

Solvent effects on Grubbs' pre-catalyst initiation rates†

Cite this: *Dalton Trans.*, 2013, **42**, 4110Ian W. Ashworth,^a David J. Nelson^{‡*b} and Jonathan M. Percy^{*b}

Received 13th October 2012,

Accepted 8th November 2012

DOI: 10.1039/c2dt32441e

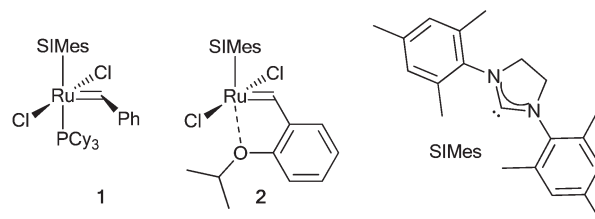
www.rsc.org/dalton

Initiation rates for Grubbs and Grubbs–Hoveyda second generation pre-catalysts have been measured accurately in a range of solvents. Solvatochromic fitting reveals different dependencies on key solvent parameters for the two pre-catalysts, consistent with different mechanisms by which the Grubbs and Grubbs–Hoveyda pre-catalysts initiate.

The alkene metathesis reaction is now a standard transformation in academic laboratories, and has been applied to the synthesis of a wide range of natural^{1,2} and unnatural products,³ fine chemicals,^{4,5} and polymers.⁶ The availability of robust and commercially available pre-catalysts such as **1** and **2** has enabled the rapid growth of alkene metathesis in the synthetic repertoire. Although transition metal catalysts have enabled many efficient large scale processes to be carried out in industry,^{5,7} the application of alkene metathesis to industrial processes has been more limited.^{7,8}

For industrial-scale syntheses, the implications of the reaction solvent must be considered carefully; these include costs of purchase, purification, drying, recycling and/or disposal, and the health and safety implications of transport, transfer, storage and use. There are also sustainability issues raised by projected uncertainty of supply and by legislative changes.⁹ These considerations have led major pharmaceutical companies to encourage their discovery chemists to anticipate scale-up in their laboratory practice and reaction design,⁹ replacing solvents which are problematic for scale-up, wherever practicable, and at the earliest stage possible.

The classical set of solvents for RCM is dichloromethane (DCM), 1,2-dichloroethane (DCE), benzene and toluene; of these, toluene raises the fewest issues while the other three are problematic. It has been reported that other solvents, including acetic acid,¹⁰ methyl *tert*-butyl ether (MTBE),¹¹ dimethyl carbonate¹² and hexafluorobenzene (HFB)^{13,14} are particularly effective for RCM, mostly on the basis of reaction yields or qualitative comparisons of kinetic profiles, so the precise locus of any solvent effect is not clear. If solvent effects on reaction chemistry could be revealed in detail, solvents could be selected for scale-up on the basis of both chemical efficacy and sustainability, and from a strong experimental starting point. We therefore sought to reveal the effects of solvent on the initiation rates of **1** and **2**, which are currently the most popular metathesis pre-catalysts.



Initiation rates were measured for **1** and **2** by reaction of the pre-catalyst with ethyl vinyl ether¹⁵ in a number of solvents. Initiation rates for **1** were obtained following the method of Sanford *et al.*;¹⁶ the rate-determining step in this reaction is known to be dissociation of the phosphane ligand. The integral *versus* time data was processed using a simple first order treatment (eqn (1)).

$$\ln[1] = \ln[1]_0 - k_{\text{obs}}t \quad (1)$$

For solvents where deuterated analogues were unavailable commercially, 10% v/v chloroform-*d* was added to enable a deuterium lock.

Initiation rates for pre-catalyst **1** cover only a *ca.* four-fold spread of values (Table 1), with similar values obtained in MTBE, dimethyl carbonate and DCM. Amongst the most frequently used solvents, the order of initiation rate constants for

^aGlobal Research and Development, AstraZeneca, Silk Road Business Park, Charter Way, Macclesfield, SK10 2NA, UK

^bWestCHEM Department of Pure and Applied Chemistry, University of Strathclyde, 295 Cathedral Street, Glasgow, G1 1XL, UK. E-mail: jonathan.percy@strath.ac.uk; Tel: +44 (0) 141 548 4398

†Electronic supplementary information (ESI) available: Initiation rate data for **1** and **2** in benzene, chloroform, dichloromethane, 1,2-difluorobenzene, dimethyl carbonate, hexafluorobenzene, methyl *tert*-butyl ether, toluene, trifluorotoluene. See DOI: 10.1039/c2dt32441e

‡Current address: EaStCHEM, School of Chemistry, University of St. Andrews, North Haugh, St. Andrews, Fife, KY16 9ST, UK. E-mail: djn5@st-andrews.ac.uk

Table 1 Initiation rates for **1** (ca. 5–15 mmol L⁻¹ with 0.5 mol L⁻¹ ethyl vinyl ether) and **2** (0.1 mmol L⁻¹ with 25–200 mmol L⁻¹ ethyl vinyl ether) in various solvents at 298 K; deuterated solvents were used for measurements of the initiation of **1** unless otherwise stated

Solvent	1		2	
	$k_{\text{init}}/\text{s}^{-1}$	rel. k_{init}^a	$k_{\text{init}}/\text{L mol}^{-1} \text{s}^{-1}$	rel. k_{init}^a
Benzene	1.0×10^{-4}	0.71	0.0460	1.74
Chloroform	4.5×10^{-4}	0.32	0.0231	0.88
Dichloromethane	1.4×10^{-4}	1.00	0.0264	1.00
1,2-Difluorobenzene	1.4×10^{-4}	1.00	0.0390	1.48
Dimethyl carbonate	1.5×10^{-4}	1.07	0.0411	1.56
Hexafluorobenzene	3.3×10^{-5}	0.24	0.0515	1.95
Methyl <i>tert</i> -butyl ether	1.3×10^{-4}	0.93	0.0592	2.24
Toluene	9.2×10^{-5}	0.66	0.0509	1.93
Trifluorotoluene	1.3×10^{-4}	0.93	0.0446	1.69

^a Relative to the initiation rate in DCM. ^b Reported in ref. 26.

^c Contained 10% v/v chloroform-*d* to enable a deuterium lock.

1 was DCM > (benzene \approx toluene). Metathesis reactions are often conducted in solvents such as DCM or toluene, and DCM has become the default solvent for laboratory scale synthetic applications. It is far from ideal for use on a large scale due to safety, toxicity and environmental concerns whereas MTBE and dimethyl carbonate are not only more industrially acceptable solvents⁹ but also support initiation at similar rates to DCM. Notably, the slowest initiation rate was obtained in hexafluorobenzene, which has been reported to be a superior solvent for metathesis reactions (*vide infra*).^{13,14}

Initiation rate constants for pre-catalyst **2** were measured (Table 1) according to the published procedure;¹⁷ while Plenio *et al.* have reported non-linear plots of k_{obs} versus [substrate] in toluene, all of our plots, which cover the typical concentration range for metathesis (25–200 mmol L⁻¹), were linear.¹⁸ A modest (ca. 2-fold) spread of values covered chloroform to MTBE; this represented a range of ΔG^\ddagger values of only ca. 0.6 kcal mol⁻¹. The order of initiation rate constants for **2** amongst the most frequently used solvents was PhMe > PhH >> DCM and initiation rates in aromatic solvents are typically higher than those in most non-aromatic solvents. In contrast to the results for **1**, the initiation rate constant of **2** in hexafluorobenzene is toward the higher end of the range obtained. The trends described here are entirely qualitative (for the PCA context of this work,¹⁹ see the ESI[†]) so we sought a way of correlating measured solvent parameters to initiation rates. Samojłowicz *et al.* attempted to link metathesis performance and solvent dielectric constant¹⁴ after suggestions by Sanford *et al.* that there may be a relationship.¹⁶ However, no link could be established; plots of our initiation rates versus solvent dielectric constants²⁰ are scattered and trendless (Fig. 1).

The concept of solvent polarity is a complex one, combining microscopic (such as hydrogen-bond donor or acceptor ability) and macroscopic (such as the dielectric constant) properties of a solvent.²¹ The $E_{\text{T}}(30)$ scale²⁰ is based upon the wavelength of absorption of solvatochromatic dye **3** and has found applications in analytical chemistry (where it can be used to predict

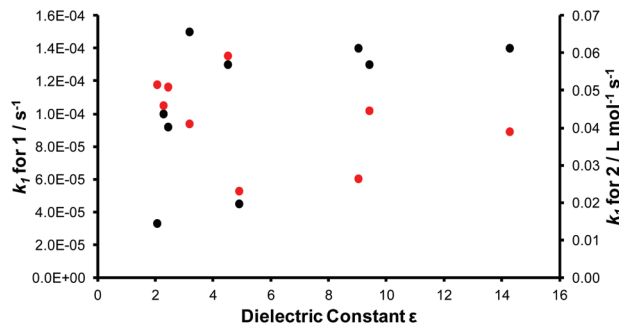


Fig. 1 Initiation rate constant versus solvent dielectric constant²⁰ for the initiation of **1** (red) and **2** (black) with ethyl vinyl ether at 298 K.

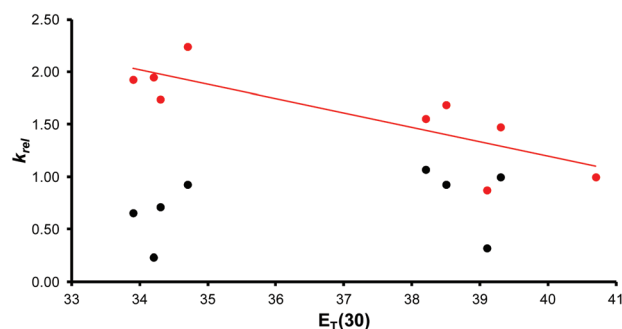
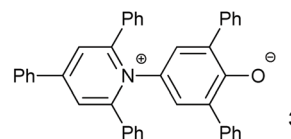


Fig. 2 Dimroth–Reichardt plots for the initiation rates of **1** (black) and **2** (red).

changes in retention time, and to measure solvent water content)²² and chemical kinetics (to predict changes in the rates of S_N2 reactions²³ and ligand exchange processes²⁴). $E_{\text{T}}(30)$ values are available for all nine solvents in which we measured initiation rate constants;²⁰ however, there was no trend at all for pre-catalyst **1** and only a scattered dependence with a negative slope for pre-catalyst **2** (Fig. 2).

The use of Kamlet–Taft–Abboud (TKA) solvatochromatic parameters was explored next;²⁵ this approach was used previously by Adjiman *et al.*¹⁰ but with unreliable rate constants.²⁶ Firstly, we treated the data with eqn (2) but found that poor correlations obtained. Parameters for use in the expanded correlation used by Adjiman are only available for five of the solvents studied here (benzene, chloroform, DCM, HFB and toluene), but if data for diethyl carbonate and di-*iso*-propyl ether are used to approximate dimethyl carbonate and methyl *tert*-butyl ether respectively,²⁰ a linear regression can be performed (eqn (3), Table 2, Fig. 3). Unfortunately, parameters are not available for 1,2-difluorobenzene and trifluorotoluene.

$$\log_{10}(k) = A\alpha + B\beta + C\pi^* + D \quad (2)$$

$$\log_{10}(k) = A\alpha + B\beta + C\pi^* + D\delta + E(\delta H)^2 + F \quad (3)$$

Table 2 Results of linear regression for **1** and **2** initiation rate constants in benzene, chloroform, dichloromethane, diethyl carbonate (for dimethyl carbonate), di-iso-propyl ether (for MTBE), hexafluorobenzene and toluene (see Table 1, eqn (3))

Pre-catalyst	A	B	C	D	$10^3 E$	F
1	-0.924 ± 0.0657	0.990 ± 0.0727	1.88 ± 0.0407	-0.170 ± 0.0278	-1.26 ± 0.118	-4.59 ± 0.0508
2	-0.793 ± 0.145	0.106 ± 0.161	0.0993 ± 0.0899	0.0487 ± 0.0614	-1.13 ± 0.261	-1.04 ± 0.112

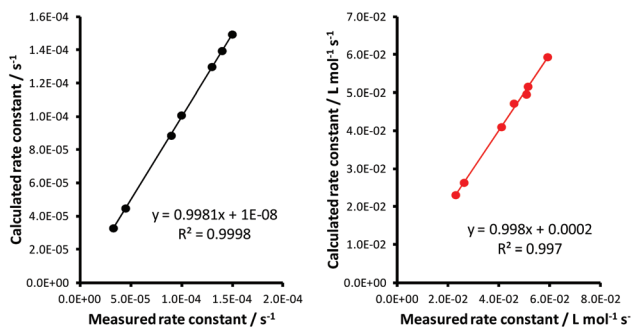


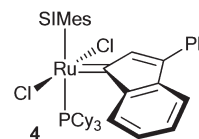
Fig. 3 Calculated rate constants for (a) **1** and (b) **2** (from the linear regression parameters in Table 2) versus measured rate constants (see Table 1).

From the coefficients (particularly *B* and *C*), it appears that the differences in the solvent effects on **1** and **2** initiation are due to significantly different dependencies on β and π^* ; **1** initiates faster in solvents with a high polarisability and low hydrogen bond donor ability, while for **2**, these parameters have a much lower influence. Clearly, solvent effects on the initiation of **2** are far smaller than those on the initiation of **1**. These values may also be useful predictively; initiation rate constants of $2.1 \times 10^{-4} \text{ s}^{-1}$ for **1** and $0.039 \text{ L mol}^{-1} \text{ s}^{-1}$ for **2** can be calculated in 1,2-dichloroethane, a popular solvent in the synthetic laboratory. This solvent not only offers an extended temperature range over DCM but may also have a higher intrinsic ability to support phosphane dissociation. Adjiman *et al.* proposed that acetic acid is a particularly effective solvent for metathesis reactions, and has some attractive features for scale-up.²⁷ While ethyl vinyl ether and acetic acid are incompatible chemically, preventing the measurement of the initiation rate, the values in Table 2 would predict slow initiation of **1** and **2**, with rate constants of $1.0 \times 10^{-5} \text{ s}^{-1}$ and $0.004 \text{ L mol}^{-1} \text{ s}^{-1}$ respectively. This suggests strongly that the successful preparative reaction of diethyl diallylmalonate described by Adjiman *et al.* is not due to rapid initiation of **1**. Slow initiation of both pre-catalysts would also be predicted in acetone and ethyl acetate.

There is a consensus that the initiation mechanism of **1** involves phosphane dissociation;¹⁶ the alkylidene rotates as the Ru...P distances increase, according to recent work by Jensen *et al.*²⁸ The dissociative mechanism for **2**, for which there is experimental and computational support,^{18,29} also involves an alkylidene rotation coupled with an extending Ru...O distance. The interchange mechanism supported by our computational work¹⁷ and also by Plenio's experimental data,¹⁸ is earlier with respect to the Ru...O extension, and includes the approach of an alkene molecule towards the Ru centre (EVE in the case of initiation) and some sharing of

alkene electron density with the Ru centre. The three types of transition states have different shapes, volumes and polarities so it is not surprising that there is no simple relationship between dielectric constant or $E_T(30)$ for either pre-catalyst, and that multi-parameter approaches are required.

The accurate accounting of solvent effects requires that both specific solvation and more continuum-based effects are represented in any model and the KTA empirical approach offers the prospect of revealing these effects. We are not aware of any organometallic reactions to which the approach has been applied so our findings represent a novel use of the methodology. In a more *a priori* approach, Samojłowicz *et al.* studied the effects of solvent on metathesis reactions;¹⁴ the authors noted that the yields of many RCM reactions (with pre-catalysts including **1** and **2**) were highest in perfluoroaromatic solvents, followed by aromatic hydrocarbon solvents, followed in turn by chlorinated aliphatic solvents. The authors argued that the facile dissociation/oxidation reaction of the phosphane ligand in C_6F_6 would allow faster and irreversible formation of catalytically active 14-electron species, which could subsequently be stabilised by π -stacking interactions between solvent molecules and the aryl groups of the NHC ligand. Our measured initiation rates are not consistent with faster initiation in the perfluoroarene solvent studied; initiation reactivity decreases in the order DCM > toluene \approx benzene > hexafluorobenzene for **1**, though initiation is faster in HFB for **2**. The differences in rates are relatively modest but even these small rate differences could cause quite large perturbations in reaction yields. Perfluorination is known to increase the solubility of oxygen in solvents; for example, oxygen is approximately twice as soluble in HFB as in benzene itself³⁰ so catalyst decomposition *via* oxidation should be faster in HFB. Grell showed that **1** decomposed more rapidly in a mixture of HFB and 1,2-DCE than in 1,2-DCE alone, consistent with this view. In contrast, pre-catalyst **4** (known as Umicore M2) was less air sensitive in HFB than in toluene; pre-catalysts of this type are highly robust and typically initiate much more slowly³¹ than **1** or **2**, so comparisons between the systems may be unsafe, but these contrasting findings do suggest that we need to improve our understanding of pre-catalyst and catalyst concentration behaviour throughout the duration of reactions in addition to quantifying the initiation event accurately.



Conclusions

In conclusion, we have presented accurately measured initiation rate constants for pre-catalysts **1** and **2** in a range of commonly-used solvents. These initiation rates could be used in reaction simulation studies of alkene metathesis, to simulate concentration/time profiles in different solvents.²⁶ MTBE and dimethyl carbonate, which are much more sustainable solvents, offer faster pre-catalyst initiation than default laboratory solvent DCM. The lack of even a qualitative link between solvent dielectric constant and initiation rate shows that this quantity cannot be used to predict pre-catalyst initiation rates; a KTA solvatochromic treatment has greater potential for this purpose. Further data in a wider variety of solvents, plus a better understanding of how to model the catalyst decomposition event, are required before more concrete conclusions about solvent effects can be made.

Notes and references

- 1 A. H. Hoveyda and A. R. Zhugralin, *Nature*, 2007, **450**, 243–251.
- 2 K. C. Nicolau, P. G. Bulger and D. Sarlah, *Angew. Chem., Int. Ed.*, 2005, **44**, 4490–4527.
- 3 E. N. Guidry, S. J. Cantrill, J. F. Stoddart and R. H. Grubbs, *Org. Lett.*, 2005, **7**, 2129–2132.
- 4 S. Monfette, M. Eyholzer, D. M. Roberge and D. E. Fogg, *Chem.–Eur. J.*, 2010, **16**, 11720–11725.
- 5 J. Magano and J. R. Dunetz, *Chem. Rev.*, 2011, **111**, 2177–2250.
- 6 C. W. Bielawski and R. H. Grubbs, *Prog. Polym. Sci.*, 2007, **32**, 1–29.
- 7 C. A. Busacca, D. R. Fandrick, J. J. Song and C. H. Senanayake, *Adv. Synth. Catal.*, 2011, **353**, 1825–1864.
- 8 V. Farina, C. Shu, X. Zeng, X. Wei, Z. Han, N. K. Yee and C. H. Senanayake, *Org. Process Res. Dev.*, 2009, **13**, 250–254.
- 9 R. K. Henderson, C. Jimenez Gonzalez, D. J. C. Constable, S. R. Alston, G. G. A. Inglis, G. Fisher, J. Sherwood, S. P. Binks and A. D. Curzons, *Green Chem.*, 2011, **13**, 854–862.
- 10 C. S. Adjiman, A. J. Clarke, G. Cooper and P. C. Taylor, *Chem. Commun.*, 2008, 2806.
- 11 K. M. Kuhn, T. M. Champagne, S. H. Hong, W.-H. Wei, A. Nickel, C. W. Lee, S. C. Virgil, R. H. Grubbs and R. L. Pederson, *Org. Lett.*, 2010, **12**, 984.
- 12 X. Miao, C. Fischmeister, C. Bruneau and P. H. Dixneuf, *ChemSusChem*, 2008, **1**, 813.
- 13 C. Samojłowicz, M. Bieniek, A. Zarecki, R. Kadyrov and K. Grella, *Chem. Commun.*, 2008, 6282.
- 14 C. Samojłowicz, M. Bieniek, A. Pazio, A. Makal, K. Wozniak, A. Poater, L. Cavallo, J. Wojeik, K. Zdanowski and K. Grella, *Chem.–Eur. J.*, 2011, **17**, 12981.
- 15 J. Louie and R. H. Grubbs, *Organometallics*, 2002, **21**, 2153.
- 16 M. S. Sanford, J. A. Love and R. H. Grubbs, *J. Am. Chem. Soc.*, 2011, **123**, 6543.
- 17 I. W. Ashworth, I. H. Hillier, D. J. Nelson, J. M. Percy and M. A. Vincent, *Chem. Commun.*, 2011, 5428.
- 18 V. Thiel, M. Hendann, K.-J. Wannowius and H. Plenio, *J. Am. Chem. Soc.*, 2012, **134**, 1104.
- 19 F. P. Ballistreri, C. G. Fortuna, G. Musumarra, D. Pavone and S. Scire, *ARKIVOC*, 2002, **xi**, 54.
- 20 J. L. M. Abboud and R. Notari, *Pure Appl. Chem.*, 1999, **71**, 645.
- 21 A. R. Katritzky, D. C. Fara, H. Yang, K. Tamm, T. Tamm and M. Karelson, *Chem. Rev.*, 2004, **104**, 175.
- 22 B. P. Johnson, B. Gabrielsen, M. Matulenko, J. G. Dorsey and C. Reichardt, *Anal. Lett.*, 1986, **19**, 939.
- 23 A. Skrzypczak and P. Neta, *Int. J. Chem. Kinet.*, 2004, **36**, 253.
- 24 G. Gumbel and H. Elias, *Inorg. Chim. Acta*, 2003, **342**, 97.
- 25 M. J. Kamlet, J. L. M. Abboud and R. W. Taft, in *Progress in Physical Organic Chemistry*, John Wiley & Sons, Inc., 2007, pp. 485–630.
- 26 D. J. Nelson, D. Carboni, I. W. Ashworth and J. M. Percy, *J. Org. Chem.*, 2011, **76**, 8386–8393.
- 27 C. Capello, U. Fischer and K. Hungerbühler, *Green Chem.*, 2007, **9**, 927934.
- 28 Y. Minenkov, G. Occhipinti, W. Heyndrickx and V. R. Jensen, *Eur. J. Inorg. Chem.*, 2012, 1507–1516.
- 29 F. Nuñez Zarur, X. Solans Monfort, L. Rodríguez Santiago and M. Sodupe, *Organometallics*, 2012, **31**, 4203.
- 30 J.-J. Delpucha, M. A. Hamza, G. Serratrice and M.-J. Stebe, *J. Chem. Phys.*, 1979, **70**, 2680–2687.
- 31 H. Clavier, C. A. Urbina Blanco and S. P. Nolan, *Organometallics*, 2009, **28**, 2848.