

# Quantification of liquid phase faecal odour to evaluate membrane technology for wastewater reuse from decentralised sanitation facilities

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## Abstract

Decentralised developing country sanitation facilities typically produce water of insufficient quality for local reuse despite the pressing need for such resources. A further limiting criterion is public willingness to use such facilities or the arising water product due to odour. Whilst odour is characterised in the gas phase, it originates in the liquid phase. Consequently, controlling odour at source could prevent gas-phase partitioning and limit produced water contamination. This study therefore developed an analytical method for the quantitation of a range of liquid phase volatile organic compounds (VOCs) classified into eight chemical groups, known to be primary indicators of faecal odour, to provide characterisation of real fluids and to permit evaluation of several potential membrane separation technologies for liquid phase odour separation. The gas chromatography mass spectrometry method provided quantitation in the range of 0.005 mg L<sup>-1</sup> to 100 mg L<sup>-1</sup> with instrument detection limits ranging from 0.005 mg L<sup>-1</sup> to 0.124 mg L<sup>-1</sup>. Linear calibration curves were achieved ( $r^2 > 0.99$ ) with acceptable accuracy (77-115%) and precision (<15%) for quantitation in the calibration range below 1 mg L<sup>-1</sup>, and good accuracy (98-104%) and precision (<2%) determined for calibration in the range 1-100 mg L<sup>-1</sup>. Pre-concentration of real samples was facilitated via solid phase extraction. For the selected VOC range, recoveries were classified as acceptable (50-63%, dimethyl trisulfide, 2-butanone and dimethyl disulfide) and excellent (77-100%, 1-butanol, benzaldehyde, indole, ethyl propionate, skatole, *p*-cresol, ethyl butyrate), and were consistent to within a relative standard deviation of <10%. Subsequent application of the method to the evaluation of two thermally driven membranes based on hydrophilic (polyvinyl alcohol) and hydrophobic (polydimethylsiloxane) polymers evidenced contrasting separation profiles and the nascent efficacy of polyvinyl alcohol based membranes for the separation of liquid phase odour. Importantly, this study demonstrates the methods utility for liquid phase VOC determination which is of use to a range of disciplines, including healthcare professionals, sentient specialists and public health engineers.

**Keywords:** wastewater; taste; sewage; pervaporation; membrane distillation; pit latrine

## 1. Introduction

Large scale centralised wastewater treatment is not economically practicable for implementation in many developing country contexts. Local communities are therefore instead dependent upon decentralised sanitation solutions such as pit latrines that do not provide a safe barrier to discharge of faecal material into local water resources. Malodour associated with these sanitation facilities has also been shown to exacerbate discharge of faecal material into the environment with users preferring open defecation to foul-smelling pit latrines [1,2]. The odour profile associated with decentralised sanitation can be considered distinct from that of centralised treatment facilities since the absence of flush water and other water sources limit the primary composition to urine and faeces. Whilst there are approximately 279 and 381 volatile organic compounds (VOCs) associated with urine and faeces respectively from healthy individuals [3], the faecal-borne VOCs indole, skatole (3-Methyl-1H-indole) and *p*-cresol (4-Methylphenol) amongst others, are considered key contributors to malodour arising from pit latrines [4]. Previous research has demonstrated that VOCs originate in the liquid phase as microbial metabolites, with factors such as diet and health influencing composition and concentration of VOCs, and the physico-chemical environmental conditions (e.g. pH and temperature) encouraging partitioning into the gas phase where odour is finally perceived.

Recent technological innovations seek to deliver alternative sustainable sanitation solutions that can facilitate sufficient water quality for safe discharge to the environment or to promote local water re-use [5,6]. As water supplies often arise from sources of unknown provenance, the production of water to re-use standards can be considered an attractive proposition. However, a major limiting criterion that governs willingness to use reclaimed water is odour [7]. Odour abatement technologies presently provide elimination or neutralisation of malodourous compounds already partitioned into the gas phase [2]. Through introducing barrier technology into this new genre of sanitation solutions for liquid phase treatment, the partitioning of odorous VOCs from the liquid phase into the gas phase could be mediated at source and potentially averted, therefore enhancing the potential willingness of users to use locally engineered sanitation solutions and the arising water product for a range of re-use applications [7]. Pervaporation fosters water transport through application of a vapour pressure gradient and permeation through a polymeric membrane. The availability of waste heat, coupled with characteristically low water volumes from these new decentralised sanitation solutions, make thermally driven membrane separation a practicable solution for water recovery [6]. For non-porous (or dense) membranes, the polymer chemistry can favour

permeation of water over VOCs thereby imparting selectivity into the separation that will exert an influence on the final odour profile of the treated water.

Whilst the management of odour in the liquid phase is an attractive proposition, there is presently not an analytical solution of sufficient resolution to characterise the separation performance of membrane technology for this application. The conventional analytical route that has been previously exploited for liquid phase VOC odour determination is headspace sampling with pre-concentration onto a sorbent (e.g. Tenax) before introduction into gas chromatography mass spectrometry (GC-MS) [2,8]. Such indirect techniques introduce temporal and sample volume restrictions in addition to limitations with respect to recovery which do not guarantee accurate quantitation of the liquid phase VOC profile. Lin et al. [2] recently introduced a direct method for liquid phase VOC odour characterisation of pit latrine faecal sludge using solid phase extraction (SPE) for pre-concentration from the liquid phase before determination by GC-MS. The authors used the method to successfully identify a discrete range of VOCs in the liquid phase representative of faecal odour. Pre-concentration by SPE was also selected for study by Chappuis et al. [4] to extract compounds from pit latrine air in which the equilibrium was shifted to the liquid phase to trap and concentrate the compounds, enabling quantitation close to the odour detection thresholds (ODTs) to be achieved.

Although SPE-GCMS has been demonstrated as a suitable method for liquid phase VOC quantitation, only a discrete range of VOCs has been determined, representing a limited range of chemical structures that is not sufficiently definitive to aid in the characterisation and development of membrane technology for the selective separation of liquid phase odour. This study therefore seeks to develop an analytical method for the determination of liquid phase odour sufficient to characterise a broad range of VOC chemistries including organo-sulphurs, aromatics, phenols, alcohols, aldehydes, ketones, esters and hydrocarbons, that are known contributors to faecal odour [3,8] and within a single elution to simplify the analytical procedure. Specific objectives are therefore to: (i) develop a method for the quantitation of liquid phase VOCs within a single elution, which present a broad range of chemistries, representative of those commonly associated with faeces and urine; (ii) develop and validate solid phase extraction for the liquid-phase pre-concentration stage; (iii) apply the method for VOC quantitation in urine and faecally contaminated urine; and (iv) confirm the methods validity through application to pervaporative membranes of differing polarity that should engender distinct differences in liquid phase VOC separation.

## 2. Materials and methods

### 2.1. Chemicals and reagents

All chemicals were sourced from Fisher Scientific (Loughborough, UK) or Sigma Aldrich (Dorset, UK). The VOCs analytes (1-butanol, 1-propanol, benzaldehyde, indole, skatole, ethyl butyrate, ethyl propionate, limonene, 2-butanone, *p*-cresol, dimethyl disulfide and dimethyl trisulfide) had a purity of at least 98%. Diethyl ether, propylene glycol, and the methyl octanoate internal standard were of extra pure grade ( $\geq 99\%$ ) and the methanol used for SPE conditioning and acetone used for glassware cleaning was laboratory grade.

### 2.2. Standards preparation

All standards were prepared in Class A volumetric glassware which was cleaned to remove residual contaminants by soaking glassware in deionised water, acetone and methanol for 10 minutes respectively, within a sonicator and then dried overnight at 50°C. For the calibration range between 1-100 mg L<sup>-1</sup>, a 1000 mg L<sup>-1</sup> stock solution of all VOCs was prepared in diethyl ether and subsequently diluted according to the calibration concentration. The standards were spiked with the internal standard (IS, methyl octanoate) for a final concentration of 10 mg L<sup>-1</sup>, the upper and lower limit of the two calibration curves (1–10 and 10-100 mg L<sup>-1</sup>), allowing for both ranges to be run on the same sample. A 10 mg L<sup>-1</sup> stock solution was prepared for the lower calibration curve (<1 mg L<sup>-1</sup>) and spiked with the internal standard for a final concentration of 1 mg L<sup>-1</sup>. Internal standard response curves were plotted for each compound with the mean response factor used to determine unknown concentrations.

### 2.3. Gas Chromatography Mass Spectrometry

Compound identification and quantification were performed using a Shimadzu-TQ8040 GC-MS (Shimadzu, Milton Keynes, UK), equipped with a semi polar ZB-624 fused silica GC column (thickness: 1.4  $\mu\text{m}$ , length: 60 m, diameter: 0.25 mm, Phenomenex, Macclesfield, UK). The initial GC oven temperature was held at 35°C for 5 minutes, equivalent to the boiling point of the solvent (diethyl ether) and then increased at a rate of 10°C min<sup>-1</sup> until reaching 170°C in order to obtain 1-propanol, 2-butanone, 1-butanol, ethyl propionate, dimethyl disulfide, and ethyl butyrate. This was followed by an isothermal temperature section for 2 minutes to provide separation between dimethyl trisulfide, benzaldehyde and limonene. The ramp was increased to a rate of 30°C min<sup>-1</sup> until reaching 240°C for the detection of the internal standard (methyl octanoate) and *p*-cresol and then further increased to 250°C at a rate of 5°C min<sup>-1</sup>. An isothermal temperature section was maintained for 5 minutes, allowing for the separation of

indole and skatole. The total runtime was 29.83 minutes. Helium was used as the carrier gas (236.1 kPa) at a linear column flow rate of 2.47 ml min<sup>-1</sup> to maintain a velocity of 40 cm s<sup>-1</sup>. The mass spectrometer was operated with a detector voltage relative to the tuning result (0.2 kV) at an ion source temperature of 200°C and interface temperature of 250°C. A solvent cut time was applied until 8.95 minutes. Initially, the MS was operated in scan mode in order to identify the retention times and target ions through in house MS libraries and NIST MS search with a scan range of 30-500 m/z. Compounds of interest were then detected in single ion monitoring (SIM) mode by the principal ion and two reference ions (Table 2).

#### 2.4. Determining SPE recovery factors

A synthetic solution was prepared in order to determine SPE recovery factors. A 1000 mg L<sup>-1</sup> stock solution containing all VOCs was prepared in propylene glycol to completely dissolve all compounds. An aliquot was subsequently added to three buffered solutions (potassium chloride buffer pH 2, potassium phosphate monobasic 6.5 and tris (hydroxymethyl aminomethane pH 9) according to Robinson and Stokes [9], within a volumetric flask for an injection concentration of 100 mg L<sup>-1</sup>.

Oasis® HLB cartridges (1 g), sourced from Waters (Milford, USA) were used and attached to an Agilent VacElut20 manifold (Agilent Technologies, Stockport, UK). The cartridges were first conditioned following a sequence of flushing with 10 mL diethyl ether, methanol then deionised water, facilitated by a vacuum pump (N 022 AN.18, KNF Neuberger, Whitney, UK). The 20 mL sample was loaded onto the cartridge ensuring that the cartridge did not completely dry out, allowing the VOCs from the sample to adsorb onto the cartridge with the output water discarded. The cartridge was eluted with 1 mL of methyl octanoate (IS) in diethyl ether (0.057 µg mL<sup>-1</sup>), and collected into a 10 mL glass centrifugal vial (Cole Parmer, London, UK). The cartridge was eluted with an additional 5 mL of pure diethyl ether, and collected into the same vial. The residual water which collected at the bottom of the beaker was removed carefully using a glass Pasteur pipette (Fisher Scientific, Loughborough, UK). The eluent was concentrated by evaporation using nitrogen gas at a flow which did not exceed surface breakage. These vials were analysed by GCMS. The response ratios were compared between the calibration standard and the sample in order to calculate the recovery factors of the compounds. All trials were triplicated at pH 2, 6.5 and 9. The method detection limit (MDL) was determined by:

$$MDL = \frac{IDL * 100}{C_f * R_f} \quad (\text{Equation 1})$$

where  $C_f$  is concentration factor and  $R_f$  is the recovery factor.

### 2.5. Characterisation of urine and faecally contaminated urine

Fresh urine and faeces samples were collected from consenting anonymous volunteers through a collection regime approved by the Cranfield University Research Ethics System (CURES, project ID 3022). A box containing instructions, urine pots, disposable sampling bowls, a waste and sample bag was placed within the designated donation toilet. The samples were collected and analysed within 12 hours of collection.

The faecally contaminated urine was prepared by producing a composite sample comprised of a 10:1 urine to faeces ratio, which represents the typical proportions produced by an individual per day [10]. A 5 g sample of fresh faeces was transferred into a 50 mL falcon tube and mixed with fresh urine (50 g) on a vortex mixer for 30 seconds. The sample was then filtered through cotton wool and sand (50 mL) and a 20 mL aliquot was processed by SPE. Urine samples (20 mL) were also processed using SPE. All samples were eluted with 0.2 mL IS solution ( $0.057 \mu\text{g mL}^{-1}$ ) and 10 mL diethyl ether and concentrated down to 100  $\mu\text{L}$ . An additional sample with a concentration factor of five was also processed to capture *p*-cresol concentrations exceeding the calibration range i.e (2.5 ml sample, 1 mL IS solution, 10 mL diethyl ether, concentrated down to 500  $\mu\text{L}$ ).

### 2.6 Membrane technology set-up

Commercially available polydimethylsiloxane (PDMS) and polyvinyl alcohol (PVA) membranes were evaluated (Table S7). The PDMS and PVA membranes exhibited contact angles of  $116 \pm 1.4^\circ$  and  $43 \pm 1.1^\circ$ , indicating them to be hydrophobic and hydrophilic polymers respectively. Vapour pressure gradient was established using a diaphragm vacuum pump (MD 4CNT, Vacuubrand, Brackley) operating at 0.05 bar on the permeate side. Permeate samples were collected (20 mL) within a liquid nitrogen cold trap ( $-196^\circ\text{C}$ ). The permeate, feed and retentate samples were analysed using the SPE-GCMS method to establish a mass balance. The feed reservoir was submerged within a thermostatic bath at  $50^\circ\text{C}$  (Grant TC120, Cambridge, UK) with a feed flowrate of  $0.2 \text{ L min}^{-1}$  applied (520s, Watson Marlow, Falmouth, UK). Separation efficiency of the PVA and PDMS membranes was expressed through removal efficiency (%):

$$\frac{C}{C_0} \times 100 \quad (\text{Equation 1})$$

and enrichment factor ( $\beta$ ) respectively:

$$\frac{C}{C_0} \quad (\text{Equation 2})$$

where  $C$  represents the permeate concentration ( $\text{mg L}^{-1}$ ) and  $C_0$  is the initial feed concentration ( $\text{mg L}^{-1}$ ). All trials were conducted in triplicate.

### 3. Results and discussion

#### 3.1 Method development

The VOC analytes comprised of alcohols (1-butanol, 1-propanol), aldehydes (benzaldehyde), aromatics (indole, skatole), esters (ethyl butyrate, ethyl propionate), hydrocarbons (limonene), ketones (2-butanone), phenols (*p*-cresol) and organo-sulphur containing compounds (dimethyl disulfide, dimethyl trisulfide) (Table 1). The compounds represent a broad range of physico-chemical properties such as acid dissociation constant ( $\text{pK}_a$ , -7 to 16.1), octanol-water partitioning coefficient ( $\log K_{ow}$  0.25 to 4.57), water solubility ( $0.013$  to  $1000 \text{ g L}^{-1}$ ) and volatility ( $0.00048$  to  $19.1 \text{ mol m}^{-3} \text{ Pa}$ ) [11–14], which confer a challenging separation for any barrier technology, and is representative of the chemistries frequently associated with faecal odour [2,4,8]. Various split ratios were trialed in scan mode to identify a method capable of detecting each VOC in this range within a single elution. The optimum split ratios were selected according to the upper limit of detector saturation which was associated to the later emerging higher boiling point compounds (aromatics) and a signal to noise ratio of  $>10$  for the lower boiling point compounds (alcohols). The injection port was operated at a split of 1:5, 1:12.5 and 1:100 for the low calibration range ( $0.005$ - $1 \text{ mg L}^{-1}$ ), medium calibration range ( $1$ - $10 \text{ mg L}^{-1}$ ) and high calibration range ( $10$ - $100 \text{ mg L}^{-1}$ ) respectively; three calibration ranges were adopted to ensure that the ‘natural’ concentration of faecally contaminated urine as well as sample concentrations post-separation could be determined. The respective injection volumes were  $2.5 \mu\text{L}$ ,  $1 \mu\text{L}$  and  $1 \mu\text{L}$ . The split ratio conditions were then applied to SIM mode to increase selectivity and sensitivity (Table 2). The final peak of the elution (Figure 1a and b) represents butylated hydrocarbon (BHT), the stabilisation agent within the diethyl ether solvent. All compounds were detected within a 27 minute runtime with narrow symmetrical peaks identified even at low concentrations. Peaks generally had good tailing factors close to one which was within the recommended analytical range of  $\leq 2$  (see S1-3)[15,16].

#### 3.2 GC-MS calibration

Calibration was based on a linear regression analysis of the mean response factor fit [17] (Table 3). A good correlation coefficient was obtained for each of the three calibration curves

( $r^2 > 0.99$ ). Residual standard deviations (RSD) of the response factors of all calibration curves were within the acceptance criteria of  $< 20\%$  [18]. The instrument limit of detection (LD) was calculated as  $3.3 \sigma/\text{slope}$ , and limit of quantification (LQ) as  $10 \sigma/\text{slope}$  where  $\sigma$  is standard deviation of seven trace ( $0.005 \text{ mg L}^{-1}$ ) replicates [17]. The LD ranged from  $0.005 \text{ mg L}^{-1}$  (*p*-cresol) to  $0.124 \text{ mg L}^{-1}$  (2-butanone) and the LQ from  $0.014 \text{ mg L}^{-1}$  to  $0.351 \text{ mg L}^{-1}$ .

Accuracy and precision for each calibration range was determined by analysis of the mid-point concentration (Table 4;  $0.5 \text{ mg L}^{-1}$ ,  $5 \text{ mg L}^{-1}$  and  $50 \text{ mg L}^{-1}$ ). Accuracy was calculated as the ratio between measured and theoretical concentrations of 6 replicate solutions in different vials and precision was calculated as the RSD of 6 replicate injections from the same vial. According to the EPA method 8000C [18] and Little [20], accuracy and precision was classed as acceptable for all compounds at all calibration levels which was  $\leq 30\%$ . This also demonstrates sample stability after standing time which then permits repeat injections from the same vial.

### 3.3 *Solid phase extraction*

Solids phase extraction recovery efficiency was evaluated to permit calculation of recovery factors. Recoveries for *p*-cresol (90%), indole (81%) and skatole (88%) are comparable to those stated by Lin et al. [2] (Table 5). Further analytes with recoveries deemed to be either 'recommended' or 'acceptable' in accordance with EPA guidelines [21] were 2-butanone (56%), dimethyl disulfide (63%), 1-butanol (100%), benzaldehyde (77%) ethyl propionate (82%) and ethyl butyrate (89%). However, poor recoveries were identified for compounds including 1-propanol and limonene. We suggest that the poor extraction efficiency of 1-propanol can be ascribed to its strong affinity for water which limits the probability for partitioning onto the solid phase (Table 1,  $\text{LogK}_{\text{ow}}$  0.25). Conversely, the poor extraction efficiency for limonene can be attributed to its high volatility which increases the probability for sample losses at the vacuum and evaporation stages of sample preparation, coupled with its significant hydrophobicity ( $\text{logK}_{\text{ow}}$  4.57) which can initiate strong interactions with the sorbent that are known to inhibit SPE recovery [22]. Wells [23] recommended inclusion of an organic modifier for compounds with  $\text{logK}_{\text{ow}}$  exceeding 4, coupled with the addition of methanol to increase eluotropic strength and is recommended for improving SPE recovery for this compound in the future. Importantly, an RSD below 10% was recorded for each compound, which evidenced that SPE can achieve consistent recovery to within the acceptance criteria specified in the SPE EPA method 3535A (SW-846) [21], which demonstrates that correction factors could be applied (Table 5) to determine method detection limits (MDL). For illustration,

method detection limits for *p*-cresol ( $C_f$  200,  $R_f$  0.9) and indole ( $C_f$  200,  $R_f$  0.81) were 0.1 and 0.03  $\mu\text{g L}^{-1}$ . These values are several orders of magnitude lower than identified by De Preter et al. [19] using purge and trap with GCMS to determine faecal fermentation, which suggests direct determination from the liquid phase may enhance method sensitivity.

### 3.4 Characterisation of faecally contaminated urine

Liquid phase concentrations in urine and faecally contaminated urine samples from eleven volunteers were determined for the full-suite of VOCs except those which exhibited poor SPE recoveries (Table 6). In general, concentrations ranged between the MDL and 1  $\text{mg kg}^{-1}$  in urine samples, which is anticipated for fresh urine samples such as those measured in this study, which generally produce little odour when compared to aged urine [24]. The presence of indole and skatole in fresh urine is also evident in the literature, though concentrations were considered sufficiently low not to be impactful as an odorant [8]. However, *p*-cresol, was present at a considerable concentration (max. 13.01  $\text{mg kg}^{-1}$ ). Para-cresol arises in urine from the breakdown of tyrosine by cresol producing bacteria in the intestine [25]. Seigfried and Zimmerman [26] reported an average *p*-cresol concentration in urine of 18  $\text{mg L}^{-1}$ . This is similar to the maximum value, the broader variation potentially arising from various factors such as protein intake [27] and the presence of specific urease positive isolates (e.g. Enterobacteriaceae) [24] which are known contributors to raised *p*-cresol concentration. The use of 'mid-stream' urine collection techniques commonly used in medical studies (and not employed in this study) will also be expected to increase average concentration. Importantly, comparable values to the literature provide confirmation of the suitability of the method to real samples. Bacteria constitute 60% of faecal solids dry mass [28], with *Escherichia coli* (S8) representing the dominant bacterial species that is primarily responsible for the oxidation of fatty acids to alcohols, and the conversion of the amino acids tyrosine and tryptophan to *p*-cresol and indole and skatole respectively [27,29,30]. Faecally contaminated urine samples therefore generally exhibited higher VOC concentrations, and specifically for 1-butanol (alcohol) and indole (aromatic) which accords with the literature data on faecally contaminated urine [2]. The analytical data was compared to thresholds compiled from literature by van Gemert [31] used simply as a reference in order to contextualise the data (Table 6). At the background concentrations provided in urine, ethyl propionate, dimethyl disulfide, ethyl butyrate, *p*-cresol, indole and skatole were greater than the lower detection threshold for odour in water. The same VOC range was also above the taste threshold for water as was benzaldehyde. Significantly, each of the identified VOCs was determined in urine and faecally

contaminated urine samples, with several at elevated concentrations, which suggests that the VOC range selected is pertinent for the development of membrane technology for liquid phase odour abatement.

### 3.5 *Pervaporative membranes govern odour transport in faecally contaminated urine*

An initial mass balance was conducted across the membrane experimental set-up to confirm minimum VOC losses. A 10 ppm synthetic solution (comprising each VOC) was introduced to the feed-side and the mass balance constructed at the end of permeation was found to be  $100\pm 10\%$  (PVA used for assessment), which demonstrates the developed methods capability for technology assessment. An RSD of  $\leq 13\%$  was identified for replicate samples from membrane experimental studies. The membranes were subsequently challenged with the 10 ppm synthetic faecally contaminated urine. For the PVA membrane, removal efficiency ranged between  $60\pm 5\%$  (benzaldehyde) and  $85\pm 0.5\%$  (dimethyl disulfide, *p*-cresol) (Figure 3a). The separation efficiency can be accounted for by the selectivity of the polymer toward water, the intrinsic polarity increasing the solubility parameter of the polymer for water, whilst the lower molecular weight of water increases the diffusivity parameter for the polymer, the product of these two parameters providing an enhanced water permeability [32]. Whilst the presence of alcohols or carbonyl groups (e.g. benzaldehyde) are generally thought to influence the solubility parameter, a trend between VOC physico-chemical or structural properties (Table 1) was not evident [33]. This can be accounted for by the comparatively low partial pressure exhibited by the VOCs relative to water, which limits the associative driving force for separation. Baelen et al. [33] also observed that polyvinyl alcohol is soluble in water and prone to swelling above 20% wt. water. This results in an open membrane structure which decreases selectivity, and is exacerbated at elevated temperatures. In this study, the PVA membrane was used for illustrative purposes and the material used is recommended for separations comprising 50% wt. water solutions. Increasing crosslinking of the PVA polymer will increase membrane stability in the presence of water [34]. Therefore whilst good VOC separation was facilitated by the PVA membrane, optimisation of cross-linking is recommended for future investigation into PVA for liquid phase odour abatement.

For the hydrophobic PDMS membrane, permeate was enriched for all VOCs with enrichment factors ( $\beta$ ) ranging  $6.1\pm 0.8$  to  $35.9\pm 0.2$  (Equation 3, Figure 3b). The selectivity toward VOCs can be ascribed to the enhanced affinity of PDMS toward non-polar compounds [35]. A broad trend between the octanol-water coefficient, which corresponds to compound hydrophobicity, and enrichment factor was identified from benzaldehyde

(log  $K_{OW}$  = 1.48,  $\beta$  = 36) to ethyl propionate (log  $K_{OW}$  = 1.21,  $\beta$  = 27), 1-butanol (log  $K_{OW}$  = 0.88,  $\beta$  = 26) and 2-butanone (log  $K_{OW}$  = 0.29,  $\beta$  = 23). However, although *p*-cresol, indole, and skatole presented a stronger hydrophobic contribution (Table 1),  $\beta$  factors of 6-17 were identified for these compounds. In addition to compound mobility and solubility within PDMS, vapour pressure difference also governs separation [36]. The relatively lower permeability of these compounds can thus be accounted for by their vapour pressure which is around an order of magnitude lower than the other compounds. Since the PDMS polymer promotes VOC enrichment of the permeate, it is rational to expect an intensification of the 'repulsive' or 'nauseating' perception ordinarily associated with faecally contaminated urine (Table 6). However, the resulting permeate odour could be described as sweet, chemical, earthy and floral, with little perceivable evidence of faecal odour, and was hedonically more pleasant than the PVA permeate (Table 7). This suggests that barrier technology could be engineered to change perception through modification of the odour profile rather than developed simply for elimination. This is analogous to the perfumery industry in which indole, one of the core constituents of odour arising from faecally contaminated urine is similarly a critical ingredient in jasmine perfume [8].

#### 4. Conclusions

In this study, an analytical method for the detection of liquid phase VOCs responsible for faecal odour has been developed and verified. The following conclusions have been drawn:

- A quantitative method has been developed to enable co-elution of a range of VOCs comprised of a broad spectrum of physicochemical properties in a single elution.
- Comparison of this direct method to indirect methods (i.e. purge and trap) used to quantify the same compounds indicate an order of magnitude lower limit of detection can be achieved. The utility of this method extends to a broad range of stakeholders including healthcare professionals, sentient specialists and public health engineers.
- Consistent recovery was identified for solid phase extraction, while acceptable recoveries were also determined for nine VOCs, which were subsequently analysed in real matrices.
- Comparison of VOC data determined in urine and faecally contaminated urine samples to literature data, provided confirmation of the appropriateness of this method for evaluation of real samples, and also that the VOCs determined, are relevant and appropriate for the quantitation of faecal odour in the liquid phase.

- The method was successfully applied for the evaluation of pervaporative membranes, where SPE coupled with the lower calibration range, was capable of quantification within PVA membrane permeate which present an analytical challenge due to the polymers capability for separation.
- Whilst further membrane development is warranted for this application, the method was capable of facilitating diagnostic investigation of VOC separation and further demonstrated that the combination of hedonic characterisation coupled with quantitative methods are demanded to develop a technical solution for liquid phase odour separation, which offers significant potential for the advancement of decentralised sanitation.

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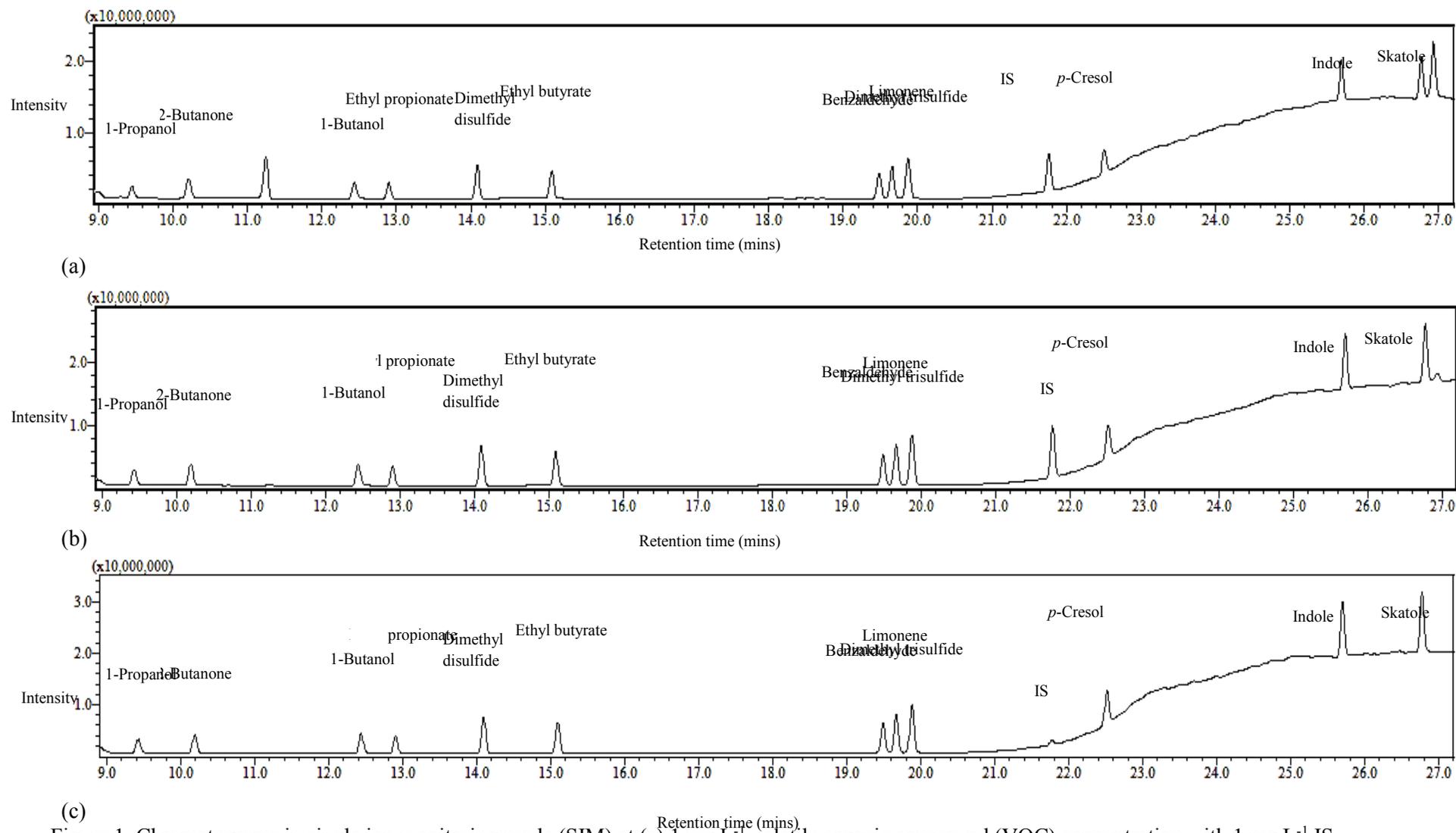


Figure 1. Chromatograms in single ion monitoring mode (SIM) at (a)  $1 \text{ mg L}^{-1}$  volatile organic compound (VOC) concentration with  $1 \text{ mg L}^{-1}$  IS concentration, (b)  $10 \text{ mg L}^{-1}$  VOC concentration with  $10 \text{ mg L}^{-1}$  IS concentration and (c)  $100 \text{ mg L}^{-1}$  VOC concentration with  $10 \text{ mg L}^{-1}$  IS concentration.

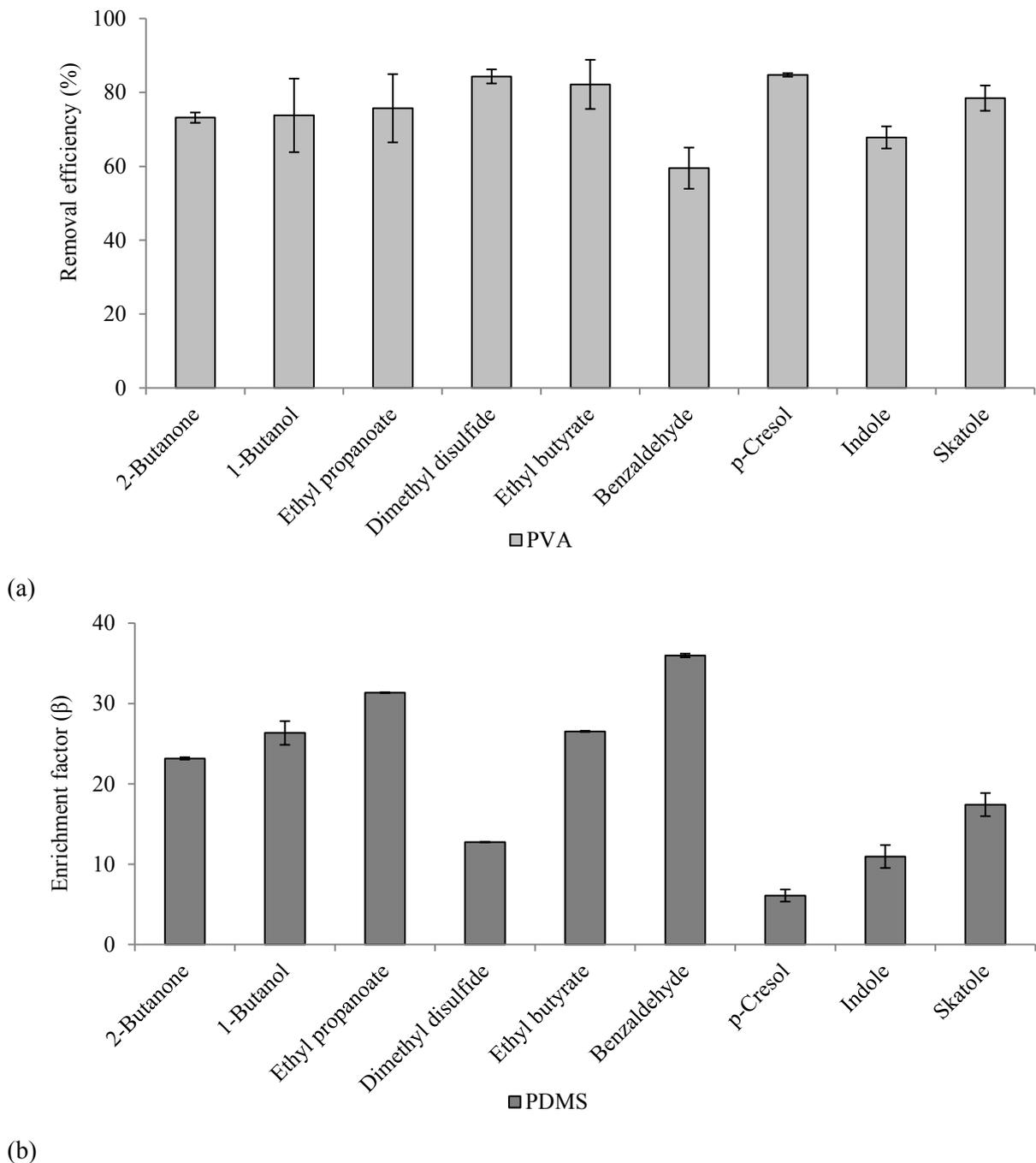
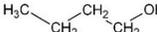
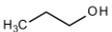
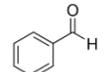
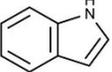
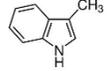
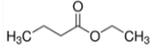
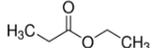
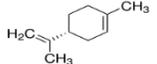
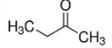
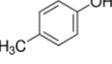
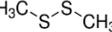
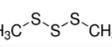


Figure 2. Assessment of pervaporative membrane processes as a liquid phase treatment to manage odour at source. Performance expressed as (a) removal efficiency for a hydrophilic membrane material and (b) enrichment factor for a hydrophobic membrane. PVA (Polyvinyl alcohol) and PDMS (Polydimethylsiloxane). Error bars represent the standard deviation of a triplicate at pH 6.5.

Table 1. Physico-chemical parameters of selected volatile organic compounds (VOCs) attributed to urine and faeces.

Compound	Chemical group	Chemical composition	Chemical structure	Molecular weight (g mol <sup>-1</sup> )	pKa	LogK <sub>ow</sub> at 20°C	Water solubility at 25°C (g L <sup>-1</sup> )	Henry's volatility constant at 25°C (mol m <sup>-3</sup> Pa)	Boiling point (°C)	Vapour pressure at 25°C (mm Hg)
1-Butanol	Alcohol	C <sub>4</sub> H <sub>9</sub> OH		74.12	16.1 <sup>a</sup>	0.88 <sup>a</sup>	63.2 <sup>a</sup>	1.2 <sup>d</sup>	111.7 <sup>a</sup>	7 <sup>a</sup>
1-Propanol	Alcohol	C <sub>3</sub> H <sub>8</sub> O		60.1	16.1 <sup>a</sup>	0.25 <sup>a</sup>	1000 <sup>a</sup>	1.5 <sup>d</sup>	97 <sup>a</sup>	14.9 <sup>a</sup>
Benzaldehyde	Aldehyde	C <sub>7</sub> H <sub>6</sub> O		106.12	14.9 <sup>a</sup>	1.48 <sup>a</sup>	6.95 <sup>a</sup>	0.38 <sup>d</sup>	178.1 <sup>a</sup>	1.27 <sup>a</sup>
Indole	Aromatic heterocycle	C <sub>8</sub> H <sub>7</sub> N		117.15	-3.6 <sup>c</sup>	2.14 <sup>a</sup>	3.56 <sup>a</sup>	19.1 <sup>d</sup>	254 <sup>a</sup>	0.0122 <sup>a</sup>
Skatole	Aromatic heterocycle	C <sub>9</sub> H <sub>9</sub> N		131.17	-4.6 <sup>c</sup>	2.6 <sup>a</sup>	0.498 <sup>a</sup>	4.7 <sup>d</sup>	265 <sup>a</sup>	0.0055 <sup>a</sup>
Ethyl butyrate	Ester	<a href="#">C<sub>6</sub>H<sub>12</sub>O<sub>2</sub></a>		116.16	-7 <sup>b</sup>	1.85 <sup>a</sup>	2.7 <sup>b</sup>	0.029 <sup>d</sup>	121 <sup>a</sup>	14 <sup>a</sup>
Ethyl propionate	Ester	C <sub>5</sub> H <sub>10</sub> O <sub>2</sub>		102.13	-7 <sup>b</sup>	1.21 <sup>a</sup>	19.2 <sup>a</sup>	0.041 <sup>d</sup>	98.9 <sup>a</sup>	35.8 <sup>a</sup>
Limonene	Hydrocarbon	C <sub>10</sub> H <sub>16</sub>		136.24	-4.2 <sup>b</sup>	4.57 <sup>a</sup>	0.013 <sup>a</sup>	0.00048 <sup>d</sup>	177 <sup>a</sup>	1.98 <sup>a</sup>
2-Butanone	Ketone	C <sub>4</sub> H <sub>8</sub> O		72.11	14.7 <sup>a</sup>	0.29 <sup>a</sup>	223 <sup>a</sup>	8.1 <sup>d</sup>	79.7 <sup>a</sup>	90.6 <sup>a</sup>
<i>p</i> -Cresol	Phenol	C <sub>7</sub> H <sub>8</sub> O		108.14	10.26 <sup>a</sup>	1.94 <sup>a</sup>	21.5 <sup>a</sup>	10 <sup>d</sup>	201.9 <sup>a</sup>	0.11 <sup>a</sup>
Dimethyl disulfide	Sulphur containing	C <sub>2</sub> HS <sub>2</sub>		94.19	-	1.77 <sup>a</sup>	3 <sup>a</sup>	0.0065 <sup>d</sup>	110 <sup>a</sup>	28.7 <sup>a</sup>
Dimethyl trisulfide	Sulphur containing	C <sub>2</sub> H <sub>6</sub> S <sub>3</sub>		126.25	-	1.926 <sup>a</sup>	2.39 <sup>a</sup>	0.021 <sup>d</sup>	170 <sup>a</sup>	1.06 <sup>a</sup>

<sup>a</sup> Pubchem (2017)

<sup>b</sup> YMDB (2017)

<sup>c</sup> Gu and Berry (1991)

<sup>d</sup> Sander (2015)

Table 2. Single ion monitoring (SIM) mass spectrometry parameters for target analytes.

Compound	Retention time (minutes)	Principal ion (m/z)	Reference ion 1 (m/z)	Reference ion 2 (m/z)
1-Propanol	9.455	31	42	59
2-Butanone	10.213	43	72	57
1-Butanol	12.437	56	41	43
Ethyl propionate	12.903	57	74	75
Dimethyl disulfide	14.087	94	79	45
Ethyl butyrate	15.087	71	43	88
Dimethyl trisulfide	19.478	126	79	45
Benzaldehyde	19.653	106	105	77
Limonene	19.862	68	93	67
<i>p</i> -Cresol	22.498	107	108	77
Indole	25.688	117	90	89
Skatole	26.76	130	131	77

Table 3. Calibration parameters for the target analytes; LD, limit of detection; LQ, limit of quantification; RF, response factor; RSD, relative standard deviation; SD, standard deviation.

	Calibration range (mg L <sup>-1</sup> )	Slope	Intercept	r <sup>2</sup>	LD <sup>a</sup> (mg L <sup>-1</sup> )	LQ <sup>b</sup> (mg L <sup>-1</sup> )	RF (% RSD)	Mean RF	RF SD
1-Propanol	10-100	0.531	0.00479	1.000			3.266	0.52	0.017
	1-10	0.456	0.00382	0.992			10.48	0.455	0.047
	0.005-1	0.674	0.00677	1.000	0.019	0.077	11.66	0.75	0.087
2-Butanone	10-100	0.874	0.07026	0.999			2.57	0.89	0.023
	1-10	0.817	0.01183	0.994			8.429	0.84	0.07
	0.005-1	1.39	0.194	0.991	0.124	0.351	14.53	1.87	0.27
1-Butanol	10-100	0.4144	0.01558	1.000			2.71	0.63	0.017
	1-10	0.381	-0.00232	0.996			9.54	0.365	0.034
	0.005-1	0.468	-0.00213	0.999	0.036	0.099	17.93	0.436	0.078
Ethyl propionate	10-100	0.614	0.0693	0.999			2.68	0.99	0.027
	1-10	0.588	0.00609	0.995			0.914	0.59	0.054
	0.005-1	0.753	0.00412	0.999	0.011	0.045	4.975	0.78	0.039
Dimethyl disulfide	10-100	0.976	0.0939	0.999			1.81	0.468	0.0085
	1-10	1.08	0.00586	0.998			9.16	1.07	0.098
	0.005-1	1.2003	0.00316	1.000	0.005	0.019	7.25	1.17	0.085
Ethyl butyrate	10-100	0.462	0.0301	1.000			2.5	0.729	0.018
	1-10	0.49	0.00362	0.997			8.82	0.49	0.043
	0.005-1	0.561	0.00184	1.000	0.006	0.026	16.54	0.545	0.09
Dimethyl trisulfide	10-100	0.718	0.047	0.999			2.03	0.71	0.014
	1-10	0.797	-0.000266	0.997			8.86	0.78	0.069
	0.005-1	0.8114	0.000512	1.000	0.010	0.031	13.01	0.769	0.1
Benzaldehyde	10-100	0.685	0.0823	0.999			3.09	0.5	0.015
	1-10	0.731	0.004297	0.997			8.03	0.73	0.0587
	0.005-1	0.76	0.00259	1.000	0.005	0.021	5.38	0.76	0.041
Limonene	10-100	0.479	0.0741	1.000			3.09	0.503	0.0156
	1-10	0.474	0.00898	0.997			5.7	0.495	0.028
	0.005-1	0.529	0.0105	0.999	0.041	0.165	8.85	0.568	0.05
<i>p</i> -Cresol	10-100	0.69	0.0809	1.000			2.42	0.741	0.0173
	1-10	0.71	0.00152	0.997			9.87	0.7	0.069
	0.005-1	0.681	-0.000495	0.999	0.019	0.057	18.46	0.698	0.129
Indole	10-100	1.39	0.545	0.996			5.74	1.56	0.0896
	1-10	1.49	0.0121	0.997			7.57	1.5	0.113
	0.005-1	1.433	0.00895	1.000	0.005	0.027	12.523	1.56	0.195
Skatole	10-100	1.509	0.558	0.994			5.3	1.67	0.089
	1-10	1.6625	0.00763	0.998			7.51	1.66	0.12
	0.005-1	1.519	0.00755	1.000	0.005	0.014	13.1862	1.6	0.211

<sup>a</sup>LD calculated as  $3.3 \sigma / \text{Slope}$ , where  $\sigma$  is standard deviation of seven 0.005 mg L<sup>-1</sup> replicates (Currie, 1999).

<sup>b</sup>LQ calculated as  $10 \sigma / \text{Slope}$ , where  $\sigma$  is standard deviation of seven 0.005 mg L<sup>-1</sup> replicates (Currie, 1999).

Note: RSD is acceptable when < 20 % (EPA, 2003)

Table 4. Precision and accuracy for each analyte within three calibration ranges; RSD, residual standard deviation.

	0.5 mg L <sup>-1</sup>			5 mg L <sup>-1</sup>			50 mg L <sup>-1</sup>		
	Mean (mg L <sup>-1</sup> )	Accuracy <sup>a</sup> (%)	Precision <sup>b</sup> (RSD)	Mean (mg L <sup>-1</sup> )	Accuracy <sup>a</sup> (%)	Precision <sup>b</sup> (RSD)	Mean (mg L <sup>-1</sup> )	Accuracy <sup>a</sup> (%)	Precision <sup>b</sup> (RSD)
1-Propanol	0.46 ± 0.08	92.6	3.7	5.25 ± 0.04	104.9	0.3	50.70 ± 1.50	101.4	1.6
2-Butanone	0.45 ± 0.07	89.6	15.6	5.16 ± 0.08	103.2	0.8	49.14 ± 3.69	98.3	1.7
1-Butanol	0.50 ± 0.02	100.6	5.34	5.15 ± 0.05	103.1	1.4	49.59 ± 5.05	99.2	2.4
Ethyl propionate	0.51 ± 0.07	102.7	3.6	5.19 ± 0.05	103.7	0.7	50.13 ± 1.89	100.3	1.1
Dimethyl disulfide	0.54 ± 0.06	107.6	3.6	5.15 ± 0.03	103.1	0.6	50.27 ± 1.24	100.5	1.5
Ethyl butyrate	0.55 ± 0.06	109.3	3.6	5.18 ± 0.03	103.6	0.7	50.04 ± 2.41	100.1	1.6
Dimethyl trisulfide	0.58 ± 0.05	115.3	2.5	5.12 ± 0.01	102.4	1.1	50.96 ± 1.82	101.9	1.7
Benzaldehyde	0.53 ± 0.07	106.7	0.9	5.16 ± 0.02	103.1	0.2	49.68 ± 1.44	99.4	1.4
Limonene	0.39 ± 0.12	77.3	0.9	5.17 ± 0.03	103.4	0.4	49.28 ± 1.02	98.6	1.6
<i>p</i> -Cresol	0.50 ± 0.07	100.4	4.6	5.13 ± 0.01	102.7	2.3	50.39 ± 1.53	100.8	2.7
Indole	0.47 ± 0.08	93.8	2.5	5.17 ± 0.02	103.4	0.7	49.55 ± 1.89	99.1	1.0
Skatole	0.49 ± 0.07	98.4	7.0	5.13 ± 0.03	102.7	0.4	49.55 ± 2.80	99.1	1.9

<sup>a</sup>Accuracy calculated as the percentage ratio between measured and theoretical concentrations of 6 replicate solutions in different vials

<sup>b</sup>Precision calculated as the RSD of 6 replicated injections from the same vial.

Note: 1. Accuracy acceptance: ≤ 30%. (EPA, 2003).

2. Criteria for precision: ≤ 25% is excellent, less than or equal to 30% is acceptable (Little, 2016)

Table 5. SPE recovery factors, RSD, residual standard deviation.

	SPE recovery <sup>a</sup> (% ± RSD) in this study			Average SPE recovery (%) All trials	SPE recovery (%) Lin et al., (2013)		
	pH 2	pH 6.5	pH 9		pH 5	pH 6	pH 7
1-Propanol	21 ± 1	26 ± 4	21 ± 5	22 ± 3			
2-Butanone	64 ± 4	52 ± 2	53 ± 3	56 ± 7			
1-Butanol	106 ± 5	106 ± 2	100 ± 6	100 ± 4			
Ethyl propionate	85 ± 2	79 ± 4	83 ± 3	82 ± 3			
Dimethyl disulfide	69 ± 4	54 ± 3	66 ± 2	63 ± 8			
Ethyl butyrate	84 ± 4	95 ± 4	89 ± 3	89 ± 6			
Dimethyl trisulfide	55 ± 2	44 ± 2	51 ± 2	50 ± 6			
Benzaldehyde	76 ± 2	77 ± 3	79 ± 2	77 ± 2			
Limonene	23 ± 2	24 ± 2	21 ± 1	22 ± 2			
<i>p</i> -Cresol	96 ± 6	90 ± 6	83 ± 4	89 ± 7	103 ± 5	97 ± 0.5	103 ± 11
Indole	80 ± 7	82 ± 6	81 ± 6	81 ± 1	89 ± 2	90 ± 16	96 ± 2
Skatole	87 ± 5	89 ± 5	89 ± 5	88 ± 2	96 ± 5	97 ± 9	100 ± 2

<sup>a</sup> SPE recovery calculated as the percentage ratio between SPE measured and theoretical concentrations (100 mg L<sup>-1</sup> injection concentration representing the upper calibration limit)

Note: 1. SPE recovery recommended as: 70 – 130 % (EPA, 2007).

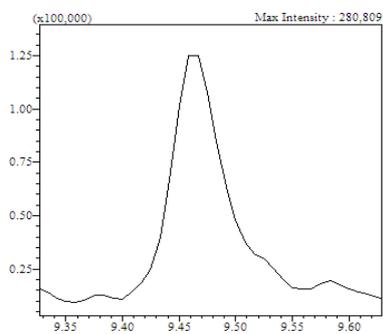
2. RSD acceptance: ≤ 30 % (EPA, 2007)

Table 6. Typical concentrations of 12 liquid phase VOCS attributed to fresh urine and faeces with associated odour descriptors and detection thresholds in air and water.

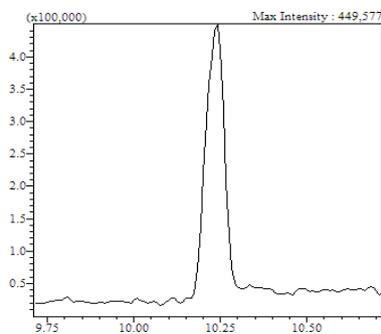
	Odour descriptor*	Urine	Faecally contaminated urine. 10:1 urine to faeces ratio		Faeces [2]	Detection threshold [31]		
		N = 11 Range (mg kg <sup>-1</sup> urine)	N = 11 Range (mg kg <sup>-1</sup> urine)      Range (mg kg <sup>-1</sup> faeces)		N = 2 Range (mg kg <sup>-1</sup> faeces)	Air (odour) Range (mg m <sup>-3</sup> )	Water (odour) Range (mg kg <sup>-1</sup> )	Water (taste) Range (mg kg <sup>-1</sup> )
2-Butanone	Acetone like	<LD-1.323	0.014-0.315	0.140-3.146		0.21-1000	7-100	3-60
1-Butanol	Alcohol like	<LD-0.016	<LD-0.185	<LD-1.846		0.015-3000	0.27-511	2-100
Ethyl propionate	Fruity, rum	<LD-0.008	<LD-0.02	<LD-0.198		0.3-1	0.0001-0.067	0.00049-0.004
Dimethyl disulfide	Rotten cabbage	<LD-0.013	<LD-0.014	<LD-0.142		0.0011-3.5	0.00016-0.09	0.03-0.068
Ethyl butyrate	Pineapple	<LD-0.006	<LD-0.02	<LD-0.197		0.000016-0.1	0.000001-0.4	0.0001-0.45
Benzaldehyde	Bitter almond	<LD-0.060	0.0009-0.012	0.009-0.107		0.01-3400	0.32-4.6	0.05-1.5
<i>p</i> -Cresol	Sweet, tar-like	0.003-13.01	0.214-2.67	2.139-26.683	20-25	0.00002	0.055-0.2	0.002-0.018
Indole	Faecal	<LD-0.514	0.012-1.001	0.113-10.015	5-8	0.00035-0.0071	0.13-0.59	0.5
Skatole	Faecal, nauseating	<LD-0.045	0.007-0.162	0.074-1.619	2-6	0.00035-0.00078	0.0002-0.052	0.05

Table 7. Odour descriptors for the membrane permeates of faecally contaminated urine

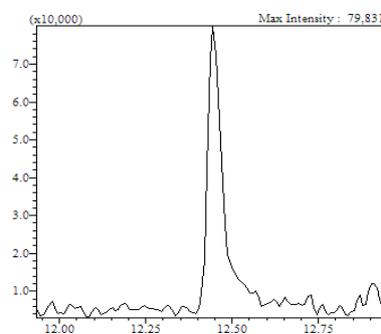
Membrane material	Permeate odour descriptor
Polyvinyl alcohol	Sweaty, chemical, sweet, onion
Polydimethylsiloxane	Sweet, chemical, earthy, floral



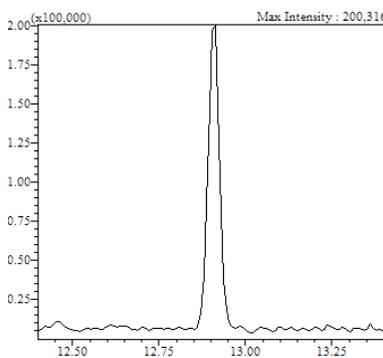
(a)



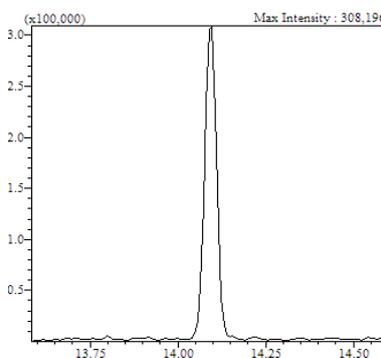
(b)



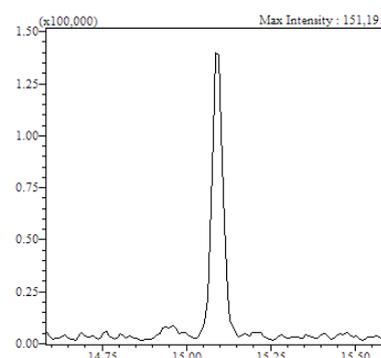
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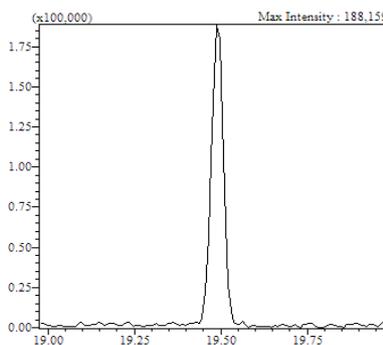
(d)



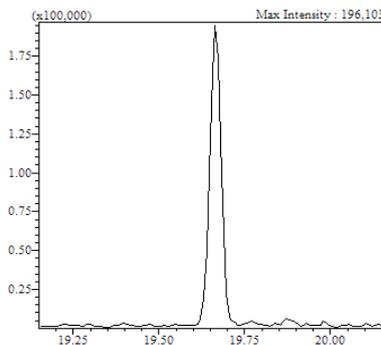
(e)



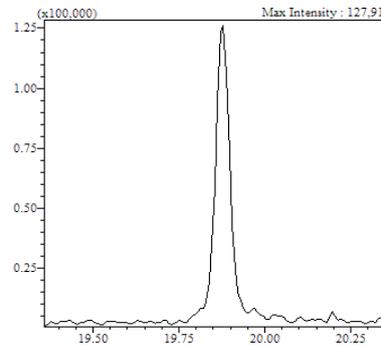
(f)



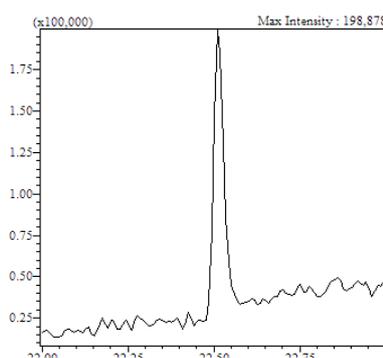
(g)



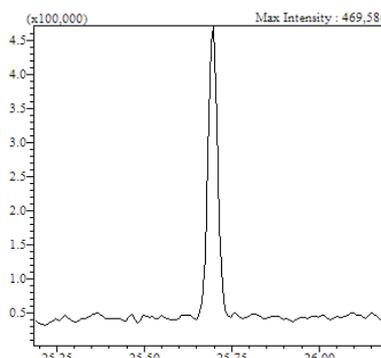
(h)



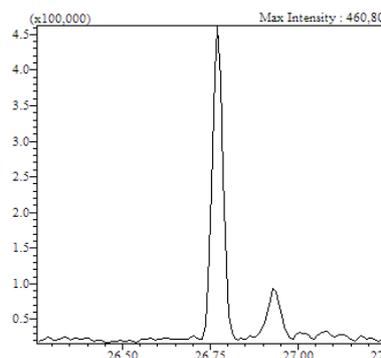
(i)



(j)

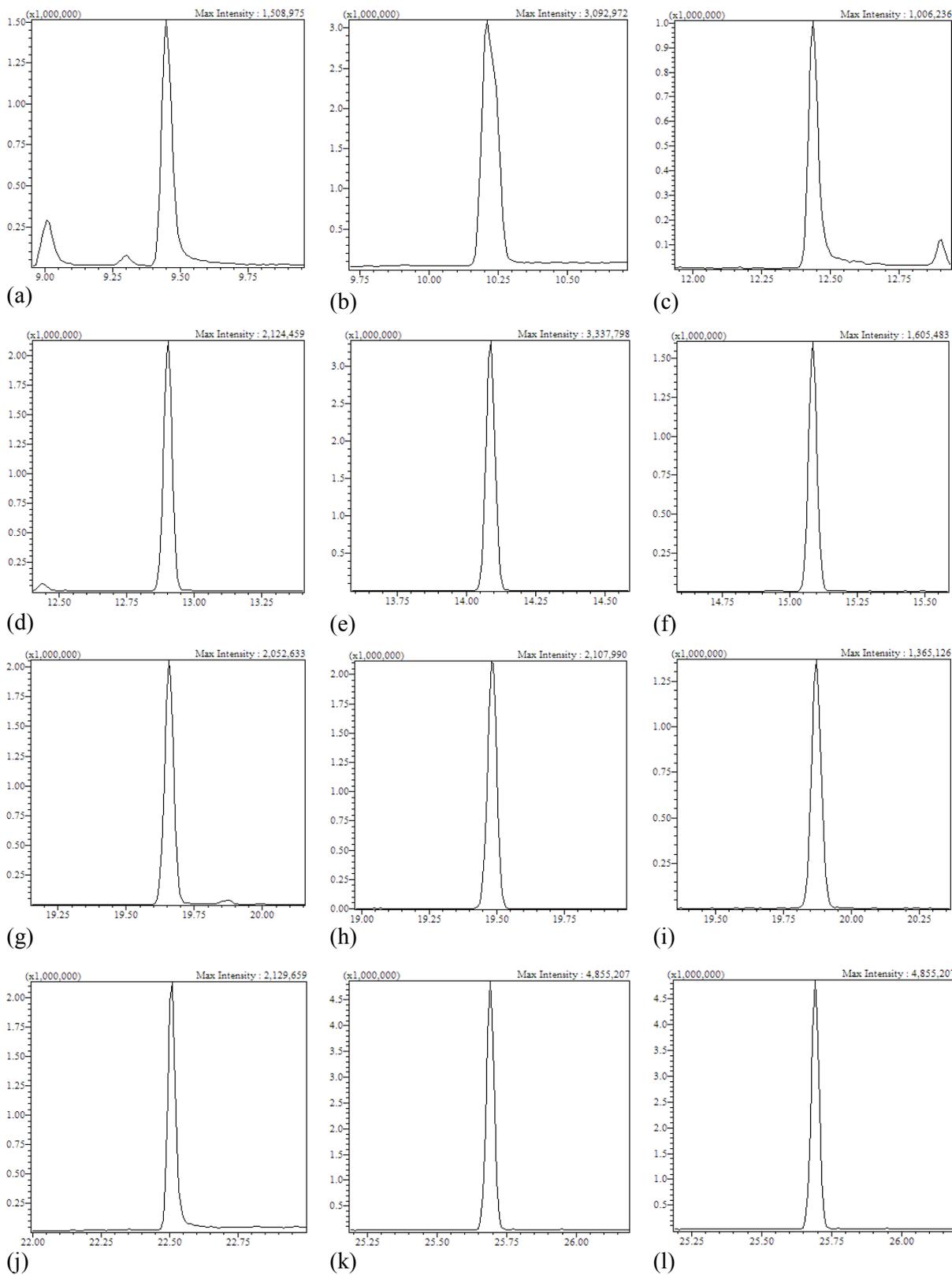


(k)



(l)

S1. Intensity vs. retention time of individual peaks at  $0.1 \text{ mg L}^{-1}$ . (a) 1-Propanol (b) 2-Butanone (c) 1-Butanol (d) Ethyl propionate (e) Dimethyl disulfide (f) Ethyl butyrate (g) Dimethyl trisulfide (h) Benzaldehyde (i) Limonene (j) *p*-Cresol (k) Indole (l) Skatole.



S2. Intensity vs. retention time of individual peaks at 1 mg L<sup>-1</sup>. (a) 1-Propanol (b) 2-Butanone (c) 1-Butanol (d) Ethyl propionate (e) Dimethyl disulfide (f) Ethyl butyrate (g) Dimethyl trisulfide (h) Benzaldehyde (i) Limonene (j) *p*-Cresol (k) Indole (l) Skatole.

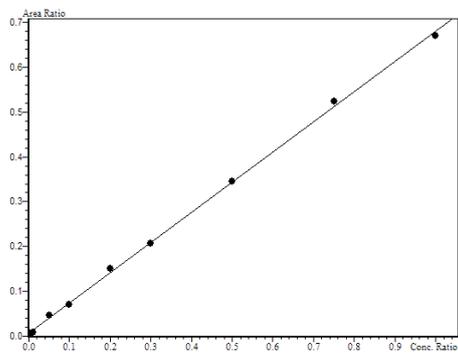
S3. USP coefficient tailing factors for 0.1 mg L<sup>-1</sup> and 1 mg L<sup>-1</sup> peaks.

	0.1 mg L <sup>-1</sup>	1 mg L <sup>-1</sup>
1-Propanol	1.40	1.39
2-Butanone	0.86	1.57
1-Butanol	2.04	1.34
Ethyl propionate	1.00	1.02
Dimethyl disulfide	1.08	1.03
Ethyl butyrate	0.96	1.00
Dimethyl trisulfide	0.96	0.90
Benzaldehyde	0.96	0.93
Limonene	1.10	1.06
<i>p</i> -Cresol	1.24	1.25
Indole	1.00	0.98
Skatole	0.95	0.92

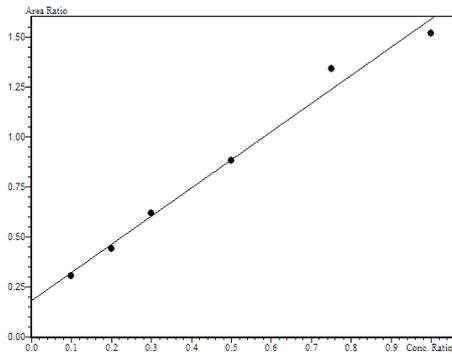
USP tailing factor calculated as  $W_{0.05}/2f_{0.05}$ , where  $W_{0.05}$  is the peak width and  $f$  is the peak front width at 5% height.

Note: 1. Good chromatographic peak shape defined as symmetrical, narrow and tailing factor of 1 (Agilent, ND).

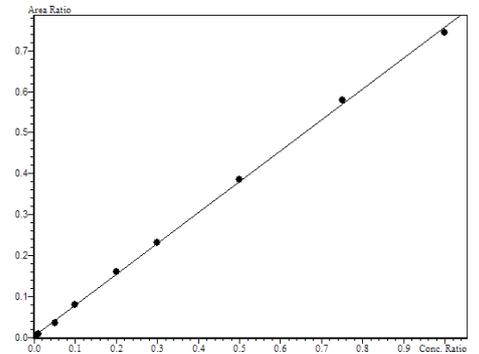
2. U.S. Food and Drug Administration (FDA). The FDA recommends a tailing factor of  $\leq 2$



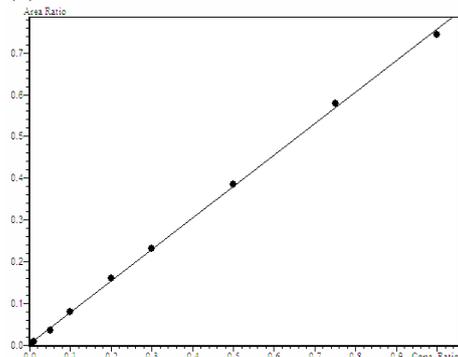
(a)



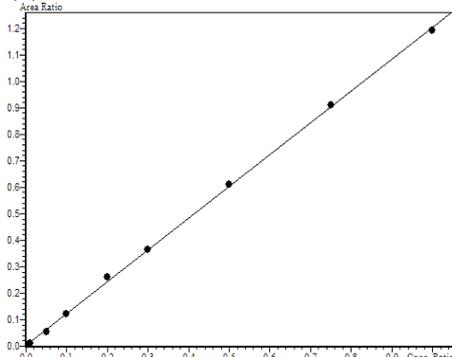
(b)



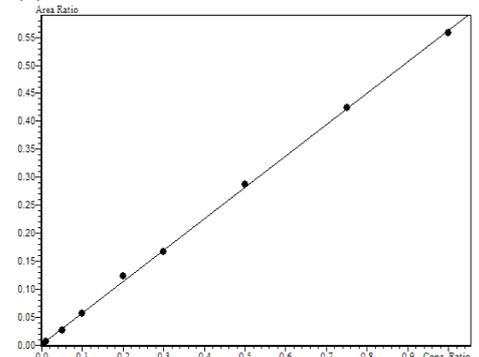
(c)



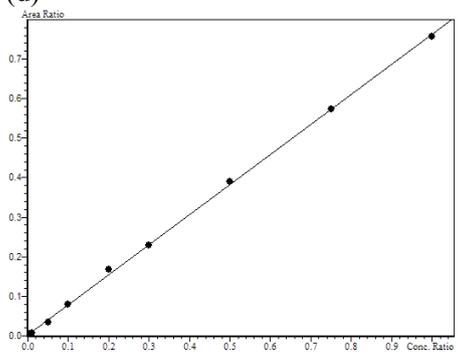
(d)



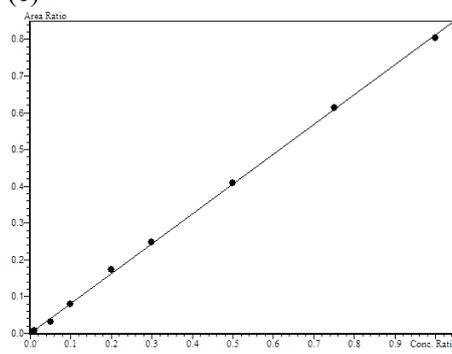
(e)



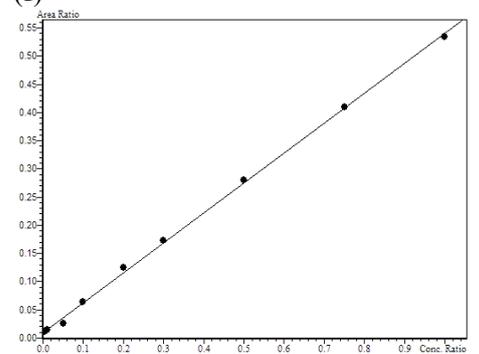
(f)



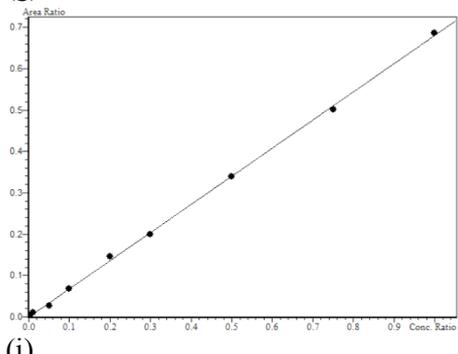
(g)



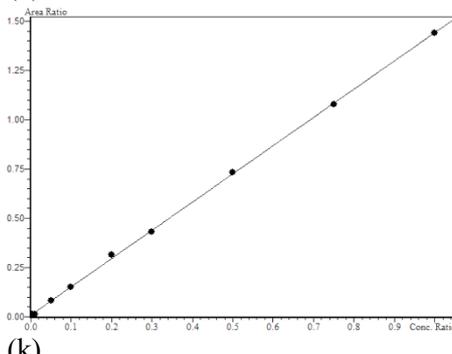
(h)



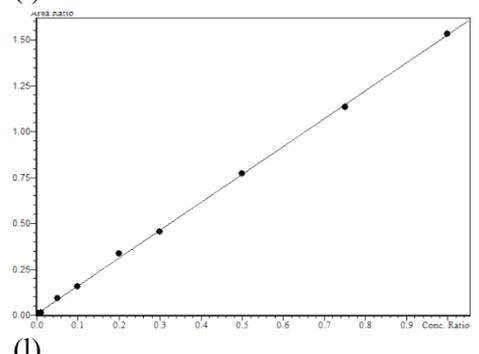
(i)



(j)

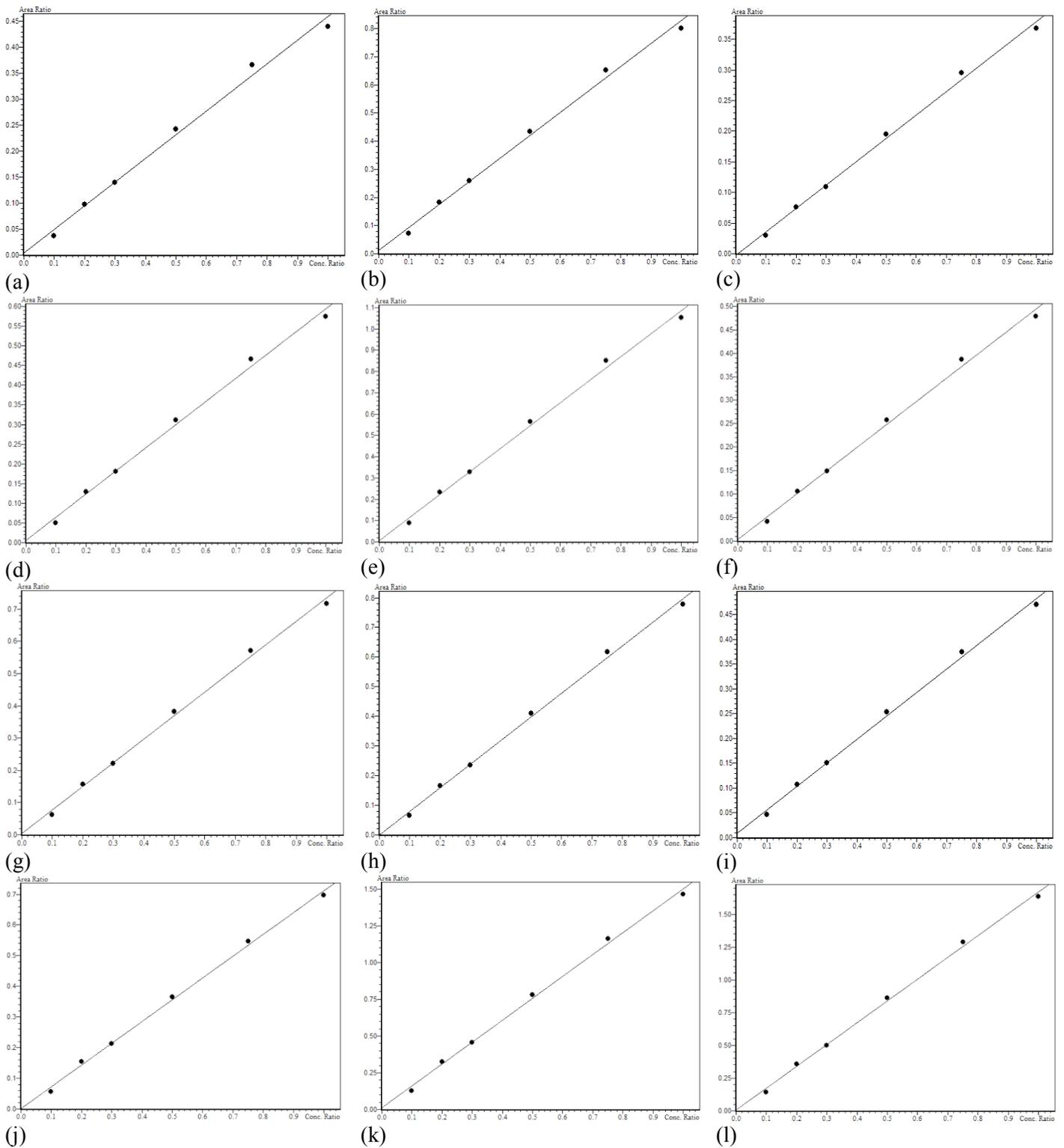


(k)

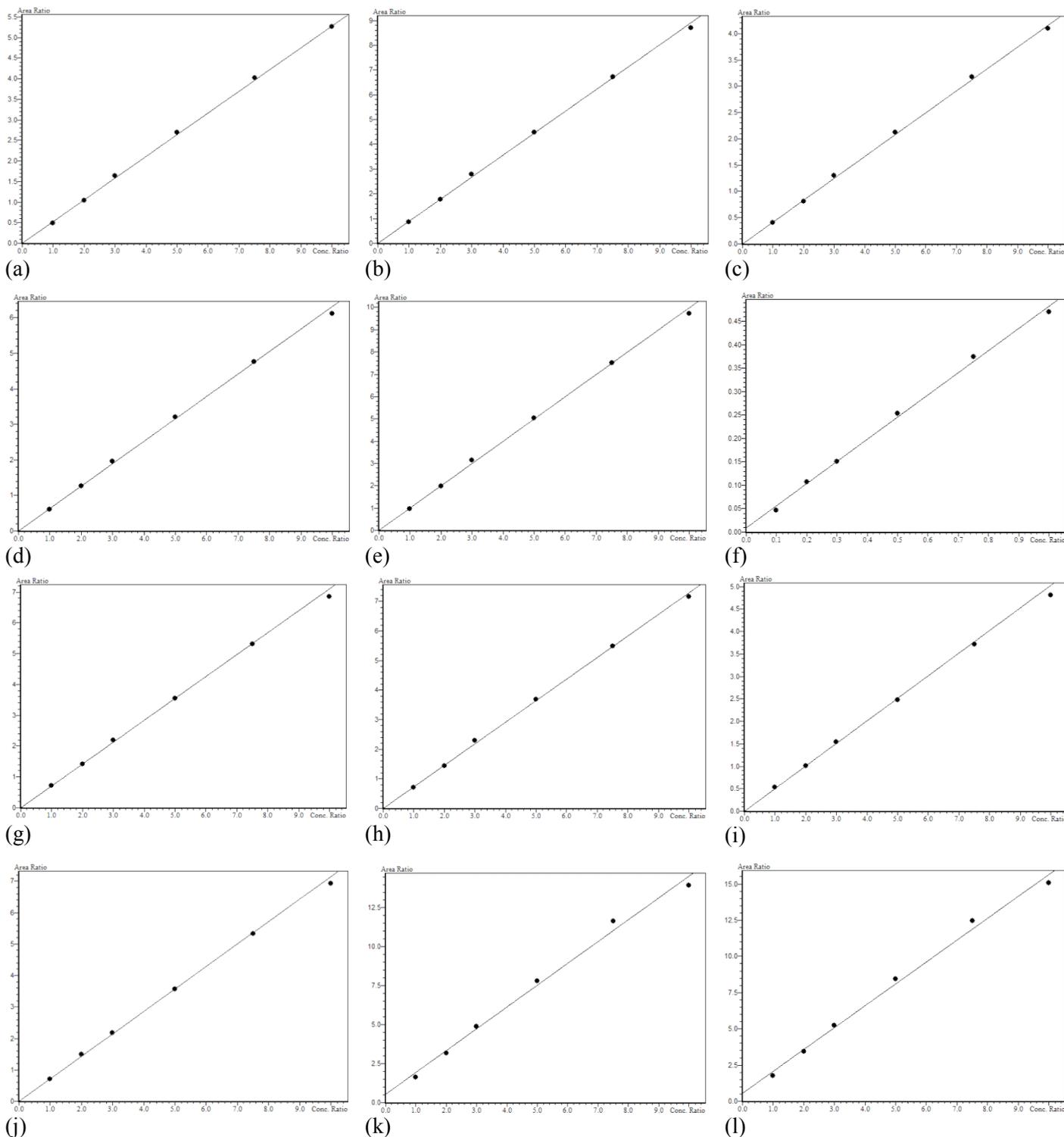


(l)

S4. Calibration curves\* for a range LD - 1 mg L<sup>-1</sup>. (a) 1-Propanol (b) 2-Butanone (c) 1-Butanol (d) Ethyl propionate (e) Dimethyl disulfide (f) Ethyl butyrate (g) Dimethyl trisulfide (h) Benzaldehyde (i) Limonene (j) *p*-Cresol (k) Indole (l) Skatole. \*Calibration curves based on area ratio (AREA<sub>VOC</sub>/AREA<sub>IS</sub>) vs concentration ratio (CONCENTRATION<sub>VOC</sub>/CONCENTRATION<sub>IS</sub>).



S5. Calibration curves\* for a range 1-10 mg L<sup>-1</sup>. (a) 1-Propanol (b) 2-Butanone (c) 1-Butanol (d) Ethyl propionate (e) Dimethyl disulfide (f) Ethyl butyrate (g) Dimethyl trisulfide (h) Benzaldehyde (i) Limonene (j) *p*-Cresol (k) Indole (l) Skatole. \*Calibration curves based on area ratio (AREA<sub>VOC</sub>/AREA<sub>IS</sub>) vs concentration ratio (CONCENTRATION<sub>VOC</sub>/CONCENTRATION<sub>IS</sub>).



S6. Calibration curves\* for a range 10-100 mg L<sup>-1</sup>. (a) 1-Propanol (b) 2-Butanone (c) 1-Butanol (d) Ethyl propionate (e) Dimethyl disulfide (f) Ethyl butyrate (g) Dimethyl trisulfide (h) Benzaldehyde (i) Limonene (j) *p*-Cresol (k) Indole (l) Skatole. \*Calibration curves based on area ratio (AREA<sub>VOC</sub>/AREA<sub>IS</sub>) vs concentration ratio (CONCENTRATION<sub>VOC</sub>/CONCENTRATION<sub>IS</sub>).

S7. Membrane characteristics and operating conditions.

	Pervaporation	
Manufacturer (model)	DeltaMem (Pervap™ 4101)	Permselect (PDMSXA–2500)
Material	Polyvinyl alcohol (PVA)	Polydimethylsiloxane (PDMS)
Membrane area (m <sup>2</sup> )	0.0153	0.25
Membrane thickness (μm)	5	55
Membrane structure	Crosslinked support layer	Symmetric
Contact angle (°)	43 (±1.1)	116 (±1.4)
Geometry	Flat sheet	Hollow fibre
Operating pressure (bar)	0.05	0.05
Operating membrane temperature (°C)	50	50

Note: >90° indicates hydrophobic polymer and <90° indicates hydrophilic polymer

Table S8. VOC microbial sources of selected compounds in this study.

Volatile organic compound	Microbial source
1-Propanol	<i>Escherichia coli</i> <sup>a</sup>
2-Butanone	<i>Pseudomonas aeruginosa</i> <sup>a</sup>
1-Butanol	<i>Staphylococcus aureus</i> , <i>Escherichia coli</i> <sup>a</sup>
Ethyl propionate	-
Dimethyl disulfide	<i>Klebsiella pneumoniae</i> <sup>a</sup> , <i>Pseudomonas aeruginosa</i> <sup>a</sup> , <i>Escherichia coli</i> <sup>a</sup>
Ethyl butyrate	<i>Enterococcus faecalis</i> <sup>a</sup> , <i>Escherichia coli</i> <sup>a</sup>
Dimethyl trisulfide	<i>Pseudomonas aeruginosa</i> <sup>a</sup>
Benzaldehyde	<i>Staphylococcus aureus</i> <sup>a</sup> , <i>Escherichia coli</i> <sup>a</sup>
Limonene	<i>Pseudomonas aeruginosa</i> <sup>a</sup>
<i>p</i> -Cresol	Most aerobic enterobacteria and anaerobic <i>Clostridium perfringens</i> <sup>b</sup>
Indole	<i>Escherichia coli</i> <sup>a</sup>
Skatole	<i>Escherichia coli</i> <sup>a</sup>

2 <sup>a</sup>B.Bos et al. 20133 <sup>b</sup>Vanholder et al., 1999

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