Controls on the Fate and Speciation of Np(V) During Iron (Oxyhydr)oxide Crystallization
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Supporting Information

ABSTRACT: The speciation and fate of neptunium as Np(V)O₄²⁻ during the crystallization of ferrihydrite to hematite and goethite was explored in a range of systems. Adsorption of NpO₂⁴⁺ to iron(III) (oxyhydr)oxide phases was reversible and, for ferrihydrite, occurred through the formation of mono-nuclear bidentate surface complexes. By contrast, chemical extractions and X-ray absorption spectroscopy (XAS) analyses showed the incorporation of Np(V) into the structure of hematite during its crystallization from ferrihydrite (pH 10.5). This occurred through direct replacement of octahedrally coordinated Fe(III) by Np(V) in neptunate-like coordination. Subsequent analyses on mixed goethite and hematite crystallization products (pH 9.5 and 11) showed that Np(V) was incorporated during crystallization. Conversely, there was limited evidence for Np(V) incorporation during goethite crystallization at the extreme pH of 13.3. This is likely due to the formation of a Np(V) hydroxide precipitate preventing incorporation into the goethite particles. Overall these data highlight the complex behavior of Np(V) during the crystallization of iron(III) (oxyhydr)oxides, and demonstrate clear evidence for neptunium incorporation into environmentally important mineral phases. This extends our knowledge of the range of geochemical conditions under which there is potential for long-term immobilization of radiotoxic Np in natural and engineered environments.

INTRODUCTION

Many nations have substantial nuclear legacy wastes due to their civil and military nuclear programs. These materials contain significant levels of Np-237, an α-emitting radionuclide with a long half-life (2.14 × 10⁸ y) which is highly radiotoxic and will be a significant dose contributing radionuclide in radioactive wastes on longer time scales. Understanding the environmental behavior of Np is a key consideration at nuclear legacy sites such as Sellafield and Hanford, and for deep geological disposal of radioactive wastes. In oxic environments, neptunium exists predominantly in the pentavalent oxidation state (Np(V) as neptunyl, NpO₂⁴⁺) with poorly soluble tetravalent neptunium (Np(IV) as Np⁴⁺) dominant in more reducing environments. Np(V) dominates in a wide range of oxic aqueous environments and is expected to be the most soluble and mobile of the actinide species (solubility ∼100 μM Np(V) at pH = 7). Np(V) environmental mobility is controlled by a number of processes including its intrinsic solubility and sorption to minerals (e.g., iron (oxyhydr)oxide phases). Due to the reduced solubility of metal ions/radionuclides at elevated pH, base addition has been used to treat radioactive wastes and cementitious materials have been included in the engineering of geological disposal scenarios with the intention that this will increase pH and thus decrease the solubility and mobility of radionuclides. For Np(V), the thermodynamic solubility decreases from ∼100 μM at pH 7 to ∼0.1 μM at pH 10. With this in mind, we have focused on alkaline pH conditions (pH > 9) to allow specific insights into the fate of Np(V) during the transformation of ferrihydrite to crystalline iron (oxyhydr)oxide phases.

Iron(III) (oxyhydr)oxide phases such as ferrihydrite (Fe₈₋₁₂(OH)₄), hematite (Fe₂O₃) and goethite (FeOOH) are ubiquitous throughout the environment and

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will be present in nuclear decommissioning scenarios and radioactive waste materials. Previous studies have shown that Np(V) strongly interacts with iron(III) (oxyhydr)oxide surfaces such as ferrihydrite,21,22 lepidocrocite,23 hematite,24–27 and goethite8,20 with high distribution coefficients ($K_D$ up to 10$^6$ mL/g). These interactions typically occur through the formation of inner sphere Np(V) complexes25 and, where carbonate is present, via Np(V)-carbonate surface complexes.24 Additionally, the surface complexation behavior of Np(V) varies significantly with both pH21,22,25–27 and carbonate concentration.24,27 Despite its potential environmental significance, the mechanism(s) of Np(V) adsorption to ferrihydrite are not well understood, in particular spectroscopic information on Np(V)-ferrihydrite surface complexes is lacking. Additionally, there is a paucity of information on Np(V) incorporation into mineral structures despite the fact that this is potentially a highly effective mechanism to immobilize radionuclides. Interestingly, an early study suggested that Np(V) adsorbed to ferrihydrite may be incorporated during subsequent crystallization to hematite.22 However, to date, no direct spectroscopic evidence exists for Np(V) incorporation into iron (oxyhydr)oxides in environmentally relevant systems. In comparison, spectroscopic studies have shown that both U(VI) and U(V) can substitute for Fe(III) in the structure of hematite.31–33 This suggests, on the basis of the expected similarities in behavior of An(V, VI) species, that Np(V) may also undergo similar incorporation. Such sequestration processes have the potential to significantly limit the mobility of Np(V) in environmental systems. Furthermore, they could lead to irreversible binding of the radionuclide as long as the mineral-radionuclide assemblage is stable. Understanding the mechanisms of Np(V) incorporation and the factors controlling these processes will be essential in exploiting these reactions for remediating radionuclide contamination, and informing the long-term behavior of Np(V) in natural environments and nuclear decommissioning/radioactive waste management scenarios.

The aim of this study was to determine the potential for Np(V) incorporation into hematite and goethite during the crystallization of ferrihydrite under alkaline conditions. We induced the crystallization of ferrihydrite to hematite, goethite, and mixtures of these two minerals (Table SI1). Ferrihydrite was prepared according to the method of Cornell and Schwertmann.34,35 For the nonactive crystallization experiments 0.2 mM Ca(OH)$_2$ was mixed with ferrihydrite at a solid solution ratio of 0.5 g l$^{-1}$. The pH was adjusted to either 9.5 or 11 using 10 M NaOH, and the crystallization reactions were induced at elevated temperature (80 °C) for up to 2 weeks. Two additional experiments designed to produce pure hematite or goethite were also performed. Hematite crystallization was induced in pH 10.5 NaOH with a 5 g l$^{-1}$ ferrihydrite loading reacted at 120 °C for 4 weeks.35–37 Goethite crystallization was induced in 0.1 M NaOH, 0.1 M KOH, and 0.3 mM Ca(OH)$_2$, at pH 13.3 with a 0.5 g l$^{-1}$ ferrihydrite loading at room temperature for 2 weeks.36,37 Throughout, experiments were run in a CO$_2$-free atmosphere. The crystallization products were analyzed using a Bruker D8 X-ray diffractometer (XRD); Rietveld analyses were performed on the XRD patterns using Topas 4. Transmission electron microscopy images were taken on the final crystallization products using a Philips CM200 field emission gun TEM.

**Neptunium Experiments.** The ferrihydrite crystallization experiments to produce pure hematite and goethite, and the systems at pH 9.5 and 11 were repeated with 0.04–100 μM Np(V)O$_4^-$ (spiked with a 0.16 or 1.6 mM Np(V) stock solution) under the same conditions as the nonactive experiments to yield parallel sets of both low and high Np(V) concentration experiments. The Np(V) stock solution was added to the experiment either at the start of the ferrihydrite crystallization, to assess incorporation processes during crystallization, or to the final crystallization products to assess adsorption processes (Table SI1). At the end of the experiments, aliquots of the experimental suspensions were filtered (<0.2 μm, nylon) to quantify the Np(V) concentration in solution. To assess the partitioning of Np(V) into the solids, a chemical extraction was performed by acidifying experiments with HCl to pH 2.5–3, which is below the adsorption edge of Np(V) on iron(III) (oxyhydr)oxides.35,30,33 After 2 h, aliquots of the extracted suspensions were filtered (<0.2 μm nylon) and analyzed to determine the acid extractable Np(V) and to calculate the fraction of incorporated Np(V). All filtered solution samples were analyzed for total Np(V) using ICP-MS with Th-232 as an internal standard (Agilent 7500cx). Finally, the Np(V) containing solid phases from the experiments with elevated Np(V) (>2000 ppm of Np(V) associated with the solids) were sampled by centrifuging the experimental slurries and decanting the supernatant. The remaining solids were prepared as wet pastes for XAS analyses, mounted in approved XAS cells, and stored under an inert atmosphere at −80 °C prior to analysis.

X-ray absorption spectroscopy (XAS) on the Np $L_{III}$-edge were collected on the high Np(V) concentration samples from the pH 9.5 and 11 adsorption and crystallization experiments at the INE-Beamline at ANKA, Karlsruhe, Germany using a Ge(422) monochromator and 5-element Ge fluorescence detector (Table SI1).8,9 XAS analyses on the high Np(V) concentration pure hematite and goethite samples were performed at B18 at Diamond Light Source, UK, at the Np $L_{III}$-edge using a Si(111) monochromator and 9-element Ge

**METHODS**

**Safety.** Np-237 is a highly radiotoxic α-emitter with β- and γ-emitting daughter isotopes. Work with radionuclides must be carried out by appropriately qualified and experienced staff at facilities designed for work with radioactive materials. The possession and use of radioactive materials is subject to statutory control.

**Experimental Design.** The experiments were designed to bracket a range of elevated pH conditions where ferrihydrite was expected to crystallize to goethite, hematite and mixtures of these two minerals (Table SI1). Ferrihydrite was prepared according to the method of Cornell and Schwertmann.34,35 For the nonactive crystallization experiments 0.2 mM Ca(OH)$_2$ was mixed with ferrihydrite at a solid solution ratio of 0.5 g l$^{-1}$. The pH was adjusted to either 9.5 or 11 using 10 M NaOH, and the crystallization reactions were induced at elevated temperature (80 °C) for up to 2 weeks. Two additional experiments designed to produce pure hematite or goethite were also performed. Hematite crystallization was induced in pH 10.5 NaOH with a 5 g l$^{-1}$ ferrihydrite loading reacted at 120 °C for 4 weeks.35–37 Goethite crystallization was induced in 0.1 M NaOH, 0.1 M KOH, and 0.3 mM Ca(OH)$_2$, at pH 13.3 with a 0.5 g l$^{-1}$ ferrihydrite loading at room temperature for 2 weeks.36,37 Throughout, experiments were run in a CO$_2$-free atmosphere. The crystallization products were analyzed using a Bruker D8 X-ray diffractometer (XRD); Rietveld analyses were performed on the XRD patterns using Topas 4. Transmission electron microscopy images were taken on the final crystallization products using a Philips CM200 field emission gun TEM.

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During the XAS analyses, a apart from the goethite formed hematite Figures SI1 and SI2. In these spectra, two b Article. In these spectra, two b As intended, for the pure
The mineral crystallite sizes are plotted in Figure 1 Table 1. The actinide system (Table 1). The chemical extraction data showed significant differences in Np(V) desorption from the solids following crystallization (Table 1). For the adsorption samples, the majority of the Np(V) (84–94%) was released during the chemical extraction (Table 1) confirming that Np(V) was predominantly reversibly bound to the surfaces of the iron (oxyhydroxides). Conversely, only 6% of the Np(V) was released from the solid phase during the chemical extraction of the pure hematite sample (Table 1), indicating that the vast majority of the Np(V) was incorporated into hematite. During the chemical extraction of the crystallization products at pH 9.5 and 11 for 1 day (Table 1), 72 and 79%, respectively, of Np(V) remained associated with the solid product after the extraction. This fraction increased to 90 and 88%, respectively, after 14 days of crystallization at pH 9.5 and 11 (Table 1). The significant proportion of nonextractable Np(V) in these experiments suggested that incorporation processes dominated. Additionally, the increase in nonextractable Np(V) with time indicates that incorporation increased during aging similar to that observed for U(VI) incorporation into hematite.33 In contrast, the chemical extraction on the pure goethite sample at pH 13.3 released 78% of the Np(V) into solution (Figure 1, black arrows, and Table S1). This fraction increased to 90 and 88%, respectively, after 14 days of crystallization at pH 9.5 and 11 (Table 1). The significant proportion of nonextractable Np(V) in these experiments suggested that incorporation processes dominated. 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species, a shift in the energy position of the resonance features is inversely proportional to the axial and equatorial oxygen distances as observed for U(VI). The positions of the two resonance features in the XANES spectra from the Np(V) adsorption samples match those of the Np(V) solution standard, indicating the Np(V) in these samples exhibit neptunyl-like coordination. In contrast, for the pure hematite sample, the first (17620.3 eV) and second (17656.0 eV) resonance features were clearly shifted to lower and higher energies, respectively, relative to the XANES from the aqueous neptunyl standard (Figure 1 and Table SI2). The XANES from the pure hematite sample is very similar to recent data from Na-neptunate (Na₃NpO₄; Figure 2 and Table SI2), suggesting a change to a more neptunate-like Np(V) coordination environment during incorporation into hematite. In the XANES spectra from the crystallization samples at pH 9.5 and 11 and the pure goethite sample, the resonance feature positions appear intermediate between the Np adsorbed samples, and the neptunate-like pure hematite sample and Na-neptunate standard (Figure 1 and Table SI2). This reveals a change in the Np(V) local bonding environment during the crystallization of hematite, goethite, or both minerals.

**Np(V) Adsorption to Iron(Oxyhydr)Oxides.** The best fits to the EXAFS spectra from the adsorption samples indicate Np(V) is present within a 7-fold oxygen coordination environment with a neptunyl like configuration (Figure SI3 and Table 2) consistent with the XANES data. The modeled Np−O₅ bond lengths ranged from 1.83 to 1.85 Å with 5 equatorial oxygens at a distance of 2.38 to 2.42 Å, confirming essentially the same neptunyl-like coordination environment in all of the adsorbed samples. The equatorial oxygen bond lengths are marginally shorter than those observed for Np(V) adsorbed to hematite and goethite at pH 7−8.8 (between 2.46−2.51 Å). This could be caused by positively charged neptunyl forming stronger and consequently shorter bonds with negatively charged hydroxide than with water at the increased pH values of the experiment, which is consistent with modeling studies of uranyl complexation with hydroxide and water.

An additional Fe shell at 3.43 Å could also be fitted with statistical validity to the data from the ferrihydrite adsorption sample at pH 11 (Table 2). This is consistent with the adsorption of Np(V) via an inner sphere bidentate mononuclear surface complex at the ferrihydrite interface, and is in line with studies of Np(V) adsorption to other ferric (oxyhydr)oxide phases. In particular, Arai et al. identified the formation of mononuclear bidentate complexes at the hematite surface with a similar Np−O coordination environment and an essentially identical Np−Fe bond length to that observed here. It was not possible to fit Np−Fe shells to the...
Table 2. Details of the Fits to the Np L\textsubscript{III}-Edge EXAFS for Adsorption and Pure Hematite and Goethite Samples\textsuperscript{a}  

<table>
<thead>
<tr>
<th>sample</th>
<th>path</th>
<th>CN\textsuperscript{b}</th>
<th>R\textsuperscript{c}</th>
<th>\sigma\textsuperscript{d}</th>
<th>confidence (%)\textsuperscript{e}</th>
</tr>
</thead>
<tbody>
<tr>
<td>ferrihydrite adsorption pH 9.5</td>
<td>Np-O\textsubscript{ax}</td>
<td>2</td>
<td>1.85(2)</td>
<td>0.004(2)</td>
<td>92.1</td>
</tr>
<tr>
<td>R-factor = 0.018\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>5</td>
<td>2.42(4)</td>
<td>0.009(3)</td>
<td>92.1</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 11.0(35)\textsuperscript{b}</td>
<td>fit range: ( k = 3-8.5; R = 1-3.5 )</td>
<td>hematite/goethite ads. pH 9.5</td>
<td>Np-O\textsubscript{ax}</td>
<td>2</td>
<td>1.85(2)</td>
</tr>
<tr>
<td>R-factor = 0.0222\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>5</td>
<td>2.39(4)</td>
<td>0.008(2)</td>
<td>96.3</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 8.5(33)\textsuperscript{b}</td>
<td>fit range: ( k = 3-8.5; R = 1-3.5 )</td>
<td>ferrihydrite adsorption pH 11</td>
<td>Np-O\textsubscript{ax}</td>
<td>2</td>
<td>1.83(1)</td>
</tr>
<tr>
<td>R-factor = 0.0094\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>5</td>
<td>2.38(2)</td>
<td>0.007(2)</td>
<td>96.0</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 8.4(17)\textsuperscript{b}</td>
<td>fit range: ( k = 3-8.5; R = 1-4 )</td>
<td>goethite/hematite ads. pH 11</td>
<td>Np-O\textsubscript{eq}</td>
<td>2</td>
<td>1.85(2)</td>
</tr>
<tr>
<td>R-factor = 0.0180\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>5</td>
<td>2.39(3)</td>
<td>0.007(2)</td>
<td>96.1</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 9.2(36)\textsuperscript{b}</td>
<td>fit range: ( k = 3-8.5; R = 1-3.5 )</td>
<td>hematite</td>
<td>Np-O\textsubscript{eq}</td>
<td>1</td>
<td>1.86(1)</td>
</tr>
<tr>
<td>R-factor = 0.0084\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>2</td>
<td>2.16(1)</td>
<td>0.006(1)</td>
<td>99.9</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 3.6(16)\textsuperscript{b}</td>
<td>fit range: ( k = 3-15; R = 1.2-3.5 )</td>
<td>goethite</td>
<td>Np-O\textsubscript{eq}</td>
<td>3</td>
<td>2.38(1)</td>
</tr>
<tr>
<td>R-factor = 0.018\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>4</td>
<td>2.38(1)</td>
<td>0.008(1)</td>
<td>100</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 9.3(24)\textsuperscript{b}</td>
<td>fit range: ( k = 3-12.5; R = 1.3-3 )</td>
<td>goethite MS</td>
<td>Np-O\textsubscript{eq}</td>
<td>2</td>
<td>1.88(1)</td>
</tr>
<tr>
<td>R-factor = 0.018\textsuperscript{b}</td>
<td>Np-O\textsubscript{eq}</td>
<td>2</td>
<td>2.23(2)</td>
<td>0.004(2)</td>
<td>98.8</td>
</tr>
<tr>
<td>\Delta E\textsubscript{f} = 9.3(24)\textsuperscript{b}</td>
<td>fit range: ( k = 3-12.5; R = 1.3-3 )</td>
<td>goethite</td>
<td>Np-O\textsubscript{eq}</td>
<td>4</td>
<td>2.41(2)</td>
</tr>
</tbody>
</table>

\textsuperscript{a}The multiple scattering (MS) paths considered were the linear paths resulting from the Np–O\textsubscript{es} and the R and \sigma\textsuperscript{2} parameters were evaluated as multiples of the corresponding single scattering parameter path. The amplitude correction factor (S0\textsuperscript{2}) in all fits was fixed at 1. \textsuperscript{b}CN refers to the coordination number, R is the radial distance of the scattering path, \sigma\textsuperscript{2} is the Debye–Waller factor, the R-factor is the goodness of fit parameter, \Delta E\textsubscript{f} refers to the shift in energy from the calculated Fermi level, and the numbers in parentheses are the uncertainties calculated by Artemis (in the absence of uncertainties the coordination numbers were manually varied and fixed at the lowest R-factor and the \sigma\textsuperscript{2} were linked to a previous scattering path). \textsuperscript{c}Confidence refers to the statistical validity that the added scattering path improved the fit, calculated using the F-test for EXAFS.\textsuperscript{44} 

Data from the remaining adsorption samples due to restricted data quality and/or, in the case of adsorption to the crystallization products at pH 9.5 and 11, the association of Np(V) with multiple surface sites offered by hematite and goethite (Table 1).

\textbf{Np(V) Incorporation into Hematite.} In the pure hematite sample, the chemical extraction showed that more than 90% of the Np(V) was incorporated into hematite particles. Modeling of the EXAFS spectra resulted in a best fit indicating a Np(V) species coordinated to 6 oxygen atoms with a range of Np–O distances from 1.86–2.39 Å (Figure 2 and Table 2). The Np(V) species could not be represented as a simple dioxygenyl species with two short Np–O\textsubscript{eq} scattering paths. Rather, the best fit included a single short oxygen at 1.86 Å and two further oxygen shells at 2.16 and 2.38 Å. This suggests significant distortion in the overall Np–O local environment in pure hematite when compared to the neptunyl coordination environment of adsorbed and aqueous Np(V).\textsuperscript{49} As EXAFS represents average coordination, it is possible that the fit represents an average of multiple Np–O local environments. The single oxygen at 1.86 Å could indicate that a fraction of the Np(V) has short axial oxygens in neptunyl coordination, with the remaining Np(V) having longer Np–O\textsubscript{es} bond lengths. Interestingly, the Np–O\textsubscript{eq} coordination environment in the hematite sample (2 × 2.16 and 3 × 2.38 Å; Table 2) seems to closely match the Np–O coordination in sodium neptunate (2 × 2.09, 2 × 2.30, and 2 × 2.43 Å),\textsuperscript{50} and the Np–O\textsubscript{eq} bond lengths in the hematite sample are significantly shorter than those observed for the adsorbed, neptunyl species (2.38–2.42 Å, Table 2). Overall, the similarities of the Np–O coordination environment and the resonance feature positions in the XANES with sodium neptunate (Figure 1 and Table S12)\textsuperscript{50} confirm that Np(V) is incorporated into the structure of hematite in a distorted neptunyl-like coordination environment. Such a coordination environment has also been observed for the incorporation of U(VI) within hematite.\textsuperscript{31–33} However, the published U–O\textsubscript{eq} bond distances when U is incorporated into hematite (2.07 and 2.23 Å)\textsuperscript{33} are 0.09 and 0.15 Å shorter than the Np–O\textsubscript{eq} bond distances we observe (2.16 and 2.38 Å, Table 2). This is despite the fact that the crystal radii of 6-fold coordinated U(VI) (0.87 Å) and Np(V) (0.89 Å) are almost identical.\textsuperscript{59} Choppin et al.\textsuperscript{60} report that the effective charge of Np(V) in uranyl and neptunyl species are +3.3 and +2.2, respectively. The reduced effective charge of Np(V) compared to U(VI) could account for the slightly elongated Np–O\textsubscript{eq} bonds when compared to U–O\textsubscript{eq} bond distances when structurally incorporated within hematite.

In addition to the neptunyl-like oxygen scattering paths, iron scattering paths at four distances were included in the best fit of the EXAFS data from this sample. The Np–Fe bond
lengths (2.88, 3.12, 3.37, and 3.57 Å, with coordination numbers of: 1, 3, 3 and 3, respectively; Table 2) are within 5% of Fe−Fe distances in hematite (2.90, 2.97, 3.36, and 3.71 Å, Table S1), and similar to the U−Fe distances (2.87, 3.11, 3.45, and 4.01 Å) when U is substituted for Fe(III) in hematite. The combination of the distorted Np−O coordination environment and the similarity of the Np−Fe to the Fe−Fe and U−Fe coordination environment in hematite confirms that Np(V) is directly substituting for octahedrally coordinated Fe(III) in the hematite structure (Figure 2c), via a similar mechanism to U(VI). The reduced coordination number of the most distant Fe shell, from 6 to 3 compared to the Fe and U(VI) coordination in hematite (Table 2 and S13), is too large to be caused by the creation of vacancies to compensate charge (see below). Rather, this reduced coordination may have resulted from the need to constrain the Debye−Waller factors for the Np−Fe scattering paths to the same value in the fitting for this sample (Table 2). Alternatively, the low coordination number may be due to higher concentration of Np(V) near to the surface of the hematite particles with a resultant reduction in the Fe coordination number of the outer Fe shell.

The direct substitution of Np(V) for Fe(III) will lead to an excess positive charge which must be compensated for e.g. via vacancy creation or changes in the Fe oxidation state. To test for the creation of Fe(III) vacancies, we systemically decreased the coordination numbers of each Fe shell by 1. In every case, this resulted in a significant worsening of the fit. This suggests that if charge compensation occurs via the creation of Fe vacancies, that these are distributed randomly within the structure. In addition, there seems to be no clear mechanism to achieve charge compensation by reduction of Fe(III) in our fully oxidized system, therefore the exact mechanism of charge compensation remains elusive. Overall, the direct substitution of Np(V) for Fe(III) in the hematite structure highlights a novel pathway for Np(V) sequestration in natural and engineered systems where hematite forms which could significantly decrease the environmental mobility of this highly radionuclide actinide.

**Np(V) Partitioning during Mixed Hematite/Goethite Crystallization at pH 9.5 and 11.** In the crystallization experiments at pH 9.5 and 11, chemical extraction data indicate > 79% of the Np(V) incorporated into the mineral following ferrihydrite crystallization to hematite and goethite (Table 1). In addition, the energy position of the resonance features in the XANES spectra became more like the pure hematite sample as increasing incorporation of Np(V) into the mineral assemblage occurred (Figure 1 and Table S12). This suggests a systematic progression toward a more neptunian-like coordination environment. However, as the fraction of goethite increased in the crystallization samples (Table 1), the EXAFS spectra from these mixed crystallization experiments deviated from the EXAFS from the pure hematite sample (Figure SI4). Combined with the chemical extraction data highlighting increased incorporation into the crystallization products in these systems, this suggests that a fraction of Np(V) is incorporated into the structure of goethite. In addition, it indicates that the EXAFS includes contributions from Np(V) incorporated into both hematite and goethite. To explore the nature of the Np(V) interaction with goethite further, we extracted a residual EXAFS spectrum from the mixed crystallization samples by subtracting the adsorbed and pure hematite components from each EXAFS spectrum (Section S11 and Figure S15). Subsequent fitting of the resultant summed and normalized residual EXAFS resulted in a best fit with two Np−O eq scattering paths at 1.87 ± 0.02 Å and 4 Np−O eq scattering paths at 2.32 ± 0.03 Å (Table S15). Interestingly, the Np−O eq bond distances are in the range of Np−O eq in hematite (Np−O eq = 2.16 and 2.38 Å, Table 2) and shorter than those observed in the adsorbed Np(V) (Np−O eq = 2.38−2.42 Å, Table 2). This shortening appears consistent with a more neptunian-like coordination environment confirming the observations from the XANES for these mixed systems. Overall, these data suggest that Np may also be incorporated into goethite at pH 9.5 and 11 with a neptunian-like coordination environment, although its precise location within the overall structure of goethite remains unresolved.

**Np(V) Partitioning during Goethite Crystallization at pH 13.3.** Increasing the pH of the crystallization from 11 to 13.3 resulted in a change in the partitioning of Np(V), with only 22% (Table 1) nonextractable Np(V) determined from the chemical extraction. In addition, the crystallization product transitioned to pure goethite with a significantly smaller crystallite size (15.7 nm; Table 1) than the goethite formed at lower pH (≥40 nm; Table 1). Subsequent EXAFS analyses of this sample (Figure 2 and Table 2), showed a best fit of Np(V) coordinated to 8 oxygens, with two axial oxygen backscatterers at 1.88 Å and two equatorial oxygen shells at 2.23 and 2.41 Å with coordination numbers of 2 and 4, respectively (Table 2). This Np(V) coordination environment is nearly identical to Np(V) in NpO2(OH) (Table 2 and S13), suggesting that the Np(V) may be present as a discrete Np(V) hydroxide phase. However, thermodynamic modeling for this system (PHREEQC, Table S11) showed undersaturation with respect to NpO2(OH) (SI: −0.42) at 4 μM Np(V) and pH 13.3 (Table S11), while the solution was predicted to be supersaturated with respect to Cango3NpO2(OH)2 (SI: 3.37). This suggests that a discrete Np phase which is soluble in the low pH chemical extraction may have formed although the exact phase is unknown. The precipitation of a U(VI) hydroxide has also been observed under similar high pH conditions (NaUO2(OH)2·H2O) and this prevented the incorporation of U(VI) into the structure of goethite. This suggests that in these extreme pH systems, Np(V) (and more generally An(VI)) incorporation into the structure of goethite may be inhibited through the formation of discrete hydroxide precipitates. However, we cannot rule out the possibility that in this sample some of the Np(V) was adsorbed to goethite, and that the decreased crystal growth at pH 13.3 reduced the fraction of Np(V) incorporated due to limited overgrowth of the adsorbed Np. Overall, the lack of incorporation indicates Np(V) may remain labile under these high pH conditions with the aqueous concentration controlled by the Np(V)-phase solubility. Due to the discrepancy between the PHREEQC modeling and phase identity indicated by EXAFS, it is clear further work is required to determine Np(V) phase stability under high pH conditions and to generate accurate solubility data for these phases.

### Implications for Neptunium Behavior in the Environment

Here, we define the fate of Np(V) during interactions with iron(III) (oxyhydr)oxide phases, including during crystallization of goethite and hematite, under a range conditions relevant to natural and engineered environments. Specifically, we show direct evidence for Np(V) adsorption to ferrihydrite via an inner sphere bidentate neptunyl surface complex. Following crystallization, this adsorbed Np(V) incorporates into hematite.
Np(V) is not incorporated during iron oxide crystallization processes because hematite forms at lower temperature, not above 60 °C 

By contrast, under conditions where hematite is actively forming, potentially stabilizing the contaminants for extended timeframes, especially systems where hematite is actively forming, potentially extending the number of phases and the geochemical conditions under which the incorporation of Np(V) can occur. By contrast, under extremely high pH conditions (>13), which are relevant to cement systems, Np(V) is not incorporated during iron (oxyhydr)oxide crystallization probably due to the precipitation of a discrete Np(V) hydroxide phase. Overall, we provide valuable new insights into Np(V)/iron oxide interactions which will open new pathways to remediate radionuclide contamination issues in Hanford soils and sediments: Discharges from the Z-Plant (PFP) complex. Physics and Chemistry of the Earth, Parts A/B/C 2010, 35 (6-8), 292–297.


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