures. In order to observe the changes of Pd metal at different temperatures, TEM images were made (Figure 5(c-f)). As the treated temperature increased from 25 to 600 °C, the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst maintained its core-shell structure and the metallic palladium particles were rarely aggregated (Figure 5(c) and (d)), whereas the Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst appeared obviously agglomeration (Figure 5(e) and (f)). This result demonstrates that the unique structure of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> make it have good resistance to sintering and agglomeration at high operating temperatures.

### 3.2 Catalyst activity

In order to select the optimum HDO catalyst, catalytic performances of the different catalysts were evaluated in the conversion of methyl palmitate. As shown in Figure 6, under the identical reaction conditions (260 °C and 3.0 MPa H<sub>2</sub> pressure), Pd@mSiO<sub>2</sub> catalyst showed a low catalytic activity (35% conversion), while the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts with different aluminum contents showed excellent catalytic performance (>80% conversion), where the main products were C<sub>15</sub>, C<sub>16</sub> alkanes and minor amount of cetyl alcohol (C<sub>16</sub>-OH). This indicates that the introduction of aluminum species greatly improved the catalytic activity. With aluminum content increased, the conversion of methyl palmitate first increased from 82% of Pd@Al<sub>1</sub>-mSiO<sub>2</sub> to the maximum value of 98% for the best Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, and then gradually declined to 92% for the Pd@Al<sub>5</sub>-mSiO<sub>2</sub>. According to the characteristic results of BET (Table S1), the reason for the decreased catalytic performance of the Pd@Al<sub>4</sub>-mSiO<sub>2</sub> and Pd@Al<sub>5</sub>-mSiO<sub>2</sub> catalysts was mainly attributed to the reduction of surface areas and average pore

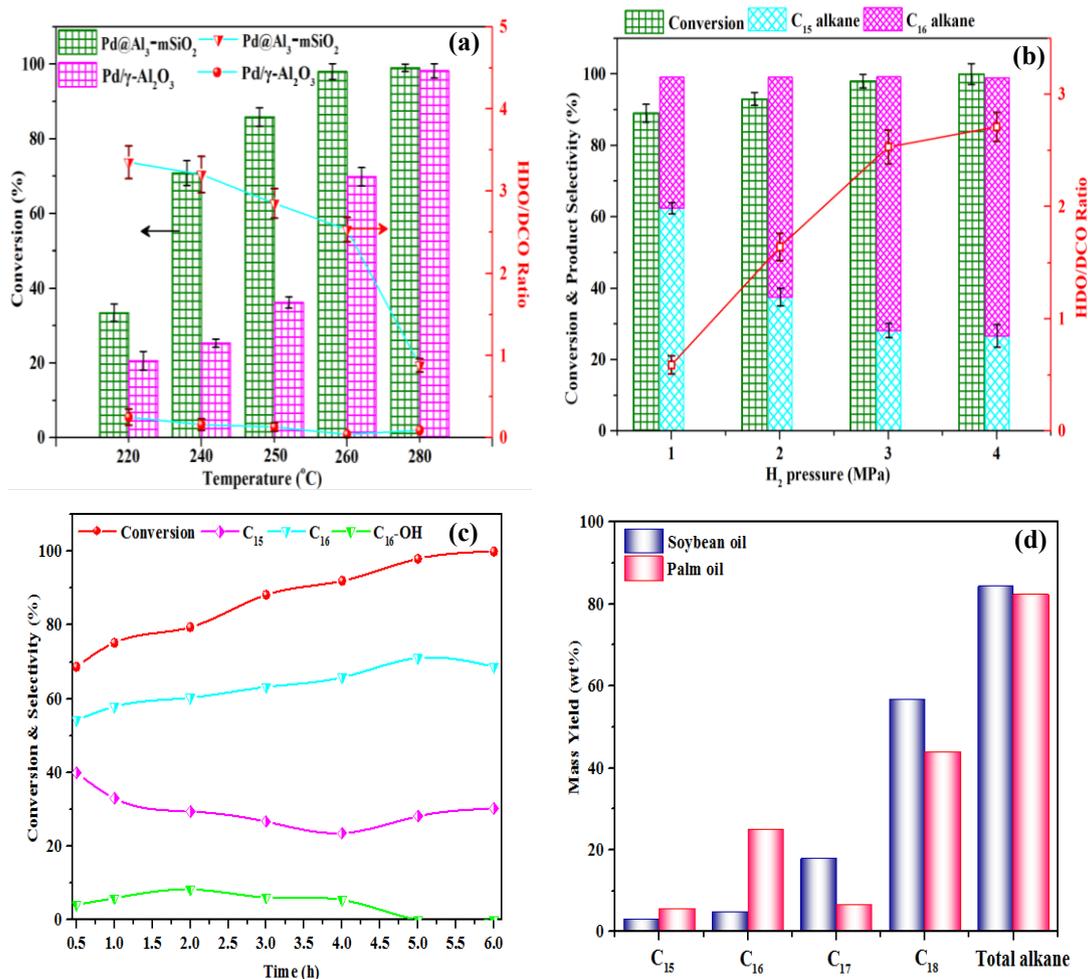
sizes, which could decrease the mass transfer efficiency of the reactant. Table S2 shows the performance of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst and comparison with the previous published studies. The Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst showed excellent catalytic performance for production of alkanes from the conversion of fatty acid esters under the mild reaction temperature of 260 °C. Moreover, compared with the conventional supported Pd metal catalysts (Pd/BaSO<sub>4</sub> and Pd/MWCNTs listed in Table S2) that exhibits a low HDO/DCO ratio (< 0.1), the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst showed a high HDO/DCO ratio (2.5). These results demonstrate that the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst has excellent catalytic performance and HDO selectivity in the conversion of fatty acid esters.



**Figure 6.** Conversion of methyl palmitate and product distributions over various catalysts. Reaction conditions: 0.03 g of catalyst; 0.1 g of reactant in 10 mL solvent at 260 °C and 3.0 MPa H<sub>2</sub> for 5.0 h.

By calculating HDO/DCO ratio, it was found that the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalysts showed higher ratio value (HDO/DCO > 0.9) in comparison with the Pd@mSiO<sub>2</sub> catalyst (HDO/DCO = 0.2). With the aluminum content increased, the ratio value exhibited first increasing and then decreasing trend. The maximum value (HDO/DCO = 2.5) was obtained over the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst. For comparison, under the identical conditions, the catalytic performance of Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst prepared by impregnated method was also studied, and the results showed that it exhibited a medium catalytic performance with 70% conversion rate. Moreover, a low HDO/DCO ratio value (HDO/DCO < 0.1)

suggested that the decarbonylation was dominant over the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst.



**Figure 7.** (a) comparison of catalytic performance between the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> and Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalysts as a function of reaction temperature; (b) effect of H<sub>2</sub> pressure on the conversion of methyl palmitate; (c) reaction composition-time profile for the conversion of methyl palmitate; (d) product distributions obtained from the conversion of soybean oil and palm oil. General conditions: 0.03 g of catalyst; 0.1 g of reactant in 10 mL solvent at 260 °C and 3.0 MPa H<sub>2</sub> for 5.0 h.

Reaction temperature would have a great effect on the catalytic performance and HDO/DCO ratio value. Therefore, with the optimum catalyst of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, the catalytic activity and HDO/DCO ratio value were studied in the temperature range 220-280 °C (Figure 7(a)). The conversion of methyl palmitate increased rapidly from 220 °C (33%) to 260 °C (98%), and remained unchanged in the range 260-280 °C (100%). In the selected temperature ranges (220-280 °C), the main products were alkane products (C<sub>15</sub> and C<sub>16</sub> alkanes). The HDO/DCO ratio value displayed a gradual decreasing trend as reaction temperature increased, revealing that high temperatures are unfavorable for the

hydrodeoxygenation selectivity of products. The ratio value was always greater than one in the selected temperatures, suggesting that the hydrodeoxygenation reaction pathway always dominates over the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst. For comparison, the catalytic performance of Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst was also evaluated with respect to reaction temperature. As shown in Figure 7(a), the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst showed medium catalytic activity (70% conversion rate) at 260 °C, and the methyl palmitate was completely converted at  $\geq$  280 °C. Besides, in the selected temperatures, the low HDO/DCO values (< 0.5) indicate that the decarbonylation reaction was always dominant on the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst. These results suggest that the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst has higher catalytic activity than the conventional Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst at a low temperature of 260 °C (98% VS 70% conversion), but also it mainly removes the oxygen atoms of fatty esters through HDO reaction pathway.

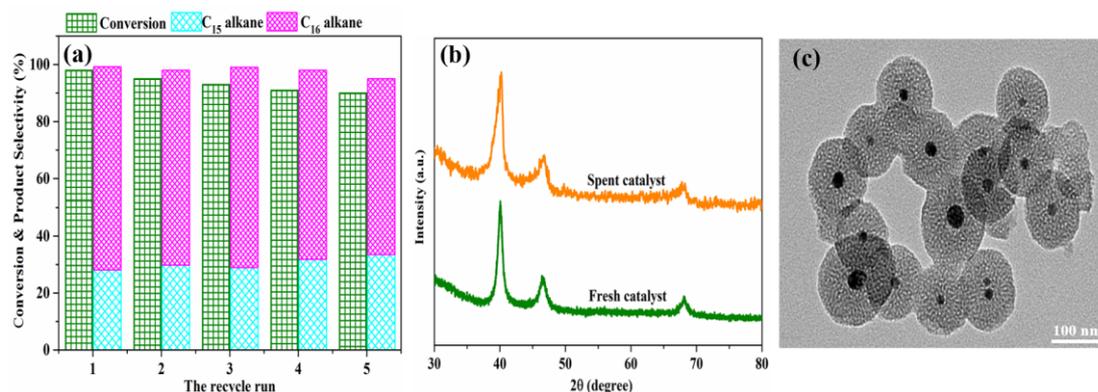
In order to obtain the optimized reaction parameters for oils conversion, other reaction conditions, such as H<sub>2</sub> pressure and reaction times, were also studied. As shown in Figure 7(b), the effect of initial hydrogen pressure on the conversion of methyl palmitate and the alkanes selectivity was not significant as that of reaction temperature. But, initial H<sub>2</sub> pressure has a great influence on the HDO/DCO value. In contrast to the decrease of HDO/DCO value caused by high reaction temperatures, the HDO/DCO value exhibited increasing trend with initial hydrogen pressure increasing from 1.0 to 3.0 MPa. Further increasing H<sub>2</sub> pressure to 4.0 MPa did not result in obvious changes in product distribution, suggesting that the optimized H<sub>2</sub> pressure was 3.0 MPa. Figure 7(c) showed that the optimal reaction time was 5.0 h, in which the reactant was basically converted and the alkanes selectivity was as high as 98%. Based on the above results and the principle of minimum energy consumption, the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, 260 °C of reaction temperature, 3.0 MPa of H<sub>2</sub> pressure and 5.0 h of reaction time were chosen.

Under the optimum catalyst and reaction conditions, the vegetable oils such as soybean oil and

palm oil were hydrotreated to produce diesel-range alkanes (Figure 7(d)). The composition of these vegetable oils was listed in Table S3. Under the selected reaction conditions, a high yield of diesel-range alkanes (84.3 wt% and 82.5 wt%, respectively) from the conversion of the two oils were obtained over the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, and these values are close to the theoretical values (soybean oil : 86.4 wt%; palm oil : 85.2 wt%) [15]. Moreover, it was found that the side reaction of the C-C bond cleavage was largely suppressed and only a small amount of cracking products were obtained (< 3 wt%) due to the low operating temperature of 260 °C. These results indicate that the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst is a promising catalyst which realizes the low-temperature deoxygenation of vegetable oils to produce the diesel-range alkanes.

Figure 8(a) shows the reusability of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst in the conversion of methyl palmitate. After the first run, the catalyst was obtained by centrifugation, dried at 80 °C, and then directly used the next run. As shown in Figure 8(a), after recycling five runs, the catalytic activity of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> showed a slight loss (conversion decreased from 98% to 90%) while the selectivity of alkanes remained above 95%. These results suggest that the synthesized Pd@Al<sub>3</sub>-mSiO<sub>2</sub> core-shell-type catalyst has good reusability. The XRD and TEM characteristic measures (Figure 8(b) and (c)) were conducted to explore the structure of the spent catalyst. After the fifth recycling test, the XRD patterns of the spent Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst revealed that the diffraction peaks assigning to Pd metal were almost no obvious changes in comparison with the fresh catalyst. TEM images showed that the spent Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst preserved its core-shell structure and no obvious metal agglomeration was observed. Based on the in-situ XRD and TEM images, we concluded that the good stability of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst could be attributed to its unique core-shell structure, where the silica shell prevents the leaching of Pd element and inhibits the agglomeration of metal nanoparticles at

high temperatures.

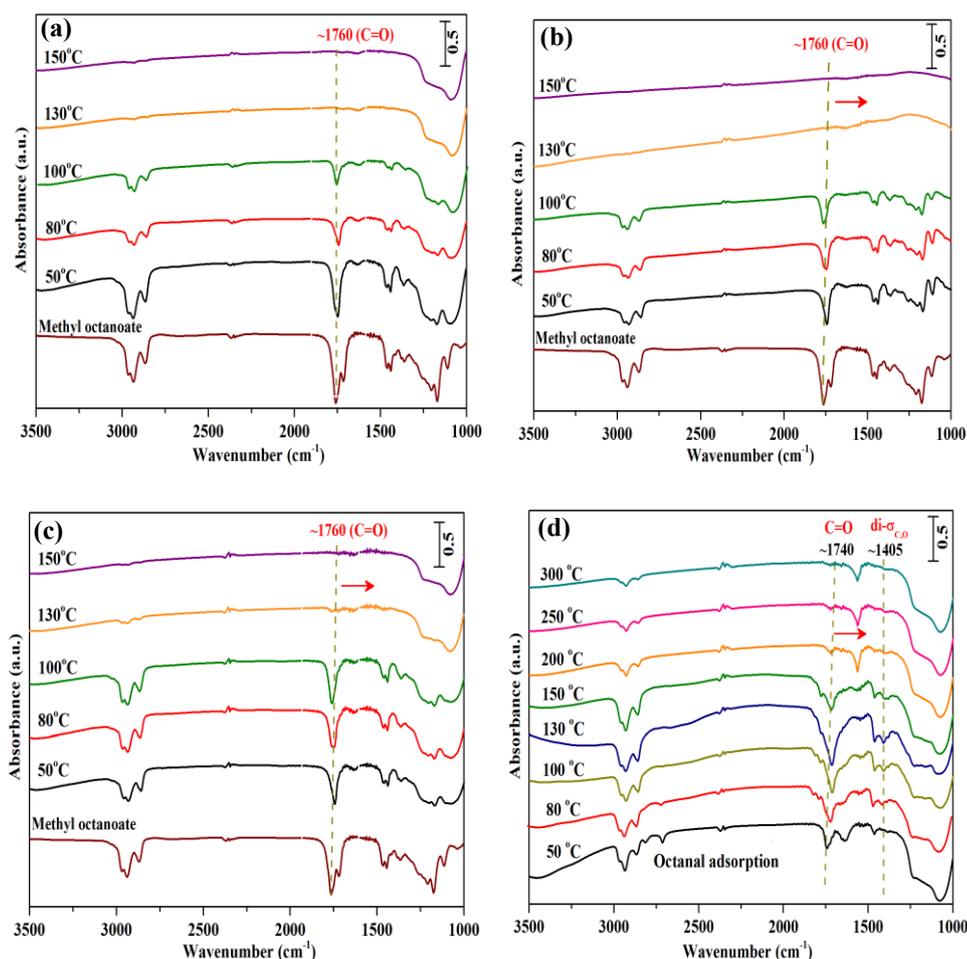


**Figure 8.** (a) Recyclability of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst; (b) XRD patterns of the spent and fresh catalysts; (c) TEM image of the spent catalyst. Reaction conditions: reactant (0.1 g), catalyst (0.03 g), cyclohexane (10 mL), 260 °C, initial H<sub>2</sub> pressure (3.0 MPa).

### 3.4 Analysis of the structure-activity relationship

According to the above results, a question we can raise is why the synthesized Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst has higher catalytic activity and HDO/DCO value than the Pd@mSiO<sub>2</sub> and Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalysts at a low temperature of 260 °C. Apparently, there are two main factors that affect these values, i.e., the strong synergistic effect of the metal-oxide and the type of acid sites. At the Pd-Al interface, H<sub>2</sub> was dissociated into active hydrogen atoms over the metallic Pd sites, while the carbonyl group of fatty esters was adsorbed and activated over the Al<sup>3+</sup> sites. The active hydrogen atoms transferred to Al sites via spillover effects and hydrogenated the adsorbed fatty esters to form fatty aldehyde intermediate. Subsequent, the resulting fatty aldehyde intermediate was converted into alkanes via two parallel reaction pathways: either via decarbonylation to produce RH alkanes, or via hydrogenation to produce fatty alcohols, and then produce RCH<sub>3</sub> alkanes through dehydrated-hydrogenation process under the action of Brønsted acids and Pd metal. In contrast to the conventional Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst, uniformly dispersed metallic Pd nanoparticles (providing more active hydrogen atoms as proved in H<sub>2</sub>-TPD (Figure 4(d)), strong interaction between the reactant and the interface of Pd-Al, and abundant the number of Brønsted acid sites of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> make it have higher catalytic activity and

HDO/DCO value.

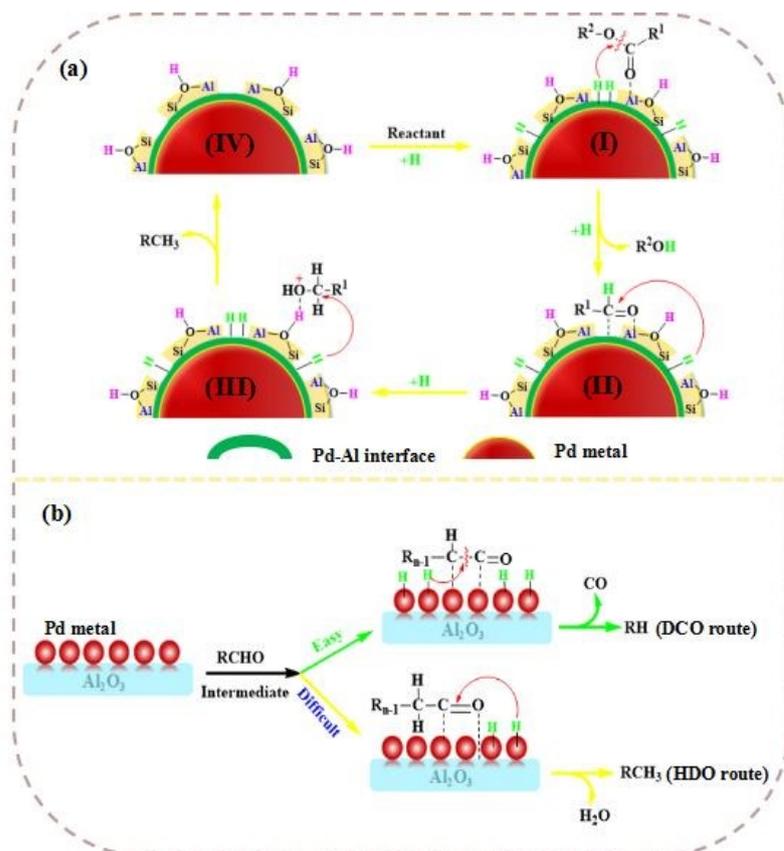


**Figure 9.** In situ FTIR for the methyl octanoate adsorption over the Pd@mSiO<sub>2</sub> (a); Pd/γ-Al<sub>2</sub>O<sub>3</sub> (b); Pd@Al<sub>3</sub>-mSiO<sub>2</sub> (c); and for the octanal adsorption over the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> (d).

To confirm the speculation, in situ FTIR of the methyl octanoate was performed. Methyl octanoate was selected as the probe molecule because of its low boiling points and similar functional groups to long-chain fatty esters. As shown in Figure 9(a), the band at 1750 cm<sup>-1</sup> was attributed to the vibration modes of  $\nu$  (C=O) [43, 44], while the bands at 2990-2885 cm<sup>-1</sup> were ascribed to the stretching frequencies of the aliphatic C-H (CH<sub>3</sub>, CH<sub>2</sub>, and CH) [17, 45]. Noted that the spectrum of methyl octanoate adsorbed on the Pd@mSiO<sub>2</sub> and Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalysts displayed weak bands at 100 °C, which cannot be detected when the temperature increased to 130 °C (Figure 9(a) and (b), respectively). In contrast, the characteristic band (C=O) of methyl octanoate was still observed on the Pd@Al<sub>3</sub>-mSiO<sub>2</sub>

catalyst even at 130 °C (Figure 9(c)). These suggest that methyl octanoate has a stronger interaction with the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst than that of the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> and Pd@mSiO<sub>2</sub> catalysts. Additionally, chemical adsorption of methyl octanoate was observed on the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> and Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalysts, where the characteristic band shifted to higher wavenumbers. In contrast, this chemical shift was not observed on the Pd@mSiO<sub>2</sub> catalyst. These results revealed that the Al<sup>3+</sup> active sites in the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> and Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalysts play an important role in adsorbing and activating carbonyl group of fatty esters, and the strongest interaction occurred on the surface of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst.

To investigate the adsorption mode of fatty aldehyde intermediate on the surface of Pd@Al<sub>3</sub>-mSiO<sub>2</sub>, in situ FTIR of octanal was performed (Figure 9(d)). The C=O stretching band of the octanal was observed at 1720 cm<sup>-1</sup>, and the characteristic band shifted to lower wavenumbers with temperature increased. Besides, the appearance of the band at 1405 cm<sup>-1</sup>, an indicator of bicoordination adsorption of carbonyl carbon and oxygen atoms in fatty aldehyde (denoted as di- $\sigma_{C,O}$  configuration) [46], indicates that the octanal was chemically adsorbed on the surface of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst in a mode of di- $\sigma_{C,O}$  configuration via C=O bonds. These results suggest that there is an interaction between the fatty aldehyde intermediate and the Pd-Al interface, i.e., the carbonyl carbon and oxygen atoms in fatty aldehyde were bonded on the Pd and Al active sites, respectively.



**Scheme 2.** (a) Surface reaction model for the hydrodeoxygenation of fatty esters over the Pd@Al<sub>x</sub>-mSiO<sub>2</sub> catalyst; (b) Surface reaction model for the deoxygenation of fatty aldehyde intermediates over the Pd/γ-Al<sub>2</sub>O<sub>3</sub> catalyst.

To study why the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst exhibited high HDO/DCO value, we normalized the HDO/DCO value as a function of B acid sites (Figure S2). As the number of B acid sites increases, the HDO/DCO value exhibits an increasing trend. This indicates that B acid sites play an important role in improving the HDO/DCO value. From Figure S2, it was found that the HDO/DCO value was not linear relationship with the increase of B acid sites, which reveals that B acid sites are not an exclusive factor for improving the HDO/DCO value. Considering the unique structure of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst, we believe that the high HDO/DCO value may be associated with its unique core-shell structure.

The core-shell structure that composed of the inner Pd metal core and the outer silica shell containing Al active sites enables the fatty aldehyde intermediate first to be adsorbed on the Al sites, and then contacts with inner Pd metal core. This leads fatty aldehyde intermediate to adsorb on the

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Pd-Al interface in a mode of di- $\sigma_{C,O}$  configuration to produce fatty alcohols (Scheme 2(a) (II)) (as proved in Figure 9(d)), followed by generating alkanes via the dehydrated-hydrogenation process under the action of Brønsted acid sites and Pd metal (Scheme 2(a) (III)). In contrast, on the surface of the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst (Scheme 2(b)), since the metallic Pd is exposed on the surface of the Al<sub>2</sub>O<sub>3</sub> support, the fatty aldehyde is easily adsorbed on the Pd metal surface through C-C bonds, which results in the cleavage of C-CHO bonds to produce alkane with one carbon atom less (Scheme 2(b) (DCO route)). Additionally, since the Pd/ $\gamma$ -Al<sub>2</sub>O<sub>3</sub> catalyst contains minor B acid sites (0.001 mmol·g<sup>-1</sup> determined by Py-FTIR at 150 °C), the generating fatty alcohol is more easier to occur dehydrogenation than the dehydration reaction on the exposed metallic Pd sites, which produced fatty aldehydes. Subsequently, the generated fatty aldehyde produced alkanes by decarbonylation reaction, which results in a low HDO/DCO ratio value. Based on the above results and analysis, we concluded that under mild conditions (260 °C and 3.0 MPa H<sub>2</sub> pressure), the high catalytic activity of the Pd@Al<sub>3</sub>-mSiO<sub>2</sub> catalyst was attributed to its excellent ability of H<sub>2</sub> dissociation/activation and strong interaction between the reactant and the Pd-Al interface active sites, while the high HDO/DCO value was ascribed to two aspects, i.e., abundant B acid sites on the surface of the outer silica shell and its core-shell structure containing inner metallic Pd core and outer silica shell doped with aluminum atoms.

#### 4. Conclusions

In this work, to solve the problems that traditional supported-palladium catalysts are prone to cause the cleavage of C-C bonds and occur agglomeration during the conversion of fatty acid esters, we prepared an efficient and stable Pd@Al<sub>x</sub>-mSiO<sub>2</sub> core-shell nanocatalysts doped with aluminum atoms. Under the mild reaction conditions (260 °C and 3.0 MPa), the optimized Pd@Al<sub>3</sub>-mSiO<sub>2</sub> not only exhibited high HDO activity and selectivity towards diesel-range alkanes, but also showed good

stability with a slight loss (conversion decreased from 98% to 90%) over five cycles. The structure-activity relationship based on the detailed characterization (TEM, HRTEM, <sup>27</sup>Al NMR NH<sub>3</sub>-TPD, and in situ FTIR) indicates that the high catalytic performance of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> mainly results from the cooperative effects of uniformly dispersed Pd metal, Lewis- and Brønsted-acid sites generated through the formation of Si-OH-Al bonds, while the good stability of Pd@Al<sub>3</sub>-mSiO<sub>2</sub> was ascribed to its unique core-shell structure, inhibiting the leaching and agglomeration of metallic Pd nanoparticles during the reaction.

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